

## Supporting Information

### Synthesis and Degradation Study of Cationic Polycaprolactone-Based Nanoparticles for Biomedical and Industrial Applications

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**Table S1:** PCL<sub>2</sub>-based NPs: average particle size (D<sub>p</sub>), polydispersity index (PDI) and ζ-potential of the NPs obtained through BEP

<b>Macromonomer</b>	<b>Surfmer</b>	<b>S/M</b>	<b>PEGMA/M</b>	<b>Dp</b>	<b>PDI</b>	<b>ζ-POT</b>
		[w/w]	[w/w]	[nm]		[mV]
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	5 %	20%	115	0.045	+36
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	10 %	20 %	90	0.063	+40
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	15%	20 %	80	0.193	+43

**Table S2:** PCL<sub>2</sub>-based NPs: average particle size (Dp), polydispersity index (PDI) and ζ-potential of the NPs obtained through MSSEP

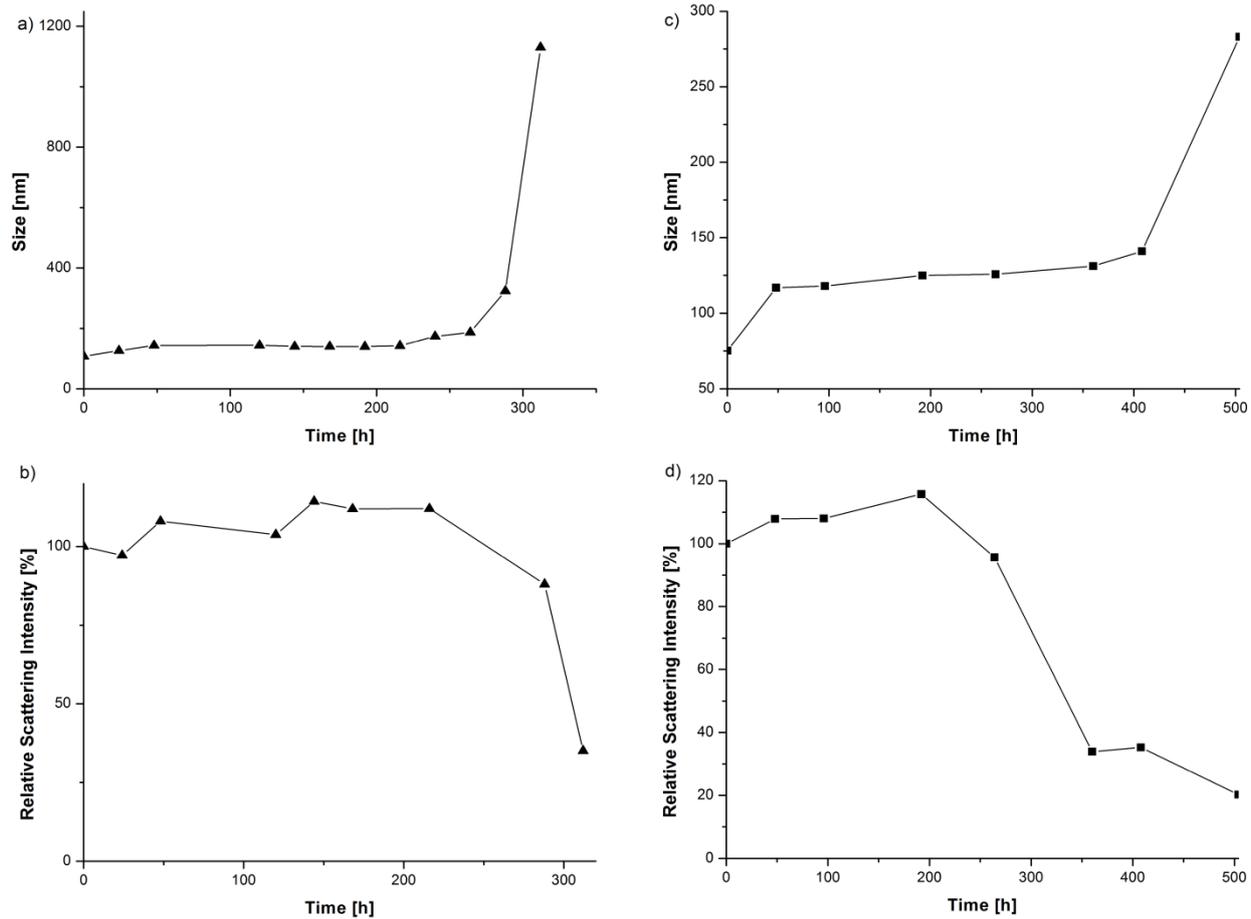
<b>Macromonomer</b>	<b>Surfmer</b>	<b>S/M</b>	<b>PEGMA/M</b>	<b>Dp</b>	<b>PDI</b>	<b>ζ-POT</b>
		[w/w]	[w/w]	[nm]		[mV]
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	2 %	20%	75	0.093	+34
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	8 %	20 %	64	0.105	+40
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	10 %	20 %	63	0.107	+41

**Table S3:** PCL<sub>n</sub>-based NPs (n=2,3,5): average particle size (Dp), polydispersity index (PDI) and ζ-potential of the NPs obtained through MSSEP.

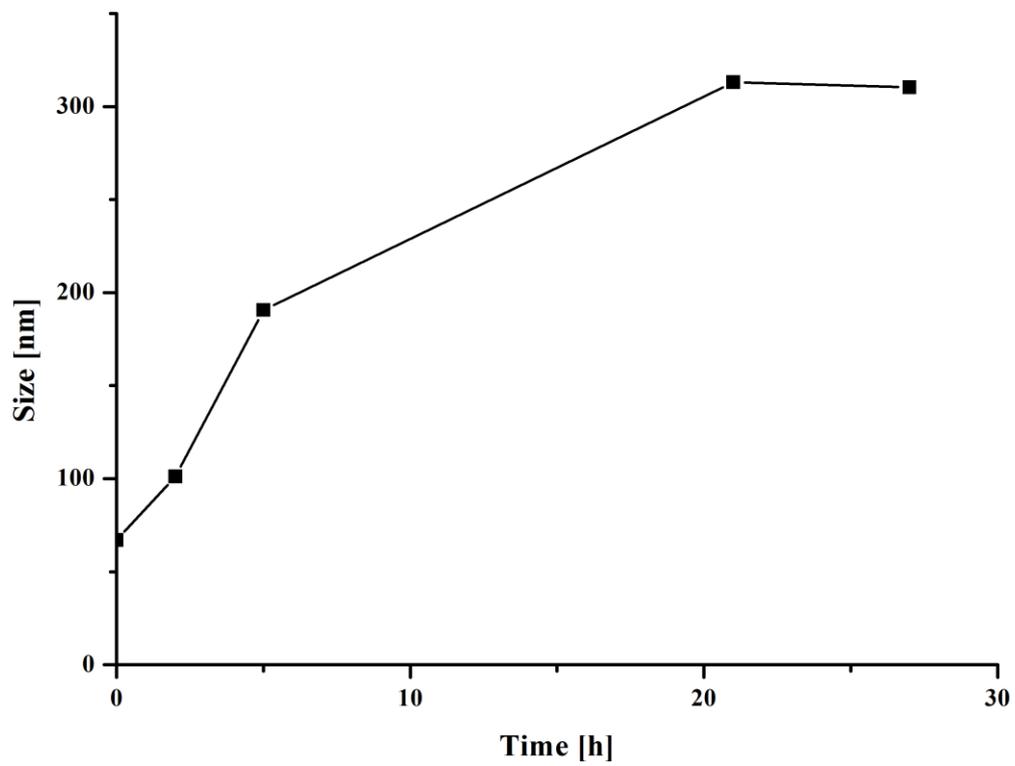
<b>Macromonomer</b>	<b>Surfmer</b>	<b>S/M</b>	<b>PEGMA/M</b>	<b>Dp</b>	<b>PDI</b>	<b>ζ-POT</b>
		[w/w]	[w/w]	[nm]		[mV]
<b>PCL<sub>2</sub>MA</b>	PCL <sub>2</sub> ChMA	10%	20%	63	0.107	+41
<b>PCL<sub>3</sub>MA</b>	PCL <sub>3</sub> ChMA	10 %	20%	56	0.158	+39
<b>PCL<sub>5</sub>MA</b>	PCL <sub>5</sub> ChMA	10 %	20%	41	0.197	+40

**Table S4:** PCL<sub>n</sub>ChMA based NPs: average particle size (Dp), polydispersity index (PDI) and ζ-potential of the NPs obtained through BEP

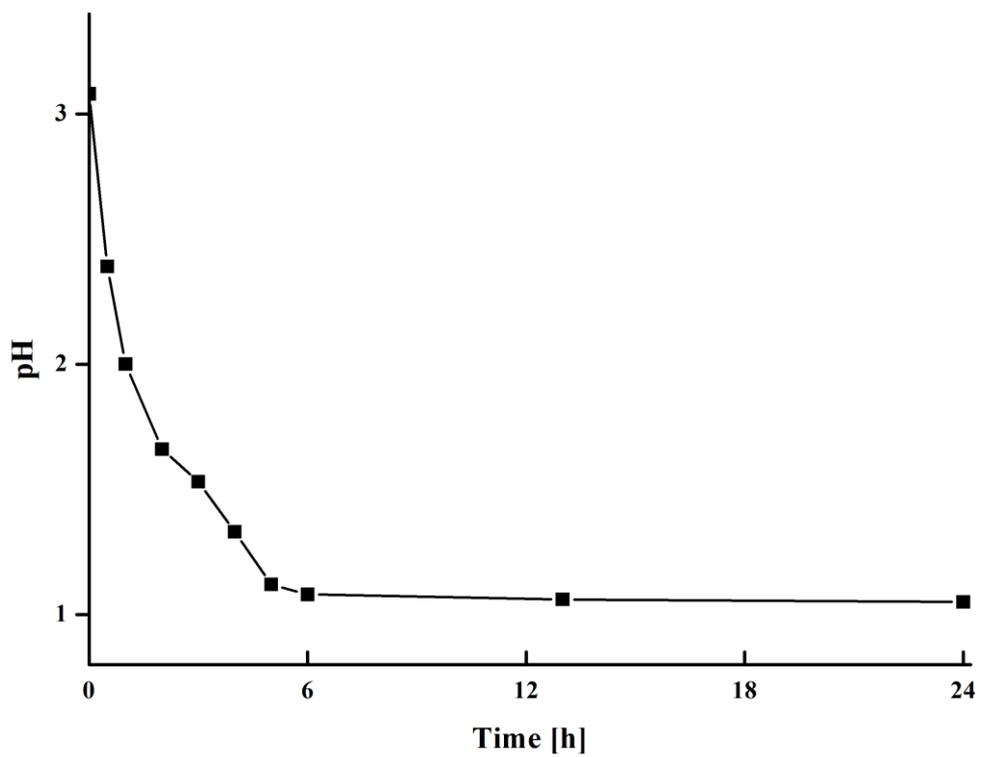
<b>Surfmer</b>	<b>PEGMA/M</b> <b>[w/w]</b>	<b>Dp</b> <b>[nm]</b>	<b>PDI</b>	<b>ζ-POT</b> <b>[mV]</b>
<b>PCL<sub>2</sub>ChMA</b>	20%	275	0.103	+54
<b>PCL<sub>3</sub>ChMA</b>	20 %	115	0.1	+51
<b>PCL<sub>5</sub>ChMA</b>	20 %	85	0.175	+49



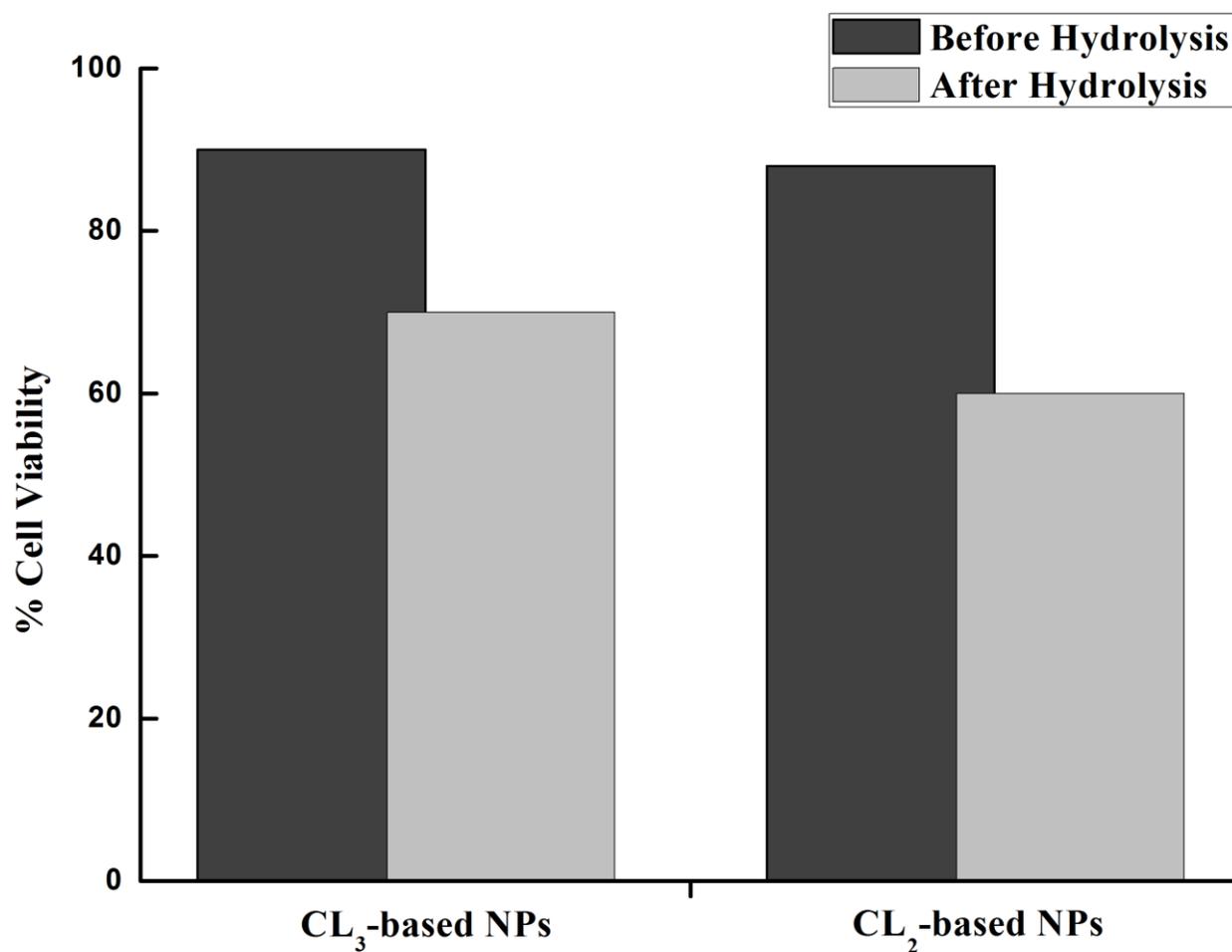
**Figure S1:** Evolution of the relative scattering intensity and the NP size for PCL<sub>2</sub>-NPs (a,b) and PCL<sub>3</sub>-NPs (c,d) in cell medium over the time at 37°C.



**Figure S2:** Variation of the PCL<sub>3</sub>-NP size during the time at 90°C in water.



**Figure S3:** Variation of the pH over the time for PCL<sub>3</sub>-NPs at 90°C.



**Figure S4:** Cell viability before and after hydrolysis of two different CL<sub>n</sub>-based NPs.

The higher toxicity after the hydrolysis is related to the pH decrease, according to the release of carboxylic acids. It is important to notice that this effect is strongly mitigated in vivo for the buffered environment.