Supplemental Information

Direct Determination of the Gibbs' Energy of Formation of Peroxynitrous Acid

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Figure 5sup: The decay of 0.21 mM ONOOH in the presence of 0.5 M H^+ , residual H_2O_2 (ca. 14 μM) and 0.75 mM added HNO_2 ($\lambda = 290 \text{ nm}$, $I = 0.75\text{M}$, $T = 25^\circ\text{C}$). The simulated solid line was obtained using reactions 1- 3 and 9 (the rate constants were taken from Table 1), $\epsilon_{290}(\text{ONOOH}) = 185 \text{ M}^{-1}\text{cm}^{-1}$, $\epsilon_{290}(\text{NO}_3^-) = 6 \text{ M}^{-1}\text{cm}^{-1}$ and $\epsilon_{290}(\text{HNO}_2) = 2 \text{ M}^{-1}\text{cm}^{-1}$.