Supporting Information for

A Structurally Perfect S = 1/2 Metal-Organic Kagomé Antiferromagnet

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Experimental Section

Sample Preparation. Water was distilled and deionized with a Milli-Q filtering system. All other reagents were obtained from commercial vendors and used without further purification: Cu(OH)₂ (99.5+%, Noah Technologies); Cu(OH)F (99.999%, Alfa Aesar); imidazole (99+%, Aldrich); and isophthalic acid (99%, Aldrich). Hydrothermal reactions were carried out under autogeneous pressure in 23 mL Teflon-lined pressure vessels purchased from Parr Instruments; large-scale sample preparation was carried out in a custom-built 1900 mL vessel. Fisher Isotemp programmable ovens with forced-air circulation were used to obtain the desired temperature profiles for the hydrothermal reactions. Elemental analysis was performed by H. Kolbe Mikroanalytisches Laboratorium.

Cu(1,3-bdc) (1). A 1900 mL liner was charged with 6.04 g (61.9 mmol) of Cu(OH)₂, 2.06 g (20.6 mmol) of Cu(OH)F, 13.71 g (82.5 mmol) of isophthalic acid, 2.81 g (41.3 mmol) of imidazole, and 410 mL of water, capped and placed into a steel hydrothermal bomb. The bomb was heated to 150 °C at a rate of 1 °C/min, maintained at this temperature for 5 days, and cooled to room temperature at a rate of 0.1 °C/min. Blue hexagonal crystals of Cu(1,3-bdc) (1) lined the walls of the vessel and intimately covered a blue-green powder and were difficult to separate mechanically from the fine powder. The product mixture, therefore, was sonicated in its mother liquor in a 1000 mL glass bottle for 80 minutes. The suspension of powder was decanted, leaving behind blue hexagonal crystals of 1, and the powder was separated from the mother liquor by vacuum filtration. The powder was identified by pXRD to consist of starting materials (Cu(OH)₂, spertinite, PDF # 80-0656; Cu(OH)F, PDF #07-0306; and isophthalic acid, PDF #45-1541) and a green impurity, $C_{32}H_{24}Cu_6O_{26}$ (3), whose structure is detailed below. The filtrate mother liquor was recombined with the blue hexagonal crystals, sonicated for an additional 25 minutes, and powder suspension was again decanted from the crystals. The blue hexagonal crystals of 1 (13.3) g, ~71%) were then washed with deionized water, isolated by filtration, and dried in air. Impurities were further removed by physical separation under a microscope. The purity of the

crystalline Cu(1,3-bdc) (1) sample was judged by pXRD. The crystal structure of 1 contains a solvent-accessible void of 45 Å³, which can be occupied by an incidental solvent molecule. Another crystal within the same batch refined as Cu(1,3-bdc)·0.11 H₂O (2), with 1/9 of a water molecule per formula unit. Further crystallographic details are provided (*vide infra*). Thermogravimetric analysis (TGA) of the crystalline sample (*vide infra*) showed loss of 0.65% by mass (or ~0.08 H₂O per formula unit) to a temperature of 257 °C, followed by decomposition of the sample at 300 °C, leaving a residue of Cu₂O (PDF # 00-0865) and Cu metal (PDF # 04-0836) as determined by pXRD. Prior to magnetic measurements, samples were dried by heating them uncovered in a ceramic crucible at 10 °C/min, maintaining them at 240 °C for 2 h, and then cooling them to room temperature at 10 °C/min. No significant sample decomposition occurred, as judged by pXRD and elemental analysis. Anal. Calcd. for C₈H₄CuO₄: C, 42.21; H, 1.77; Cu, 27.91. Found: C, 41.90; H, 1.87; Cu, 27.74.

TGA (*vide infra*) of sample dried in this manner showed no significant loss of mass before sample decomposition, indicating removal of the incidental solvent.

X-ray Structure Determinations. Structures were determined for Cu(1,3-bdc) (1), Cu(1,3-bdc)·0.11 H₂O (2), and C₃₂H₂₄Cu₆O₂₆ (3). Single crystals were transferred onto a microscope slide coated with mineral oil. Crystals were affixed to a glass fiber or a cryoloop using the oil, frozen in a cold nitrogen gas stream, and optically centered. The data were collected on a Siemens three-circle platform goniometer equipped with a Bruker Smart Apex CCD detector with graphite-monochromated Mo K α radiation ($\lambda = 0.71073$ Å), using both phi and omega scans at -173 °C. Data reduction and unit cell refinements were performed with the program SAINT,¹ and the program SADABS² was used to perform scaling and absorption corrections. Space group assignments were based on systematic absences, E statistics, and successful refinement of the structures. The structures were solved by direct methods (SHELXS)³ and

¹ Bruker SAINT, Bruker-AXS Inc.: Madison, Wisconsin, USA, 2005.

² Sheldrick, G. M. SADABS, University of Göttingen, Germany, 2007.

³ Sheldrick, G. M. Acta Cryst. **1990**, A46, 467.

refined against F^2 on all data by full matrix least squares with SHELXL-97.⁴ All non-hydrogen atoms were refined anisotropically; all hydrogen atoms were placed at idealized positions and refined using a riding model.

The structures of **1** and **2** are identical with the exception of the partially occupied water molecule on a special position in the latter case. The two structures, however, are not strictly isostructural and it is not possible to use one model in a refinement against the other dataset. Although the *x* and *z* positions of all atoms in both structures are identical within error, the *y* coordinate of Cu(1) differs by 0.5, and the *y* coordinates of all other atoms are different. The additional water molecule is located on the crystallographic $\overline{3}$ axis (coordinates 0,0,0; Wykoff *b*); its occupancy was refined freely and eventually set at the convergence value of 0.322 × 0.16667.

The structure of the $C_{32}H_{24}Cu_6O_{26}$ byproduct (**3**) contains a 16-membered ring. The structure provides evidence that some 1,3-benzenedicarboxylate is oxidized in the 2-position under the given conditions. Excluding their bond to one another, Cu(1) and Cu(2) are each in pseudo-square planar coordination environments, each with two bonds to monodentate carboxylate oxygens and two bonds to the newly formed oxide. The Cu(1)–Cu(2) distance is 2.9714(5) Å, which is within bonding distance for two Cu(II) ions. Cu(3) is pseudo-square pyramidal; the base of the pyramid contains two bonds to monodentate carboxylate oxygens and two bonds to worker molecules, and the apex of the pyramid is a bonded water molecule. Remarkably, in this incomplete octahedral coordination sphere, no residual electron density can be found near the site of the missing ligand. The two Cu(3) atoms bridge the two Cu(1)–Cu(2) dimers in a ring-like structure in the formula unit. The compound crystallizes in the centrosymmetric space group *P*–*1*. The packing diagram viewed along the *c* axis of the unit cell shows that the 16-membered ring is not planar and the planes of the dimers are staggered with respect to one another.

⁴ Sheldrick, G. M. SHELXL 97, University of Göttingen, Germany, 1997.

Magnetic Measurements. DC magnetic susceptibility data were collected on (ground-up) crystalline samples contained in gelatin capsules using Quantum Design MPMS-5S and MPMS-XL SQUID magnetometers at temperatures ranging from 2 to 350 K and field strengths varying from –50 to 50 kOe. The data were corrected for diamagnetic contributions of the sample holder by measurement of an empty capsule, and of the sample itself by use of Pascal's constants.

Other Physical Measurements. Infrared spectra of samples in KBr pellets were recorded on a Nicolet Magna-IR 860 spectrometer equipped with a KBr beam splitter and a DTGS detector. Powder X-ray diffraction patterns were measured using a Rigaku RU300 rotating anode X-ray diffractometer with Cu K α radiation ($\lambda = 1.5405$ Å), which was wavelength-selected with a single-crystal graphite monochromator. Samples were spread onto a glass slide fixed with double-sided Scotch tape. Samples were rotated through $2\theta/\theta$ space and intensity was recorded as a function of 20 from 5 to 60°. Patterns were indexed with MDI Jade software version 7.0 and referenced using the JCPDS powder diffraction database. Specific heat measurements were performed using a Quantum Design Model 6000 Physical Properties Measurement System with Helium-3 option. This system uses relaxation method calorimetry and determines specific heat by fitting the temperature response to a two exponential model. A single crystal of mass 0.999 \pm 0.002 mg was held to the measurement platform using Apiezon N grease. The long time constant (τ_1) varied between 20 and 80 seconds for the temperature range measured. The sample coupling always exceeded 96%, implying that the ratio of the thermal conductance of the grease to that of the platform wires was at least 25. Thermogravimetric analysis measurements were performed using a Seiko Dual TG/DTA 320 Thermogravimetric/Differential Thermal Analyzer (SC). Samples were referenced and measured in aluminum pans, with N₂ as a purge gas flowing at 150 cc/min. Data were recorded at 0.2 s increments from 20 °C to 500 °C at a 5 °C/min heating rate.



Figure S1. FTIR of a crystalline sample of the title compound (a) before, and (b) after heating at 240 °C for 2 hours. Before heating, the sample shows an absorbance at 1693 cm⁻¹ due to a bending mode of solvent water. After heating, this absorbance is largely absent. Data for (a): 3736(w), 3086(w), 3043(w), 2016(w), 1950(w), 1909(w), 1880(w), 1795(w), 1693(w), 1603(s), 1531(i, s), 1485(sh), 1456(sh), 1404(i, s), 1282(sh), 1165(s), 1088(s), 997(s), 957(sh), 943(s), 933(sh), 858(s), 839(s), 808(sh), 748(i, s), 687(w), 663(s), 594(s), 548(sh), 499(s), 453(s), 440(sh), 401(s). s = sharp, i = intense, w = weak, sh = shoulder.



Figure S2. (a) Thermogravimetric analysis for a crystalline sample of the title compound as prepared, showing loss of 0.65% by mass (or ~0.08 H₂O) to 257 °C, followed by decomposition of the sample at 300 °C. (b) A differential plot shows the time profile of the loss of mass to 257 °C more clearly. (c) Thermogravimetric analysis for a crystalline sample of the title compound following further sample preparation (heating for 2 hours at 240 °C). No significant loss of mass occurs until decomposition of the sample at 292 °C. (d) A differential plot shows the time profile for the loss of mass, with no significant loss occurring to 292 °C.



Figure S3. Temperature dependence of χ^{-1} for a ground sample of the title compound (a) with no heating and (b) after heating crystals to 240 °C for 2 h prior to grinding them for SQUID measurement. Values for Θ_{CW} were determined by fitting data in the temperature range $150 \le T \le 350$ K. Lines indicate best fit obtained.



Figure S4. Temperature dependence of χT for a ground sample of the title compound after heating crystals to 240 °C for 2 h prior to grinding them for SQUID measurement. The data show an upturn at low *T*, indicative of net ferromagnetism.



Figure S5. Field dependence of the magnetization per mole of Cu measured at 2 K. A ground sample of the title compound was measured after heating crystals to 240 °C for 2 h prior to grinding them for SQUID measurement. The line is drawn to guide the eye. The magnetization saturates at 0.91 μ_B / mol Cu, which is consistent with 1 unpaired electron per Cu center.



Figure S6. A portion of the crystal structure of Cu(1,3-bdc), with unit cell. 50% ellipsoids are shown. Symmetry equivalent atoms (shown but not labeled here) are given in Tables S3 and S4. Phenyl-ring carbon and hydrogen atoms are omitted for clarity.



Figure S7. Calculated and observed powder pattern for a crystalline sample of the title compound, after heating for 2 hours at 240 °C.

Identification code	07060	
Empirical formula	$C_8H_4CuO_4$	
Formula weight	227.65	
Temperature	100(2) K	
Wavelength	0.71073 Å	
Crystal system	Hexagonal	
Space group	$P6_3/m$	
Unit cell dimensions	a = 9.1081(2) Å	$\alpha = 90^{\circ}$
	b = 9.1081(2) Å	$\beta = 90^{\circ}$
	c = 15.9432(5) Å	$\gamma = 120^{\circ}$
Volume, Z	1145.41(5) Å ³	6
Density (calculated)	1.980 g/cm^3	
Absorption coefficient	2.831 mm^{-1}	
F(000)	678	
Crystal size	$0.48 \times 0.35 \times 0.15 \text{ mm}$	
θ range for data collection	2.56 to 29.57°	
Limiting indices	$-12 \le h \le 12$	
	$-12 \le k \le 12$	
	$-22 \le l \le 22$	
Reflections collected	25614	
Independent reflections	$1119 (R_{int} = 0.0247)$	
Completeness to $\theta = 29.57^{\circ}$	100.0 %	
Refinement method	Full-matrix least-squares on F ²	
Data / restraints / parameters	1119 / 0 / 64	
Goodness-of-fit on F ²	1.098	
Final R indices $[I > 2\sigma(I)]$	$R_1 = 0.0212$	$wR_2 = 0.0667$
R indices (all data)	$R_1 = 0.0217$	$wR_2 = 0.0674$
Largest diff. peak and hole	$0.528 \text{ e}\text{\AA}^{-3}$	$-0.633 \text{ e}\text{\AA}^{-3}$

Table S1. Crystal data and structure refinement for Cu(1,3-bdc).

	x	У	Ζ.	U(eq)
Cu(1)	5000	0	0	7(1)
O(1)	6360(1)	570(1)	1002(1)	9(1)
O(2)	4262(1)	1614(1)	261(1)	11(1)
C(1)	7828(2)	1867(2)	950(1)	8(1)
C(2)	8839(2)	2404(2)	1741(1)	8(1)
C(3)	10610(2)	3362(2)	1740(1)	10(1)
C(4)	11485(2)	3829(2)	2500	11(1)
C(5)	7962(2)	1915(2)	2500	8(1)

Table S2. Atomic coordinates $[\times 10^4]$ and equivalent isotropic displacement parameters $[\text{\AA}^2 \times 10^3]$ for Cu(1,3-bdc). U(eq) is defined as one third of the trace of the orthogonalized U_{ij} tensor.

Table S3. Bond lengths [Å] for Cu(1,3-bdc).

$Cu(1)-O(1)^{a}$	1.9265(9)	C(1)-C(2)	1.4929(16)	
Cu(1)-O(1)	1.9265(9)	C(2)-C(5)	1.3948(14)	
$Cu(1)-O(2)^{a}$	1.9428(9)	C(2)-C(3)	1.3985(16)	
Cu(1)-O(2)	1.9428(9)	C(3)-C(4)	1.3949(14)	
O(1)-C(1)	1.2685(15)	$C(4)-C(3)^{d}$	1.3949(14)	
$O(2)-C(1)^{b}$	1.2680(14)	$C(5)-C(2)^{d}$	1.3948(14)	
C(1)-O(2) ^c	1.2680(14)			
Symmetry transformations used to generate equivalent atoms:				

Symmetry transion	mations used to generate eq		
$a^{a} - x + 1, -y, -z$	b - x + y + 1, -x + 1, z	$^{c} -y + 1, x - y, z$	$d x, y, -z + \frac{1}{2}$

$O(1)^{a}$ -Cu(1)-O(1)	180.00(7)	$O(2)^{c}-C(1)-C(2)$	122.61(11)
$O(1)^{a}$ -Cu(1)-O(2) ^a	92.14(4)	O(1)-C(1)-C(2)	116.13(10)
$O(1)-Cu(1)-O(2)^{a}$	87.86(4)	C(5)-C(2)-C(3)	119.90(12)
$O(1)^{a}$ -Cu(1)-O(2)	87.86(4)	C(5)-C(2)-C(1)	117.85(11)
O(1)-Cu(1)-O(2)	92.14(4)	C(3)-C(2)-C(1)	122.25(11)
$O(2)^{a}$ -Cu(1)-O(2)	180.00(7)	C(4)-C(3)-C(2)	119.60(12)
C(1)-O(1)-Cu(1)	115.67(8)	$C(3)-C(4)-C(3)^d$	120.64(16)
$C(1)^{b}$ -O(2)-Cu(1)	132.42(8)	$C(2)^{d}$ - $C(5)$ - $C(2)$	120.34(15)
$O(2)^{c}-C(1)-O(1)$	121.26(11)		

 Table S4. Bond angles [°] for Cu(1,3-bdc).

Symmetry transfor	mations used to generate equ	uivalent atoms:	
a - x + 1, -y, -z	b - x + y + 1, -x + 1, z	$^{c} -y + 1, x - y, z$	$d x, y, -z + \frac{1}{2}$

	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
Cu(1)	8(1)	7(1)	4(1)	-1(1)	-1(1)	4(1)
O(1)	9(1)	10(1)	7(1)	-1(1)	-1(1)	3(1)
O(2)	16(1)	13(1)	6(1)	-1(1)	0(1)	9(1)
C(1)	10(1)	9(1)	6(1)	-1(1)	0(1)	5(1)
C(2)	9(1)	9(1)	6(1)	-1(1)	-1(1)	4(1)
C(3)	9(1)	12(1)	8(1)	1(1)	1(1)	5(1)
C(4)	8(1)	13(1)	11(1)	0	0	5(1)
C(5)	8(1)	8(1)	7(1)	0	0	3(1)

Table S5. Anisotropic displacement parameters $[Å^2 \times 10^3]$ for Cu(1,3-bdc). The anisotropic displacement factor exponent takes the form: $-2\pi^2[(ha^*)^2U_{11} + ... + 2hka^*b^*U_{12}]$.

Table S6. Hydrogen coordinates $[\times 10^4]$ and isotropic displacement parameters $[\text{Å}^2 \times 10^3]$ for Cu(1,3-bdc).

	x	у	Z	U(eq)
H(3)	11214	3693	1224	12
H(4)	12689	4472	2500	13
H(5)	6760	1246	2500	10



Figure S8. A portion of the crystal structure of $Cu(1,3-bdc)\cdot 0.11 H_2O$, with unit cell. 50% ellipsoids are shown. O(3) is the oxygen atom of a solvated water molecule statistically distributed over 1/9 of sites. Symmetry equivalent atoms (shown but not labeled here) are given in Tables S9 and S10. Phenyl-ring carbon atoms and all hydrogen atoms are omitted for clarity. The distances between O(3) and other atoms in the structure are as follows: O(3)–Cu(1), 4.5528(2) Å. O(3)–O(1A), 3.9405(8) Å. O(3)–C(1A), 3.5303(11) Å. O(3)–O(2B), 3.4166(8) Å.

Identification code	05251	
Empirical formula	$C_8H_4CuO_{4.11}$	
Formula weight	229.44	
Temperature	100(2) K	
Wavelength	0.71073 Å	
Crystal system	Hexagonal	
Space group	P6 ₃ /m	
Unit cell dimensions	a = 9.1056(3) Å	$\alpha = 90^{\circ}$
	b = 9.1056(3) Å	$\beta = 90^{\circ}$
	c = 15.9550(12) Å	$\gamma = 120^{\circ}$
Volume, Z	1145.63(10) Å ³	6
Density (calculated)	1.995 g/cm ³	
Absorption coefficient	2.832 mm^{-1}	
F(000)	683	
Crystal size	$0.15 \times 0.10 \times 0.10 \text{ mm}$	
θ range for data collection	2.55 to 29.55°	
Limiting indices	$-12 \le h \le 12$	
	$-12 \le k \le 12$	
	$-22 \le l \le 22$	
Reflections collected	22913	
Independent reflections	1118 ($R_{int} = 0.0311$)	
Completeness to $\theta = 29.55^{\circ}$	100.0 %	
Refinement method	Full-matrix least-squares on F ²	
Data / restraints / parameters	1118 / 0 / 66	
Goodness-of-fit on F ²	1.101	
Final R indices $[I > 2\sigma(I)]$	$R_1 = 0.0217$	$wR_2 = 0.0621$
R indices (all data)	$R_1 = 0.0235$	$wR_2 = 0.0637$
Largest diff. peak and hole	$0.578 \text{ e}\text{\AA}^{-3}$	-1.128 eÅ ⁻³

Table S7. Crystal data and structure refinement for $Cu(1,3-bdc) \cdot 0.11 H_2O$.

	x	у	Z.	U(eq)
Cu(1)	5000	5000	0	8(1)
O(1)	6361(1)	5790(1)	1002(1)	11(1)
O(2)	4259(1)	2646(1)	261(1)	13(1)
C(1)	7826(1)	5961(1)	950(1)	10(1)
C(2)	8838(1)	6435(1)	1741(1)	10(1)
C(3)	10612(1)	7250(1)	1741(1)	12(1)
C(4)	11484(2)	7656(2)	2500	12(1)
C(5)	7963(2)	6049(2)	2500	10(1)
O(3)	0	0	0	61(4)

Table S8. Atomic coordinates $[\times 10^4]$ and equivalent isotropic displacement parameters $[\text{\AA}^2 \times 10^3]$ for Cu(1,3-bdc)·0.11 H₂O. *U(eq)* is defined as one third of the trace of the orthogonalized U_{ij} tensor.

Table S9. Bond lengths [Å] for Cu(1,3-bdc)·0.11 H₂O.

$Cu(1)-O(1)^{a}$	1.9275(8)	C(1)-C(2)	1.4932(15)
Cu(1)-O(1)	1.9275(8)	C(2)-C(5)	1.3940(13)
Cu(1)-O(2)	1.9437(8)	C(2)-C(3)	1.4006(16)
$Cu(1)-O(2)^{a}$	1.9437(8)	C(3)-C(4)	1.3934(14)
O(1)-C(1)	1.2665(14)	$C(4)-C(3)^{d}$	1.3934(14)
$O(2)-C(1)^{b}$	1.2692(14)	$C(5)-C(2)^{d}$	1.3940(13)
C(1)-O(2) ^c	1.2692(14)		

Symmetry transformations used to generate equivalent atoms: a -x + 1, -y + 1, -z b -y + 1, x - y, z c -x + y + 1, -x + 1, z $d x, y, -z + \frac{1}{2}$

$O(1)^{a}$ -Cu(1)-O(1)	180.00(3)	O(1)-C(1)-C(2)	116.12(10)
$O(1)^{a}$ -Cu(1)-O(2)	87.84(3)	$O(2)^{c}-C(1)-C(2)$	122.63(10)
O(1)-Cu(1)-O(2)	92.16(3)	C(5)-C(2)-C(3)	119.78(11)
$O(1)^{a}$ -Cu(1)-O(2)^{a}	92.16(3)	C(5)-C(2)-C(1)	117.97(10)
$O(1)-Cu(1)-O(2)^{a}$	87.84(3)	C(3)-C(2)-C(1)	122.25(10)
$O(2)-Cu(1)-O(2)^{a}$	180.00(5)	C(4)-C(3)-C(2)	119.55(11)
C(1)-O(1)-Cu(1)	115.69(7)	$C(3)-C(4)-C(3)^d$	120.81(15)
$C(1)^{b}-O(2)-Cu(1)$	132.36(8)	$C(2)^{d}$ - $C(5)$ - $C(2)$	120.53(14)
$O(1)-C(1)-O(2)^{c}$	121.25(10)		

Table S10. Bond angles [°] for $Cu(1,3-bdc) \cdot 0.11 H_2O$.

Symmetry t	ransformation	ns used to gen	nerate equival	lent atoms:	
<i>a</i> 1	1 1 -	b 1	с		1 _ d

 $a^{a} - x + 1, -y + 1, -z$ $b^{b} - y + 1, x - y, z$ $c^{-} - x + y + 1, -x + 1, z$ $d^{a} x, y, -z + \frac{1}{2}$

	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
Cu(1)	10(1)	10(1)	6(1)	0(1)	-1(1)	5(1)
O(1)	12(1)	14(1)	9(1)	0(1)	-1(1)	7(1)
O(2)	19(1)	12(1)	8(1)	0(1)	-1(1)	8(1)
C(1)	11(1)	9(1)	9(1)	1(1)	0(1)	5(1)
C(2)	12(1)	11(1)	8(1)	0(1)	-1(1)	6(1)
C(3)	11(1)	12(1)	11(1)	0(1)	2(1)	5(1)
C(4)	9(1)	13(1)	13(1)	0	0	3(1)
C(5)	10(1)	10(1)	10(1)	0	0	5(1)
O(3)	54(6)	54(6)	75(12)	0	0	27(3)

Table S11. Anisotropic displacement parameters $[\text{\AA}^2 \times 10^3]$ for Cu(1,3-bdc)·0.11 H₂O. The anisotropic displacement factor exponent takes the form: $-2\pi^2[(ha^*)^2U_{11} + ... + 2hka^*b^*U_{12}]$.

Table S12. Hydrogen coordinates [× 10^4] and isotropic displacement parameters [Å² × 10^3] for Cu(1,3-bdc)·0.11 H₂O.

	x	у	Z	U(eq)
H(3)	11216	7523	1226	14
H(4)	12687	8216	2500	15
H(5)	6761	5519	2500	12



Figure S9. A portion of the crystal structure of $C_{32}H_{24}Cu_6O_{26}$. 50% ellipsoids are shown. Symmetry equivalent atoms (shown but not labeled here) are given in Tables S15 and S16.



Figure S10. Packing diagram for $C_{32}H_{24}Cu_6O_{26}$, as viewed down the crystallographic *a* axis. 50% thermal ellipsoids are shown. Cu atoms are shown in teal, O atoms in red, C atoms in gray, and H atoms in white.



Figure S11. Packing diagram for $C_{32}H_{24}Cu_6O_{26}$, as viewed down the crystallographic *c* axis. 50% thermal ellipsoids are shown. Cu atoms are shown in teal, O atoms in red, C atoms in gray, and H atoms in white.

0(210	
06218	
$C_{32}H_{24}Cu_6O_{26}$	
1205.75	
100(2) K	
0.71073 Å	
Triclinic	
<i>P</i> –1	
a = 7.0738(6) Å	$\alpha = 104.0140(10)^{\circ}$
b = 10.8611(8) Å	$\beta = 99.4120(10)^{\circ}$
c = 12.0790(9) Å	$\gamma = 98.2080(10)^{\circ}$
872.01(12) Å ³	1
2.296 g/cm ³	
3.705mm^{-1}	
598	
$0.20 \times 0.10 \times 0.02 \text{ mm}$	
1.97 to 29.13°	
$-9 \le h \le 9$	
$-14 \le k \le 14$	
$-16 \le l \le 16$	
18271	
$4692 (R_{int} = 0.0450)$	
99.7 %	
Full-matrix least-squares on F ²	
4692 / 9 / 307	
1.021	
$R_1 = 0.0333$	$wR_2 = 0.0699$
$R_1 = 0.0512$	$wR_2 = 0.0768$
$0.689 \text{ e}\text{\AA}^{-3}$	$-0.540 \text{ e}\text{\AA}^{-3}$
	C ₃₂ H ₂₄ Cu ₆ O ₂₆ 1205.75 100(2) K 0.71073 Å Triclinic P-1 a = 7.0738(6) Å b = 10.8611(8) Å c = 12.0790(9) Å 872.01(12) Å ³ 2.296 g/cm ³ 3.705mm ⁻¹ 598 0.20 × 0.10 × 0.02 mm 1.97 to 29.13° $-9 \le h \le 9$ $-14 \le k \le 14$ $-16 \le l \le 16$ 18271 4692 (R _{int} = 0.0450) 99.7 % Full-matrix least-squares on F ² 4692 / 9 / 307 1.021 $R_1 = 0.0333$ $R_1 = 0.0512$ 0.689 eÅ ⁻³

Table S13. Crystal data and structure refinement for $C_{32}H_{24}Cu_6O_{26}$.

	X	У	Z.	U(eq)
Cu(1)	6472(1)	3610(1)	-80(1)	10(1)
Cu(2)	8589(1)	6226(1)	50(1)	10(1)
Cu(3)	3842(1)	996(1)	2305(1)	13(1)
O(1)	8901(3)	7858(2)	1121(2)	12(1)
O(2)	8597(3)	9266(2)	2712(2)	13(1)
O(3)	4939(3)	3019(2)	907(2)	14(1)
O(4)	4416(3)	2887(2)	2629(2)	12(1)
O(5)	10652(3)	6967(2)	-2650(2)	12(1)
O(6)	10140(3)	6797(2)	-936(2)	13(1)
O(7)	6048(3)	1993(2)	-1191(2)	13(1)
O(8)	6403(3)	605(2)	-2776(2)	15(1)
O(9)	2253(4)	379(2)	463(2)	37(1)
O(10)	1550(3)	1210(2)	2982(2)	15(1)
O(11)	6415(3)	882(2)	1946(3)	32(1)
O(12)	8049(3)	4469(2)	-930(2)	10(1)
O(13)	6970(3)	5370(2)	880(2)	11(1)
C(11)	7636(4)	7083(3)	2651(2)	10(1)
C(12)	6904(4)	5777(3)	2024(2)	10(1)
C(13)	6079(4)	4898(2)	2589(2)	9(1)
C(14)	6121(4)	5343(3)	3788(2)	11(1)
C(15)	6894(4)	6599(3)	4412(2)	13(1)
C(16)	7610(4)	7471(3)	3842(2)	12(1)
C(21)	8860(4)	4960(3)	-2649(2)	10(1)
C(22)	8048(4)	4083(3)	-2084(2)	9(1)
C(23)	7295(4)	2780(3)	-2715(2)	11(1)
C(24)	7297(4)	2400(3)	-3909(2)	12(1)
C(25)	7996(4)	3275(3)	-4483(2)	13(1)
C(26)	8790(4)	4537(3)	-3849(2)	12(1)
C(110)	8423(4)	8119(3)	2124(2)	11(1)
C(130)	5117(4)	3530(3)	2000(2)	11(1)
C(210)	9923(4)	6319(3)	-2047(2)	11(1)

Table S14. Atomic coordinates $[\times 10^4]$ and equivalent isotropic displacement parameters $[\text{\AA}^2 \times 10^3]$ for C₃₂H₂₄Cu₆O₂₆. U(eq) is defined as one third of the trace of the orthogonalized U_{ij} tensor.

C(230) 6541(4) 1750(3) -2171(2) 1	11(1)
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Cu(1)-O(7)	1.8846(19)	O(8)-C(230)	1.262(3)
Cu(1)-O(3)	1.8980(19)	$O(8)-Cu(3)^{a}$	1.9500(19)
Cu(1)-O(12)	1.9267(18)	O(13)-C(12)	1.355(3)
Cu(1)-O(13)	1.9295(18)	O(12)-C(22)	1.354(3)
Cu(1)-Cu(2)	2.9714(5)	C(11)-C(16)	1.401(4)
Cu(2)-O(1)	1.8840(19)	C(11)-C(12)	1.412(4)
Cu(2)-O(6)	1.8969(19)	C(11)-C(110)	1.505(4)
Cu(2)-O(13)	1.9244(18)	C(12)-C(13)	1.416(3)
Cu(2)-O(12)	1.9334(18)	C(13)-C(14)	1.405(4)
$Cu(3)-O(8)^{a}$	1.9500(19)	C(13)-C(130)	1.490(4)
Cu(3)-O(10)	1.950(2)	C(14)-C(15)	1.371(4)
Cu(3)-O(11)	1.952(2)	C(15)-C(16)	1.386(4)
Cu(3)-O(4)	1.9647(18)	C(21)-C(26)	1.400(4)
Cu(3)-O(9)	2.216(3)	C(21)-C(22)	1.411(4)
O(1)-C(110)	1.288(3)	C(21)-C(210)	1.503(4)
O(2)-C(110)	1.252(3)	C(22)-C(23)	1.413(4)
O(3)-C(130)	1.280(3)	C(23)-C(24)	1.401(4)
O(4)-C(130)	1.262(3)	C(23)-C(230)	1.505(4)
O(5)-C(210)	1.248(3)	C(24)-C(25)	1.388(4)
O(6)-C(210)	1.291(3)	C(25)-C(26)	1.383(4)
O(7)-C(230)	1.266(3)		

Table S15. Bond lengths [Å] for $C_{32}H_{24}Cu_6O_{26}$.

Symmetry transformations used to generate equivalent atoms: $a^{a}-x+1, -y, -z$

O(7)-Cu(1)-O(3)	95.27(8)	C(130)-O(3)-Cu(1)	126.16(17)
O(7)-Cu(1)-O(12)	93.37(8)	C(130)-O(4)-Cu(3)	126.70(17)
O(3)-Cu(1)-O(12)	171.28(8)	C(210)-O(6)-Cu(2)	126.77(17)
O(7)-Cu(1)-O(13)	171.89(8)	C(230)-O(7)-Cu(1)	127.52(17)
O(3)-Cu(1)-O(13)	92.07(8)	$C(230)-O(8)-Cu(3)^{a}$	129.34(18)
O(12)-Cu(1)-O(13)	79.23(8)	C(12)-O(13)-Cu(2)	126.82(16)
O(7)-Cu(1)-Cu(2)	133.08(6)	C(12)-O(13)-Cu(1)	127.37(16)
O(3)-Cu(1)-Cu(2)	131.56(6)	Cu(2)-O(13)-Cu(1)	100.89(8)
O(12)-Cu(1)-Cu(2)	39.75(5)	C(22)-O(12)-Cu(1)	127.00(16)
O(13)-Cu(1)-Cu(2)	39.49(5)	C(22)-O(12)-Cu(2)	127.04(16)
O(1)-Cu(2)-O(6)	95.90(8)	Cu(1)-O(12)-Cu(2)	100.67(8)
O(1)-Cu(2)-O(13)	93.45(8)	C(16)-C(11)-C(12)	119.0(2)
O(6)-Cu(2)-O(13)	170.65(8)	C(16)-C(11)-C(110)	116.5(2)
O(1)-Cu(2)-O(12)	172.63(8)	C(12)-C(11)-C(110)	124.5(2)
O(6)-Cu(2)-O(12)	91.46(8)	O(13)-C(12)-C(11)	120.1(2)
O(13)-Cu(2)-O(12)	79.19(8)	O(13)-C(12)-C(13)	120.1(2)
O(1)-Cu(2)-Cu(1)	133.06(6)	C(11)-C(12)-C(13)	119.8(2)
O(6)-Cu(2)-Cu(1)	131.03(6)	C(14)-C(13)-C(12)	118.3(2)
O(13)-Cu(2)-Cu(1)	39.62(5)	C(14)-C(13)-C(130)	116.9(2)
O(12)-Cu(2)-Cu(1)	39.58(5)	C(12)-C(13)-C(130)	124.8(2)
$O(8)^{a}$ -Cu(3)-O(10)	83.46(8)	C(15)-C(14)-C(13)	122.1(2)
$O(8)^{a}$ -Cu(3)-O(11)	93.71(9)	C(14)-C(15)-C(16)	119.3(3)
O(10)-Cu(3)-O(11)	168.69(11)	C(15)-C(16)-C(11)	121.3(2)
$O(8)^{a}$ -Cu(3)-O(4)	152.98(8)	C(26)-C(21)-C(22)	119.0(2)
O(10)-Cu(3)-O(4)	86.66(8)	C(26)-C(21)-C(210)	116.2(2)
O(11)-Cu(3)-O(4)	91.08(9)	C(22)-C(21)-C(210)	124.8(2)
$O(8)^{a}$ -Cu(3)-O(9)	103.46(9)	O(12)-C(22)-C(21)	120.5(2)
O(10)-Cu(3)-O(9)	96.01(10)	O(12)-C(22)-C(23)	119.7(2)
O(11)-Cu(3)-O(9)	95.29(12)	C(21)-C(22)-C(23)	119.8(2)
O(4)-Cu(3)-O(9)	102.55(9)	C(24)-C(23)-C(22)	118.9(2)
C(110)-O(1)-Cu(2)	127.92(17)	C(24)-C(23)-C(230)	117.4(2)

Table S16. Bond angles [°] for $C_{32}H_{24}Cu_6O_{26}$.

C(22)-C(23)-C(230)	123.7(2)	O(4)-C(130)-C(13)	116.6(2)
C(25)-C(24)-C(23)	121.5(3)	O(3)-C(130)-C(13)	122.9(2)
C(26)-C(25)-C(24)	119.0(3)	O(5)-C(210)-O(6)	120.6(2)
C(25)-C(26)-C(21)	121.7(2)	O(5)-C(210)-C(21)	117.8(2)
O(2)-C(110)-O(1)	120.9(2)	O(6)-C(210)-C(21)	121.6(2)
O(2)-C(110)-C(11)	116.7(2)	O(8)-C(230)-O(7)	121.6(2)
O(1)-C(110)-C(11)	122.5(2)	O(8)-C(230)-C(23)	114.9(2)
O(4)-C(130)-O(3)	120.5(2)	O(7)-C(230)-C(23)	123.5(2)

Table S17. Bond angles [°] for $C_{32}H_{24}Cu_6O_{26}$ (continued).

Symmetry transformations used to generate equivalent atoms: a - x + 1, -y, -z

	<i>U</i> ₁₁	U_{22}	U_{33}	U_{23}	<i>U</i> ₁₃	U_{12}
Cu(1)	14(1)	7(1)	11(1)	3(1)	4(1)	-1(1)
Cu(2)	13(1)	7(1)	12(1)	4(1)	4(1)	-1(1)
Cu(3)	15(1)	6(1)	20(1)	5(1)	7(1)	1(1)
O(1)	17(1)	9(1)	14(1)	5(1)	6(1)	2(1)
O(2)	16(1)	7(1)	17(1)	3(1)	6(1)	0(1)
O(3)	17(1)	10(1)	13(1)	3(1)	4(1)	-2(1)
O(4)	16(1)	6(1)	15(1)	4(1)	6(1)	0(1)
O(5)	15(1)	9(1)	15(1)	6(1)	6(1)	0(1)
O(6)	16(1)	9(1)	14(1)	4(1)	4(1)	-2(1)
O(7)	19(1)	8(1)	12(1)	5(1)	5(1)	0(1)
O(8)	20(1)	5(1)	19(1)	3(1)	6(1)	0(1)
O(9)	66(2)	20(1)	21(1)	7(1)	3(1)	-4(1)
O(10)	15(1)	6(1)	28(1)	7(1)	8(1)	2(1)
O(11)	22(1)	14(1)	68(2)	19(1)	19(1)	6(1)
O(12)	15(1)	6(1)	10(1)	4(1)	4(1)	-2(1)
O(13)	14(1)	8(1)	12(1)	4(1)	5(1)	0(1)
C(11)	9(1)	9(1)	13(1)	5(1)	2(1)	2(1)
C(12)	9(1)	10(1)	12(1)	6(1)	3(1)	4(1)
C(13)	7(1)	8(1)	13(1)	5(1)	3(1)	2(1)
C(14)	11(1)	10(1)	14(1)	6(1)	5(1)	3(1)
C(15)	16(1)	13(1)	11(1)	4(1)	4(1)	3(1)
C(16)	14(1)	8(1)	14(1)	1(1)	4(1)	2(1)
C(21)	10(1)	9(1)	14(1)	6(1)	3(1)	4(1)
C(22)	8(1)	10(1)	11(1)	4(1)	2(1)	4(1)
C(23)	11(1)	10(1)	13(1)	4(1)	2(1)	2(1)
C(24)	12(1)	10(1)	13(1)	4(1)	1(1)	2(1)
C(25)	14(1)	15(1)	13(1)	5(1)	3(1)	4(1)
C(26)	11(1)	13(1)	14(1)	8(1)	3(1)	4(1)
C(110)	9(1)	9(1)	17(1)	5(1)	2(1)	2(1)
C(130)	11(1)	8(1)	15(1)	4(1)	5(1)	3(1)

Table S18. Anisotropic displacement parameters $[Å^2 \times 10^3]$ for $C_{32}H_{24}Cu_6O_{26}$. The anisotropic displacement factor exponent takes the form: $-2\pi^2[(ha^*)^2U_{11} + ... + 2hka^*b^*U_{12}]$.

Cu(210)	9(1)	11(1)	16(1)	7(1)	4(1)	5(1)
Cu(230)	9(1)	9(1)	15(1)	4(1)	-1(1)	2(1)

Table S19. Hydrogen coordinates $[\times 10^4]$ and isotropic displacement parameters $[\text{Å}^2 \times 10^3]$ for $C_{32}H_{24}Cu_6O_{26}$.

	X	у	Z.	U(eq)
H(9A)	1100(30)	200(40)	470(40)	56
H(9B)	2350(60)	880(30)	60(30)	56
H(10A)	970(40)	1750(20)	2780(30)	23
H(10B)	660(40)	570(20)	2860(30)	23
H(11A)	7210(50)	1560(20)	2120(40)	47
H(11B)	7070(50)	340(30)	2100(40)	47
H(14)	5596	4756	4178	13
H(15)	6939	6868	5226	15
H(16)	8092	8349	4266	14
H(24)	6809	1523	-4336	14
H(25)	7930	3012	-5301	16
H(26)	9300	5132	-4236	14

D–H…A	<i>d</i> (D–H)	$d(\mathbf{H}\cdots\mathbf{A})$	$d(\mathbf{D}\cdots\mathbf{A})$	<(DHA)
O(9)-H(9A)····O(9) ^b	0.809(18)	2.38(3)	3.119(6)	153(4)
O(9)-H(9B)···O(1) ^c	0.820(18)	2.36(3)	3.115(3)	154(4)
O(9)-H(9B)····O(3)	0.820(18)	2.61(4)	3.080(3)	118(3)
O(10)-H(10A)····O(5) ^c	0.821(17)	1.946(19)	2.752(3)	166(3)
O(10)-H(10B)···O(2) ^d	0.837(17)	1.833(18)	2.667(3)	174(3)
O(11)-H(11A)···O(5) ^e	0.823(18)	1.93(2)	2.751(3)	171(4)
O(11)-H(11B)···O(2) ^f	0.836(18)	1.89(2)	2.723(3)	171(4)

Table S20. Hydrogen bond lengths [Å] and angles [°] for $C_{32}H_{24}Cu_6O_{26}$.

Symmetry transformations used to generate equivalent atoms: a -x + 1, -y, -z b -x, -y, -z c -x + 1, -y + 1, -z d x - 1, y - 1, z e -x + 2, -y + 1, -z f x, y - 1, z