

Supporting Information for the Manuscript:

The Catalytic Role of Nickel in the Decarbonylative Addition of Phthalimides to Alkynes

Albert Poater,^{†,*} Sai Vikrama Chaitanya Vummaleti,[‡] and Luigi Cavallo^{‡,†}

[†] *Institut de Química Computacional i Catàlisi and Departament de Química, Universitat de Girona, Campus de Montilivi, E-17071 Girona, Spain.*

[‡] *KAUST Catalyst Center, 4700 King Abdullah University of Science and Technology, Thuwal 23955-6900, Kingdom of Saudi Arabia.*

Computational Details

All the density functional theory (DFT) calculations were performed using the Gaussian09 package.ⁱ The BP86 GGA functional of Becke and Perdewⁱⁱ was used with the standard split-valence basis set with a polarization function of Ahlrichs and coworkers for H, C, N and P atoms (SVP keyword in Gaussian)ⁱⁱⁱ while the quasi relativistic small-core Stuttgart effective core potential (ECP) was used for Nickel (SDD keyword in Gaussian09), for geometry optimizations.^{iv} The reported energies have been obtained via single point energy calculations with the M06 functional^v with the triple- ζ basis set of Ahlrichs for main group atoms (TZVP keyword in Gaussian09).^{vi} Solvent effects, toluene, were included with the default Gaussian PCM implementation.^{vii} The M06 energy calculations were carried out with the `scf=tight`, and `integral(grid=ultrafinegrid)` keywords. Zero point energies and thermal corrections calculated at the BP86 level were added to the M06 in solvent energies to approximate free energies in solvent. As reference system we used the separated reactants, instead of using the pre-reactant complex approach.

In some cases we discuss the change in the local aromaticity of a given ring. As a structure-based measure, we used the harmonic oscillator model of aromaticity (HOMA) index, defined by Kruszewski and Krygowski as:^{viii}

$$HOMA = 1 - \frac{\alpha}{n} \sum_{i=1}^n (R_{opt} - R_i)^2, \quad (1)$$

where n is the number of bonds considered, and α is an empirical constant (for CC bonds $\alpha = 257.7$ and for CN bonds $\alpha = 93.2$) fixed to give $HOMA = 0$ for a model nonaromatic system, and $HOMA = 1$ for a system with all bonds equal to an optimal value R_{opt} , which is 1.388 Å for CC bonds (1.334 Å for CN bonds), assumed to be achieved for fully aromatic systems.^{viiiia} R_i stands for a running bond length. HOMA calculations have proved to be useful in similar studies. **¡Error! Marcador no definido.**^{d,ix}

To evaluate the strength of bonds, a Mayer Bond Order (MBO) analysis has been envisaged. This index is derived from the definition of bond order due to Wiberg^x based on the P matrix.

$$P_{st} = \sum_i^{occupied} n_i c_{rs} c_{rt} \quad (2) \quad (3) \square$$

The Wiberg bond order, applicable to NDO-type theories where the atomic orbital basis forms an orthonormal set, uses the square of the off-diagonal elements of P,

$$B_{AB}^{Wiberg} = \sum_s^{onA} \sum_t^{onB} P_{st}^2 \quad (4) \square \quad (3)$$

Mayer^{xi} has suggested a method for calculating bond orders from the P matrix,

$$B_{AB}^{Mayer} = \sum_s^{onA} \sum_t^{onB} (PS)_{st} (PS)_{ts} \quad (4)$$

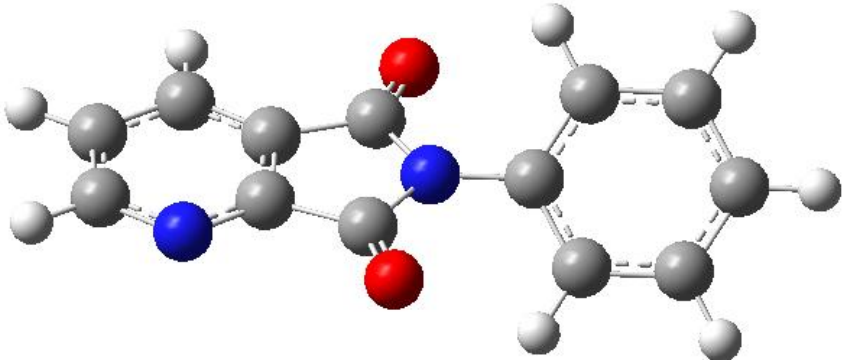
The Mayer definition can be seen as an extension of the Wiberg index. This leads to the classical integer values for homonuclear diatomics when minimal or small basis sets are used. Non-integer values are found for larger basis sets and in more complicated molecules and these reflect the polarized character of the bonds as well as delocalization and multicenter effects. Mayer bond orders are a valuable tool in the analysis of the bonding in main group^{xii} and transition metal^{ix,xiii} systems.

It is worth mentioning here that for all the species presented in the manuscript, we have computed the closed-shell singlet and the triplet electronic states. Since the triplet state is always much higher in energy, in the following sections, we confine our discussion only to the singlet ground state.

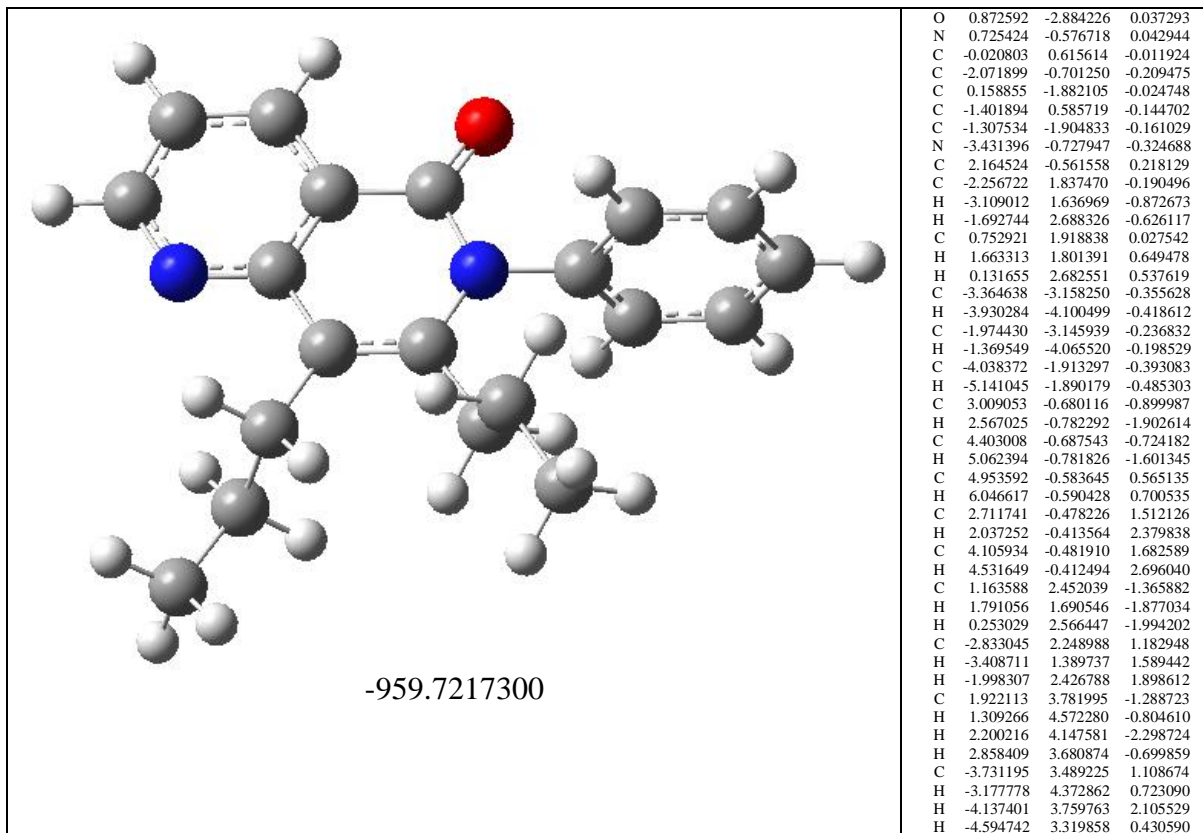
Figure S1. Potential parallel stationary points the reaction pathway. Free energies in solution (toluene as the solvent) relative to species **3** are given. The imaginary frequency characterizing the transition state structures is given in brackets. Selected distances are given in Å.

Table S1. xyz coordinate data sets and absolute energies in a.u. for DFT optimized complexes.

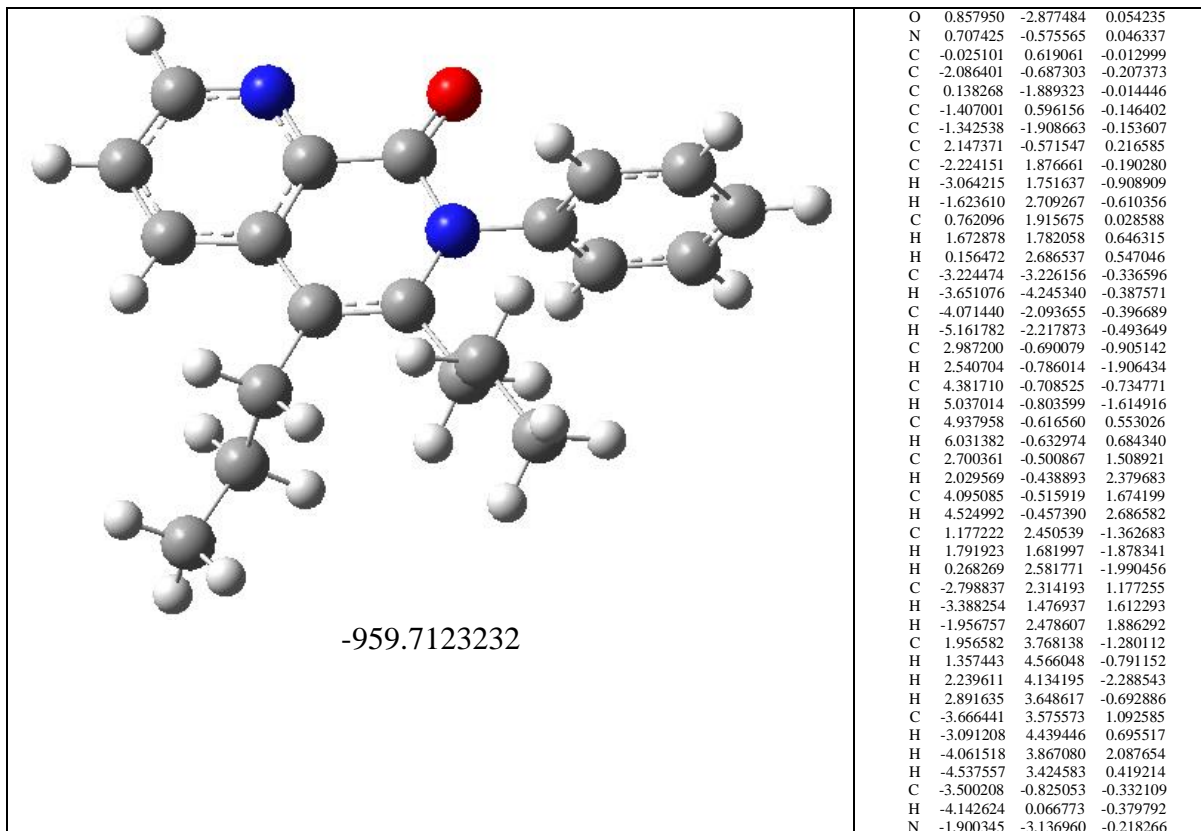
1

 <p style="text-align: center;">-759.8956565</p>	O	0.074530	2.307176	0.360924
	N	-0.315810	0.005130	0.002412
	C	1.918106	-0.678965	-0.107330
	C	0.484147	1.173217	0.183905
	C	1.903209	0.707744	0.111511
	N	3.010536	-1.436254	-0.227868
	C	-1.745029	0.002479	0.001746
	C	4.279840	0.627812	0.098360
	H	5.275037	1.093062	0.171484
	C	3.110872	1.403384	0.221558
	H	3.136684	2.490818	0.393241
	C	4.177032	-0.764603	-0.122596
	H	5.095608	-1.370818	-0.219383
	C	-2.453521	1.064509	-0.601335
	H	-1.906551	1.897631	-1.061649
	C	-3.856998	1.056036	-0.593332
	H	-4.403114	1.890466	-1.060987
	C	-4.562067	-0.006549	-0.002167
	H	-5.663377	-0.010177	-0.003784
	C	-2.448190	-1.064240	0.602688
H	-1.897394	-1.893963	1.064512	
C	-3.851759	-1.064450	0.590926	
H	-4.393668	-1.902521	1.056906	
C	0.489947	-1.167750	-0.181692	
O	0.077990	-2.295722	-0.357872	

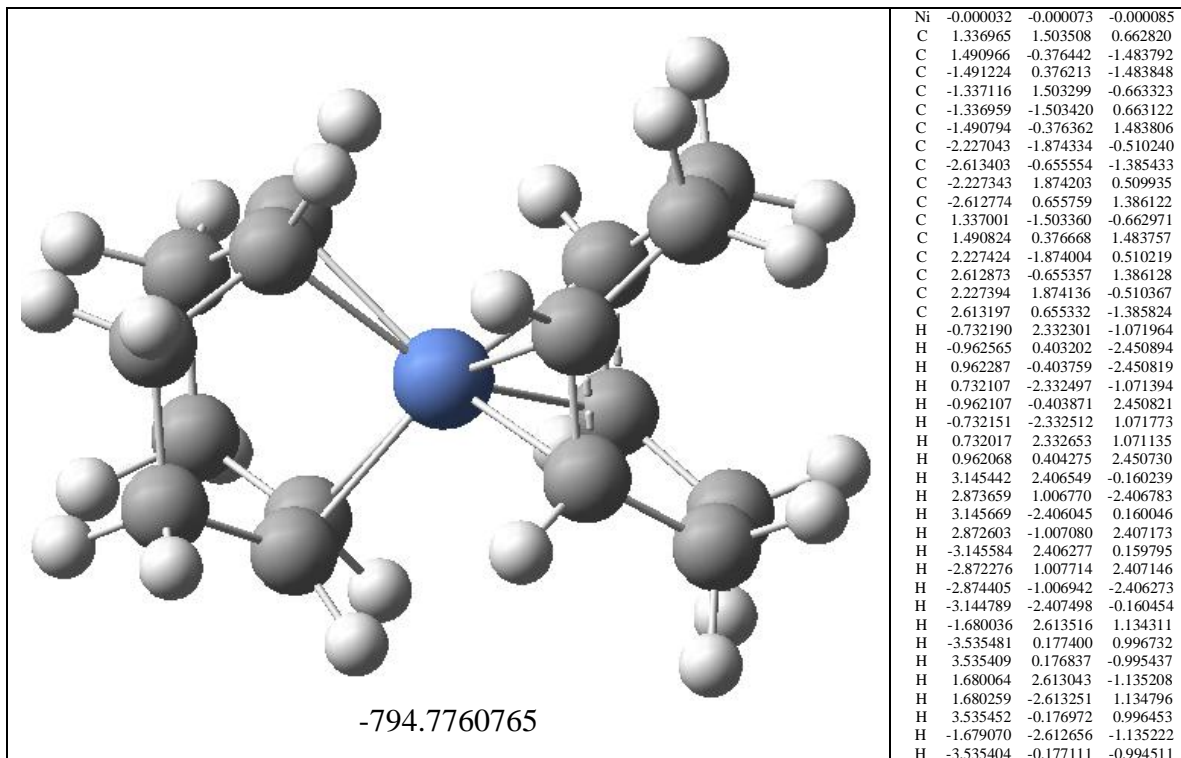
2



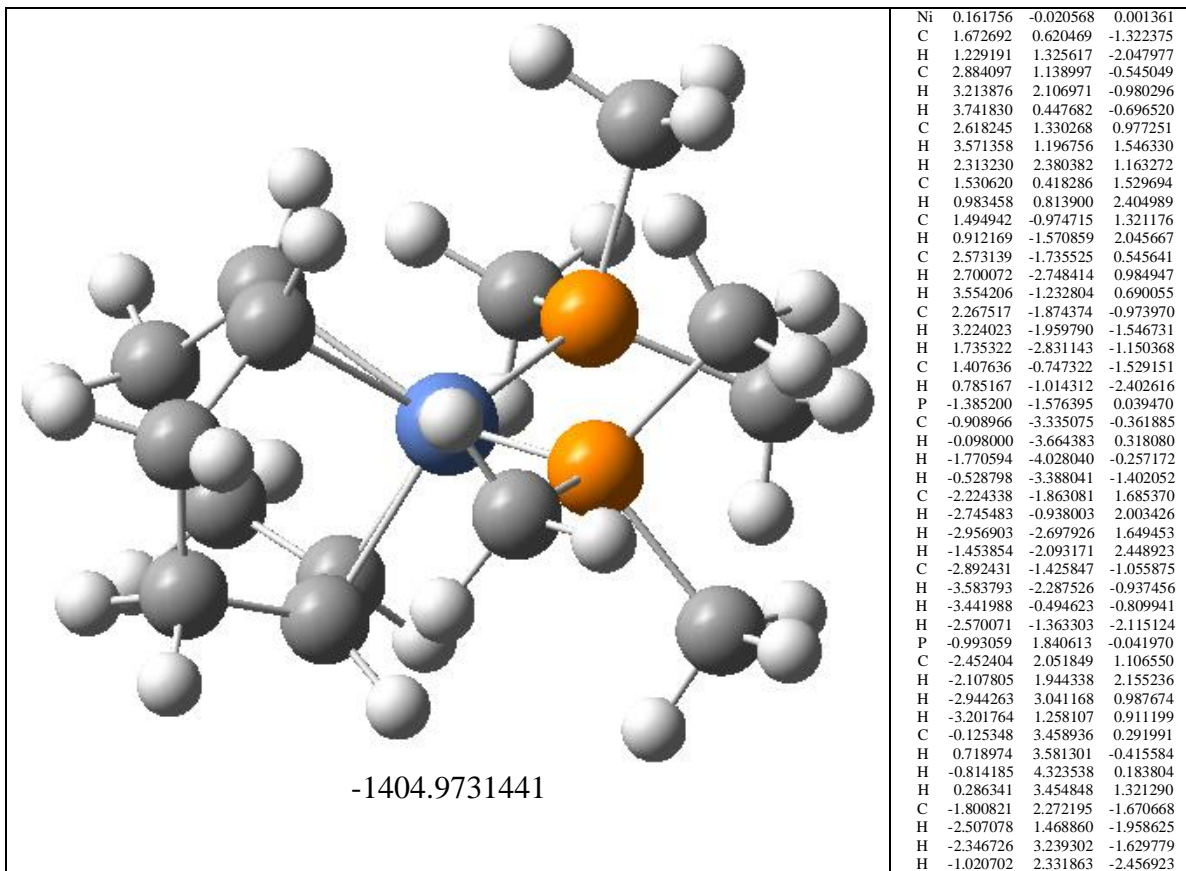
2'



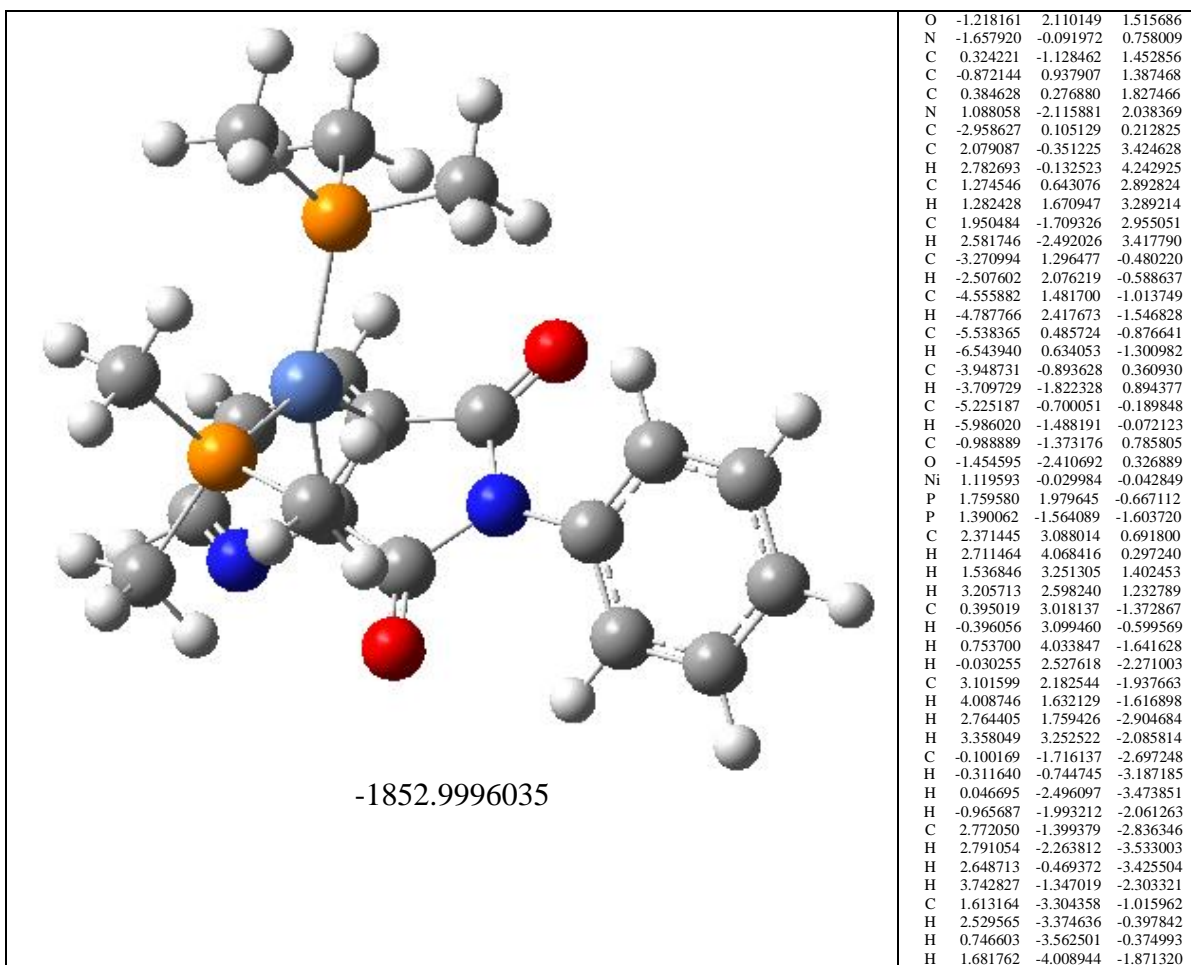
3



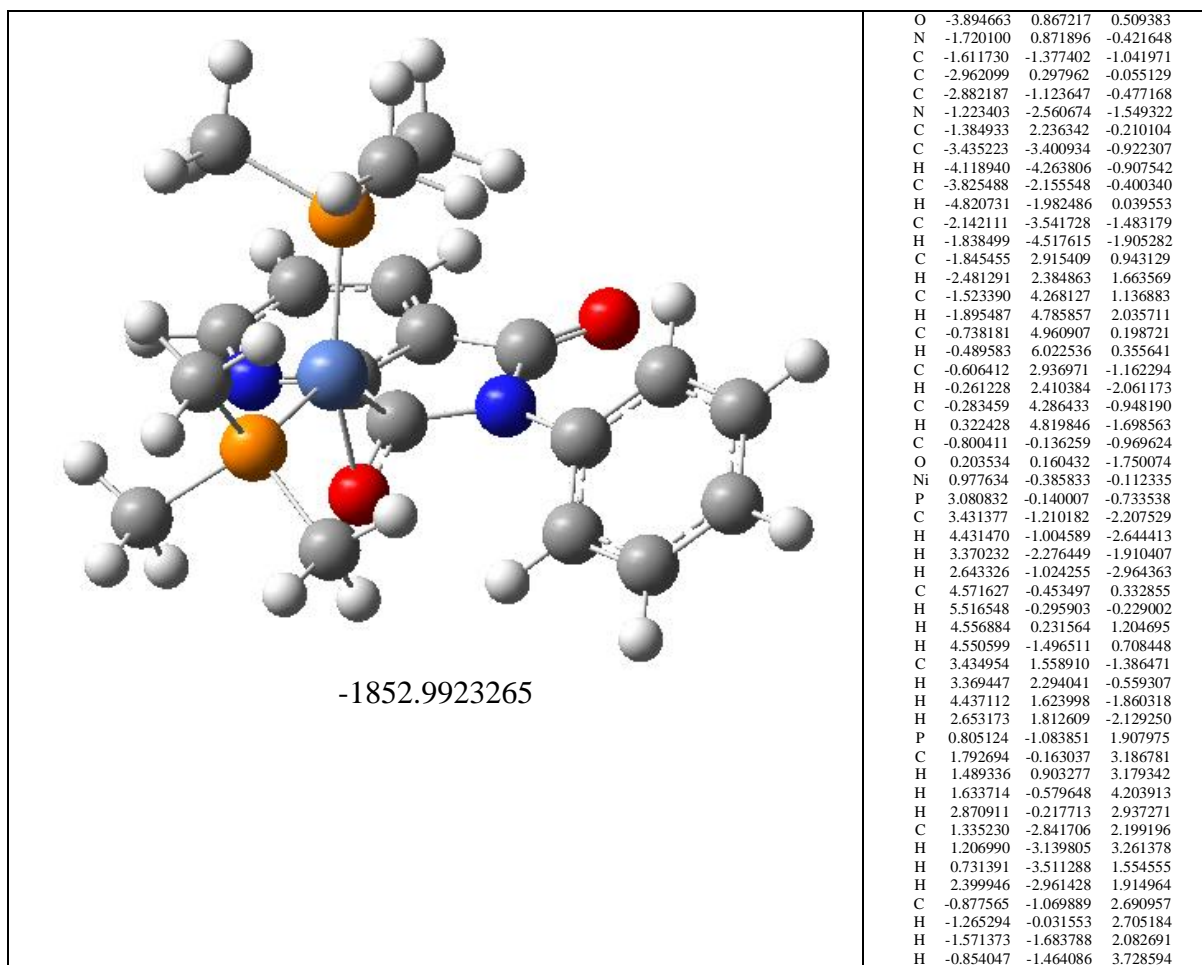
4



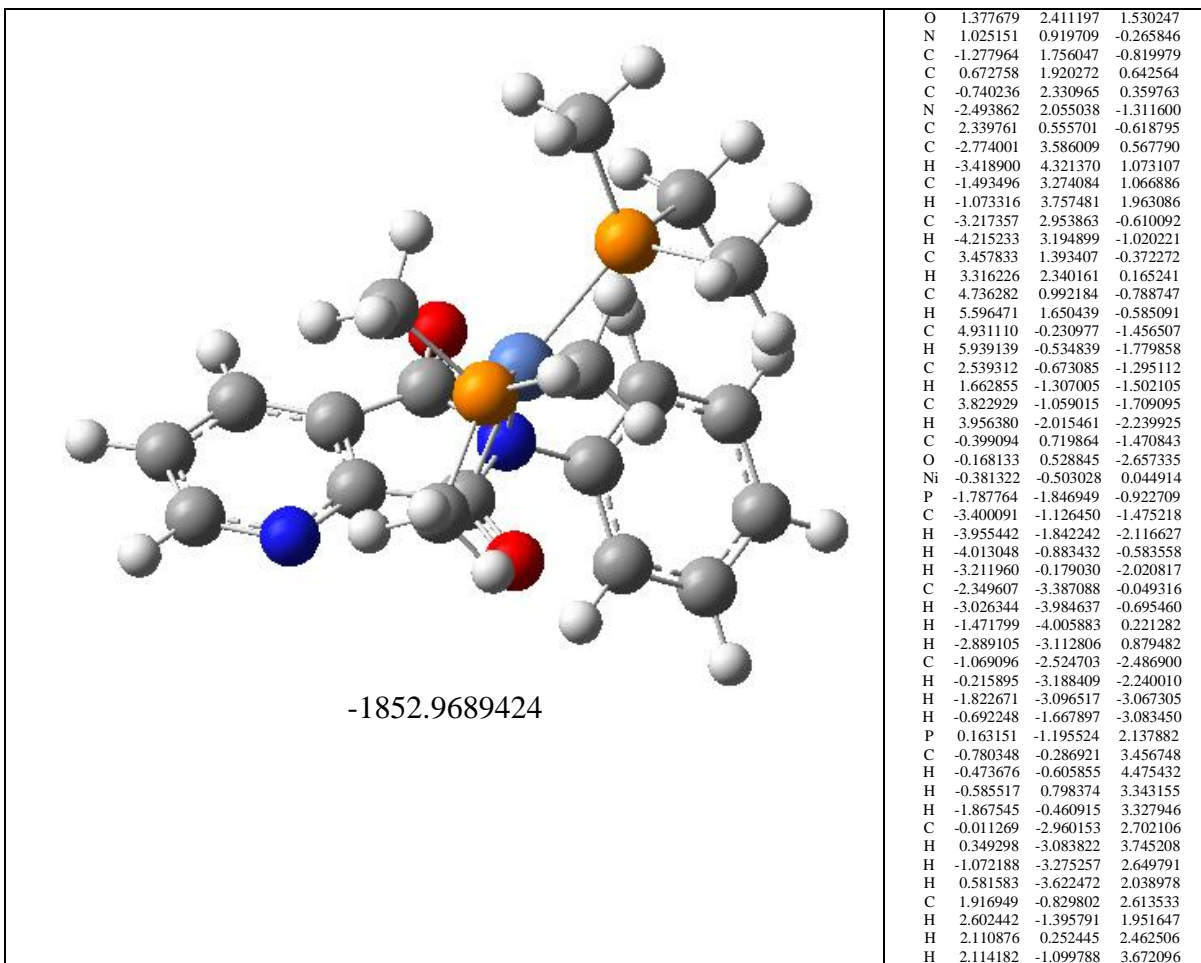
5



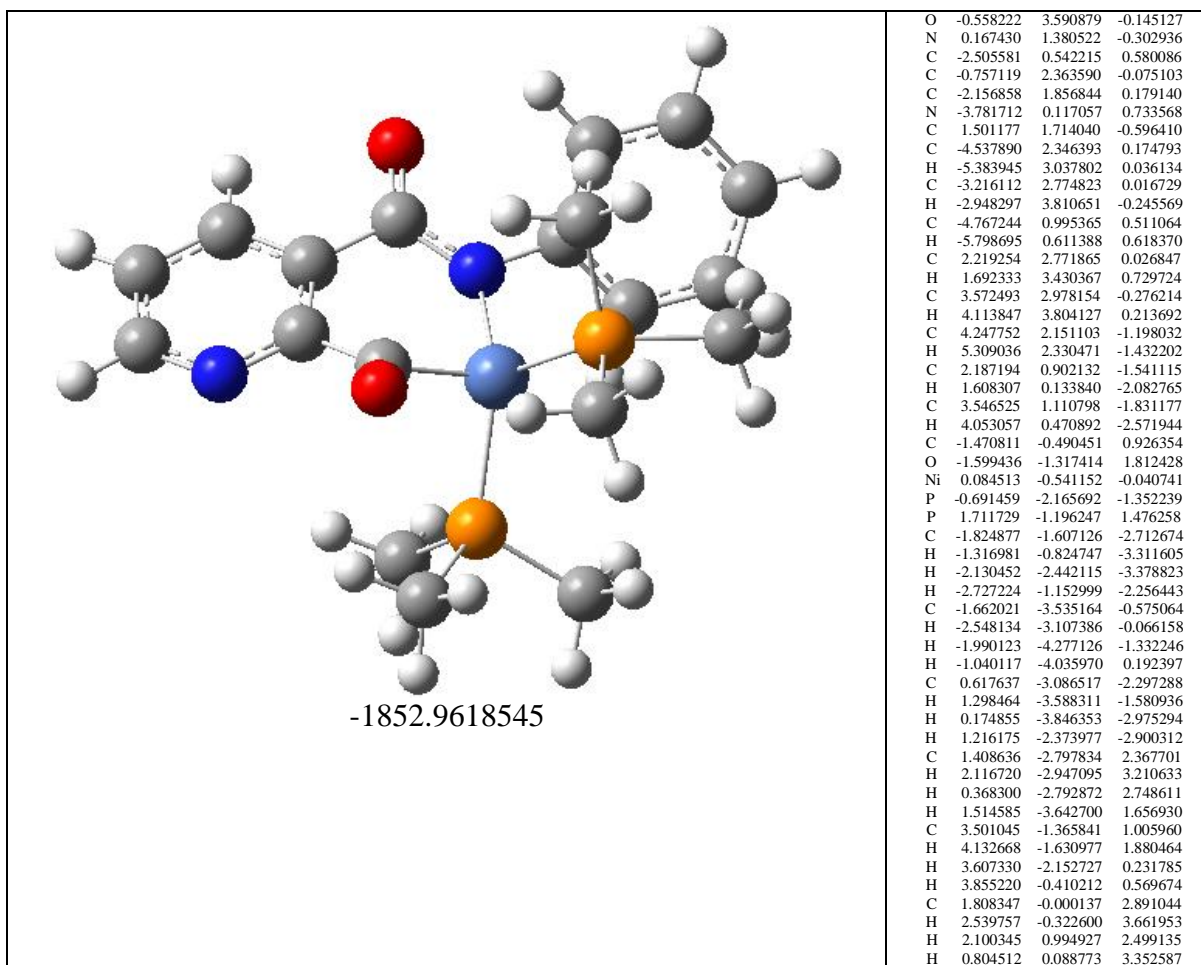
5isomer



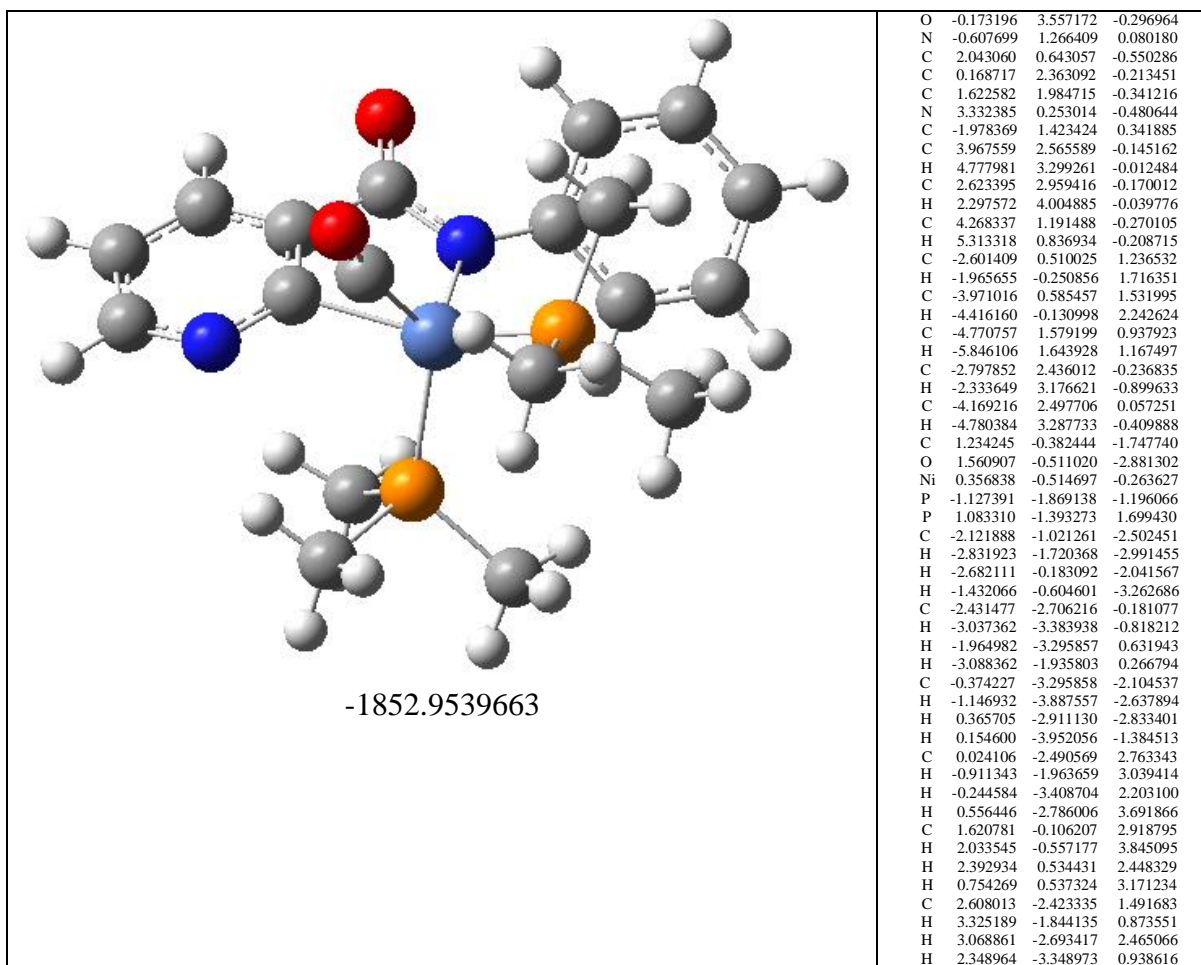
5-6



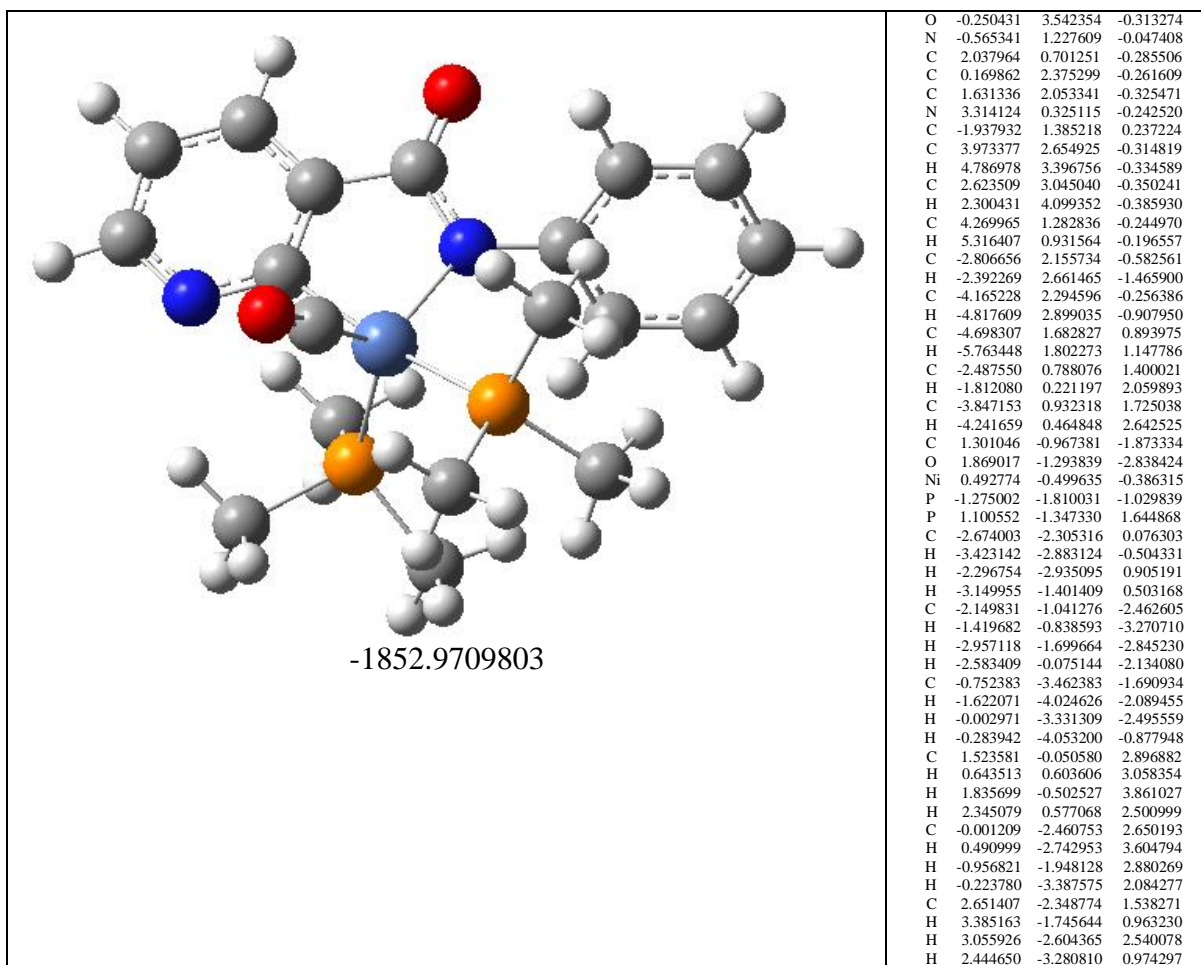
6



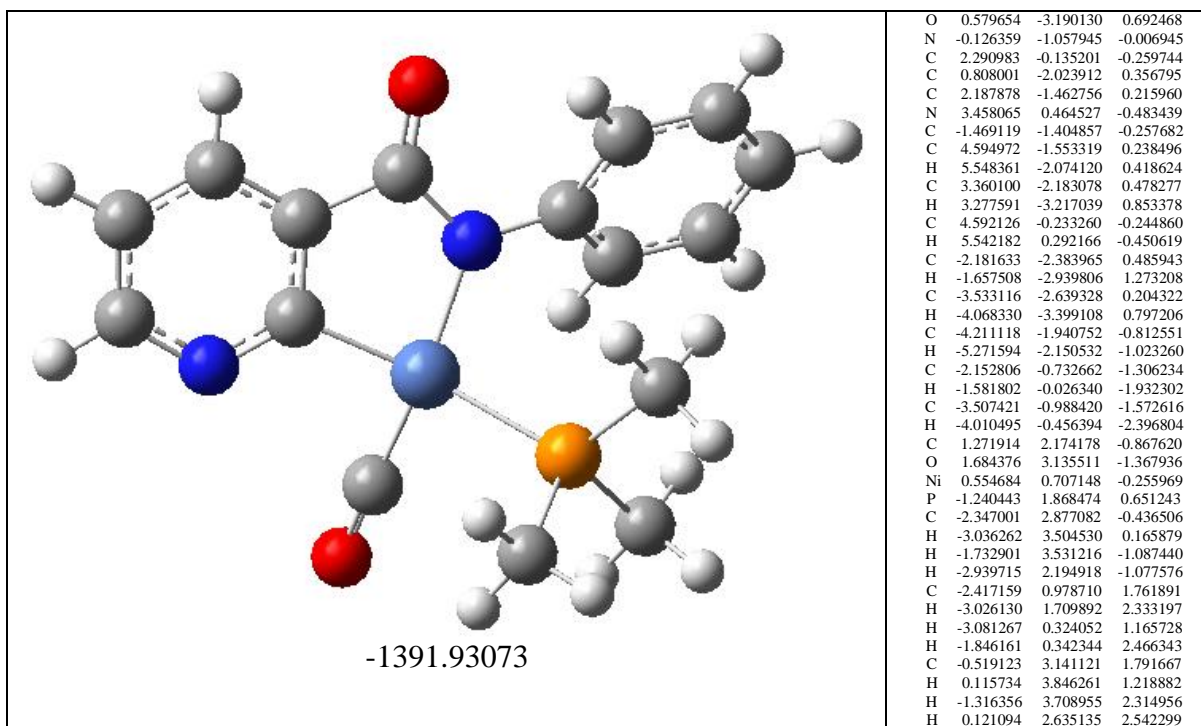
6-7

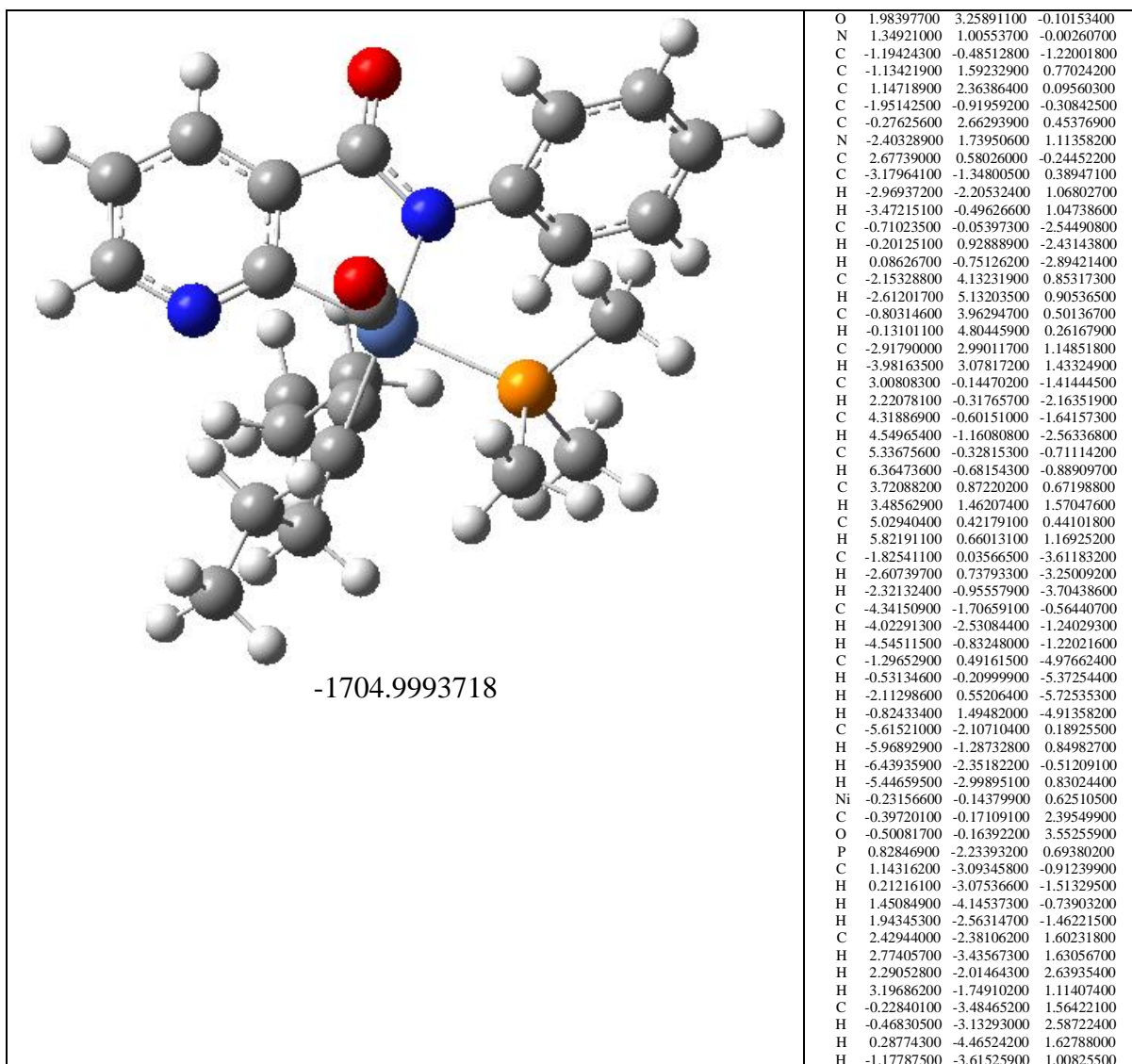


7

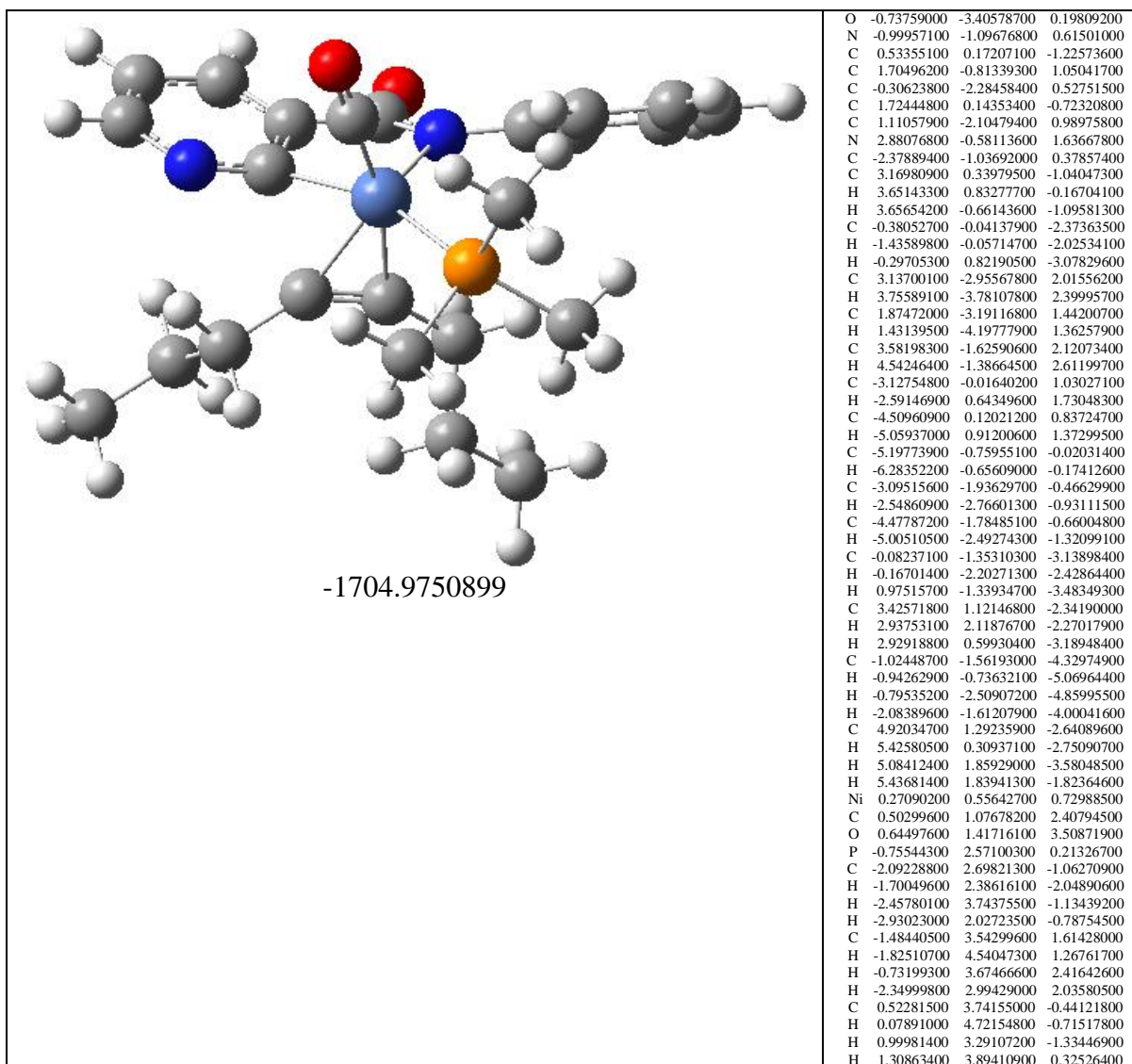


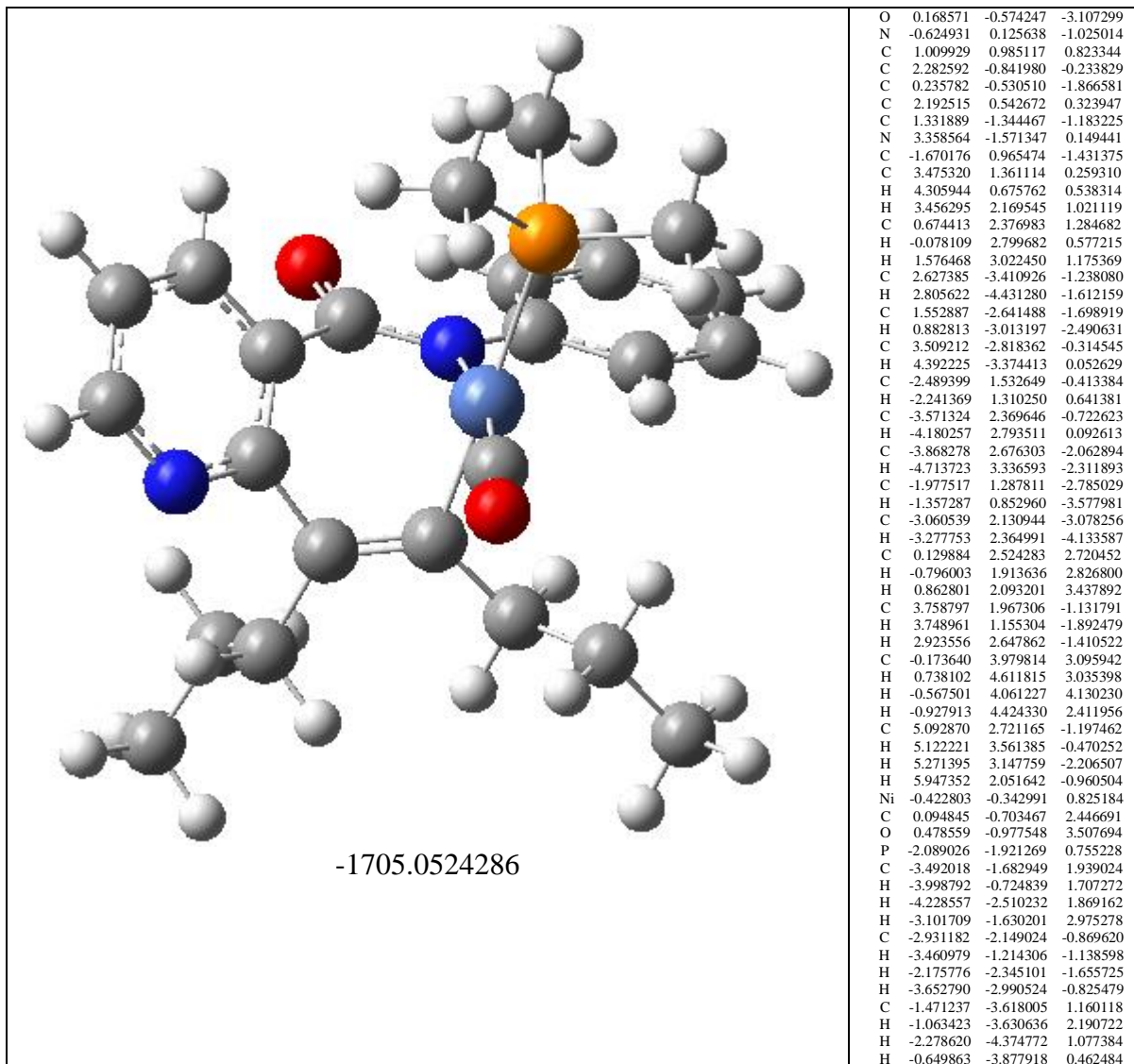
8



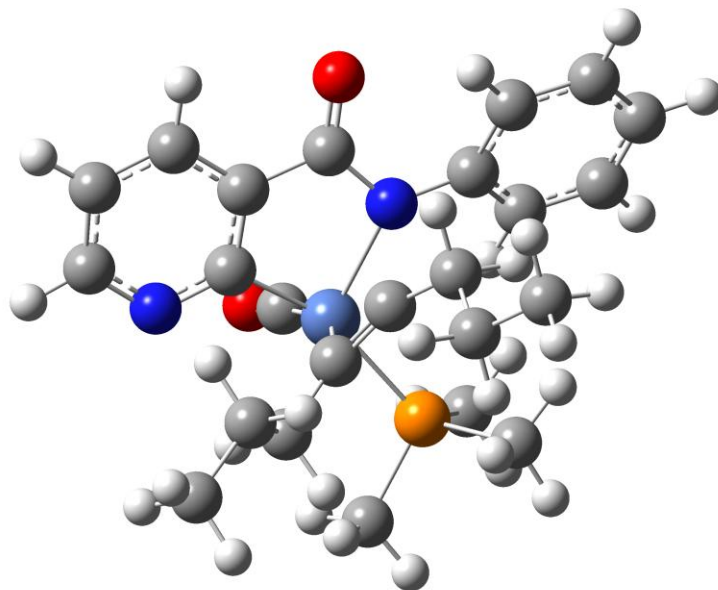


9-10





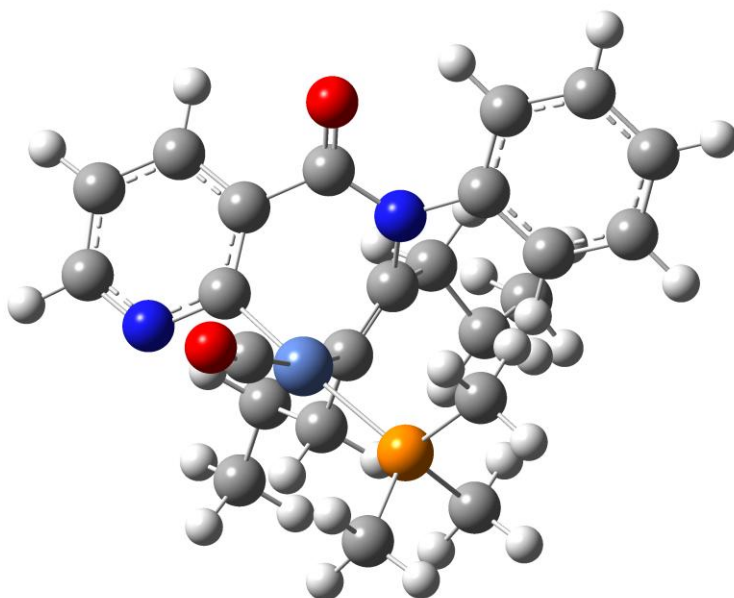
9-10'



-1704.9607339

O	-0.77568500	-2.63723100	2.24746800
N	-0.93355500	-0.68849400	0.92974900
C	0.00297500	-0.91107300	-1.01127300
C	1.76572300	-0.16859000	1.32921200
C	-0.26502900	-1.64963700	1.69935300
C	1.03906400	-0.14482300	-1.04937900
C	1.19831500	-1.36046700	1.84763700
N	3.06617800	0.11253700	1.40775400
C	-2.34384700	-0.68886500	0.87460800
C	2.30632100	0.05589100	-1.83114600
H	2.08030800	0.65007900	-2.74817400
H	2.99944100	0.67351000	-1.21819500
C	-0.95625500	-1.87135000	-1.58446400
H	-1.01474300	-2.78334600	-0.94825700
H	-1.97877200	-1.43087600	-1.55270600
C	3.42010900	-1.97251700	2.59113200
H	4.12227000	-2.65483000	3.09510900
C	2.05054700	-2.26980900	2.49826500
H	1.61133800	-3.19020700	2.91695600
C	3.88183700	-0.77386200	2.01962300
H	4.95372600	-0.50686800	2.06038900
C	-3.16343000	-1.84993100	0.97777000
H	-2.68760800	-2.82009500	1.16186900
C	-4.56075900	-1.74574100	0.88520000
H	-5.17107900	-2.65971900	0.97192800
C	-5.18810500	-0.50224500	0.69098700
H	-6.28535100	-0.43311000	0.62299200
C	-2.99651700	0.55777100	0.68818100
H	-2.36187700	1.45533700	0.64740500
C	-4.39250400	0.65413300	0.59697300
H	-4.86482300	1.64143900	0.46384500
C	-0.62405700	-2.26710500	-3.04317200
H	0.38121300	-2.73959800	-3.07650000
H	-0.55480200	-1.34927800	-3.66767600
C	3.02914700	-1.24649100	-2.23540500
H	2.35873000	-1.85816700	-2.87749300
H	3.21081900	-1.84767000	-1.31852400
C	-1.66797600	-3.22558900	-3.62961600
H	-2.67944500	-2.76715400	-3.64189300
H	-1.41257400	-3.50705300	-4.67196200
H	-1.73597100	-4.16179200	-3.03633300
C	4.35247800	-0.98224100	-2.96329200
H	5.05348900	-0.40053300	-2.32772500
H	4.85951400	-1.92999200	-3.23887800
H	4.19515100	-0.40368800	-3.89920200
Ni	0.47319400	1.00793100	0.50848200
C	0.78801800	2.13230800	1.80505600
O	1.05059700	2.90665600	2.63507800
P	-0.37316600	2.63678300	-0.85593100
C	-1.35973900	2.11663300	-2.33302800
H	-0.78251500	1.35979600	-2.90045700
H	-1.59307100	2.97867300	-2.99155800
H	-2.30131900	1.64408100	-1.98842000
C	-1.39660100	4.00643600	-0.13409800
H	-1.62348300	4.78384700	-0.89288800
H	-0.84154000	4.47026400	0.70583100
H	-2.34890800	3.60523600	0.26576200
C	1.02449400	3.59577700	-1.60507500
H	1.61697900	4.06082800	-0.79203000
H	0.65738500	4.38570200	-2.29336200
H	1.69267300	2.90602000	-2.15556900

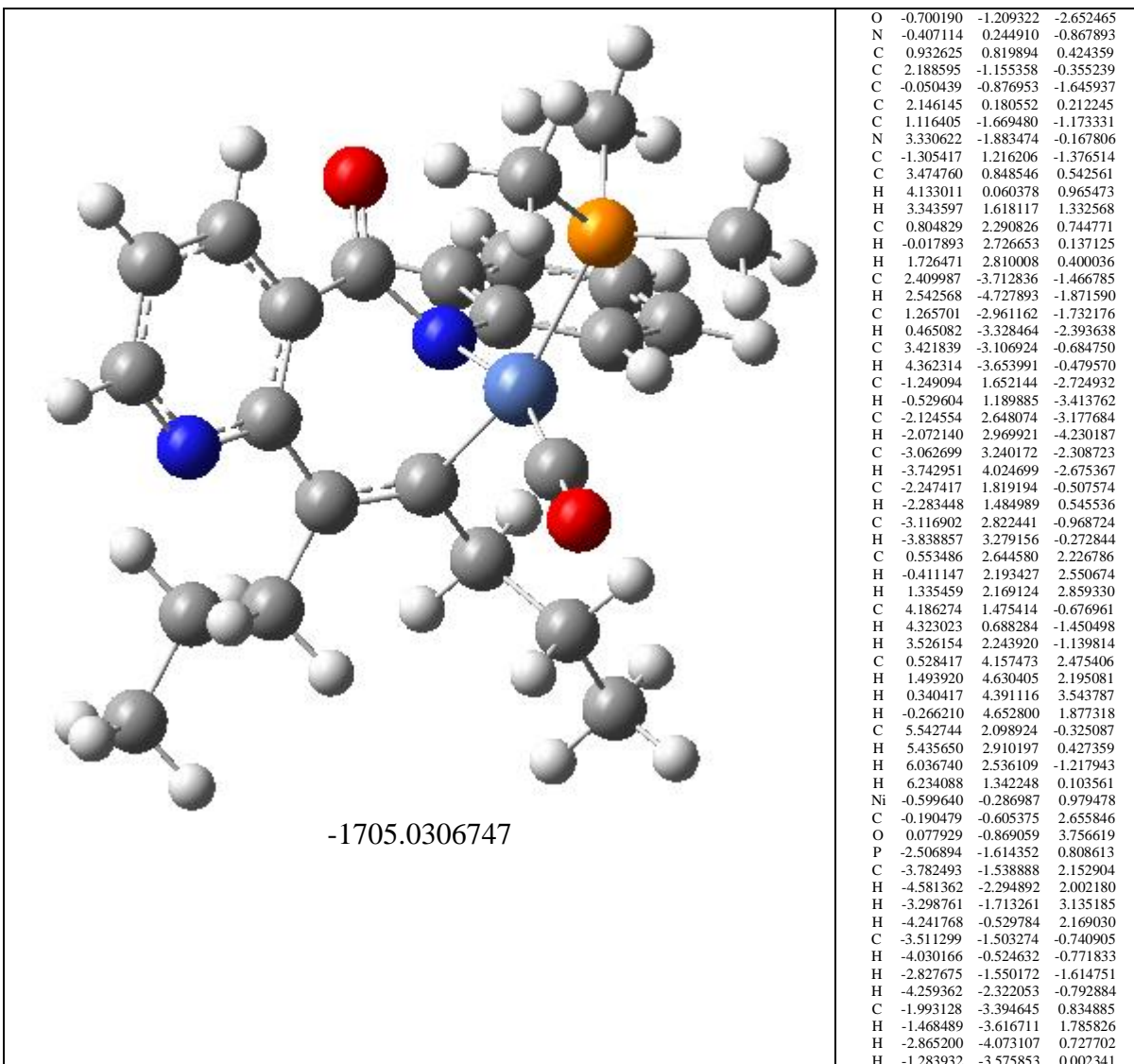
10⁶



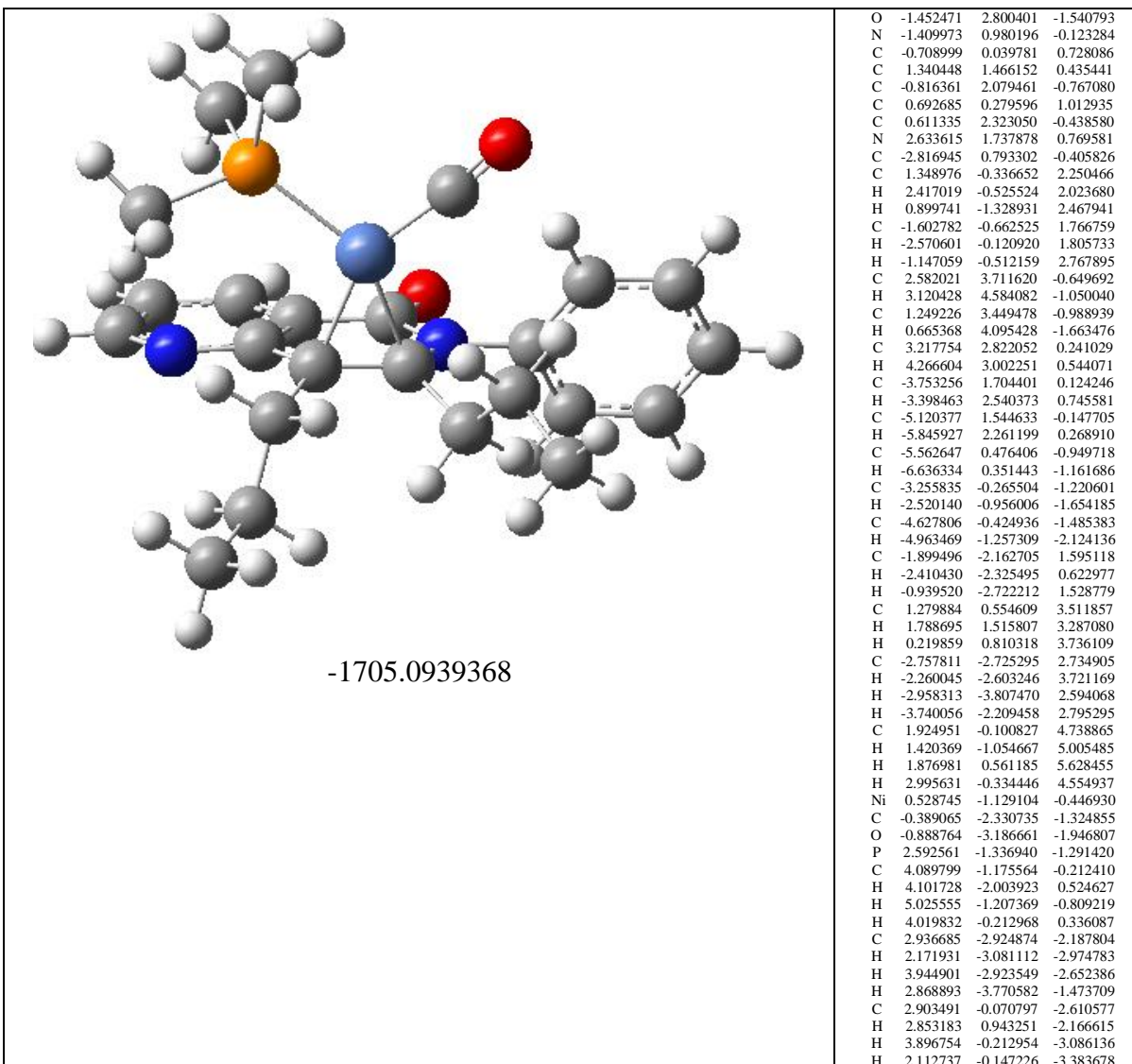
-1705.0209526

O	-0.73543400	-0.89295100	3.26054000
N	-1.07966700	-0.77434100	0.95786400
C	-0.54624100	-1.11404500	-0.36441200
C	1.95950600	-0.36951400	1.01257700
C	-0.28560100	-0.93369500	2.11173800
C	0.51541000	-0.41149300	-0.84863800
C	1.18647900	-1.15826900	1.90584300
N	3.27921700	-0.54726000	0.85836400
C	-2.46016200	-0.39821900	1.02549400
C	1.32141300	-0.76216100	-2.09227800
H	0.66534000	-0.71115200	-2.99275400
H	2.09431100	0.02408700	-2.25262400
C	-1.20589800	-2.36439100	-0.95422400
H	-0.53797900	-3.23741000	-0.75609300
H	-2.12374700	-2.57077600	-0.36309000
C	3.23854200	-2.27409800	2.55521700
H	3.79301900	-3.00887500	3.15916400
C	1.85540400	-2.08940400	2.72013600
H	1.28162400	-2.66316400	3.46555800
C	3.90081000	-1.49331400	1.59408300
H	4.98419600	-1.62091800	1.41433400
C	-3.25739800	-0.50947500	2.19496600
H	-2.80353700	-0.89163300	3.11452000
C	-4.60505500	-0.11445600	2.16990200
H	-5.19934600	-0.21028700	3.09310500
C	-5.20198500	0.38817900	1.00151700
H	-6.26100100	0.69030400	0.99581400
C	-3.07223100	0.10839900	-0.15053100
H	-2.46873100	0.19290800	-1.06562700
C	-4.42183100	0.49226300	-0.16233500
H	-4.86414600	0.87841700	-1.09513400
C	-1.59991600	-2.36673500	-2.44367200
H	-0.69938900	-2.28948600	-3.08875500
H	-2.21120400	-1.46320400	-2.66301600
C	2.04910200	-2.12506400	-2.07755700
H	1.31057100	-2.94704000	-1.95866000
H	2.70418100	-2.16727100	-1.18317900
C	-2.38810500	-3.62491800	-2.83322300
H	-3.32773400	-3.71087300	-2.24692200
H	-2.66252500	-3.61892500	-3.90854200
H	-1.79662700	-4.54665300	-2.64544700
C	2.88109200	-2.36264200	-3.34403300
H	3.65967800	-1.57961100	-3.46923500
H	3.39916300	-3.34372700	-3.31215600
H	2.24905600	-2.35016600	-4.25872300
Ni	1.19178200	1.13246000	0.09601900
C	1.93355700	2.27369900	1.23803200
O	2.40337100	2.97883000	2.03096900
P	0.15407800	2.67980000	-1.17946400
C	-0.53501400	2.19075000	-2.82462300
H	0.29193100	1.86026700	-3.48370200
H	-1.07084000	3.03537500	-3.30461000
H	-1.22939600	1.33827900	-2.69147900
C	-1.29739000	3.44050400	-0.32120900
H	-1.76710400	4.22950300	-0.94484100
H	-0.96344600	3.88279600	0.63875200
H	-2.04448400	2.65297300	-0.09537400
C	1.19084100	4.15323400	-1.61888100
H	1.55731300	4.63494600	-0.69031700
H	0.61562100	4.89530600	-2.21082900
H	2.07186900	3.82347200	-2.20570800

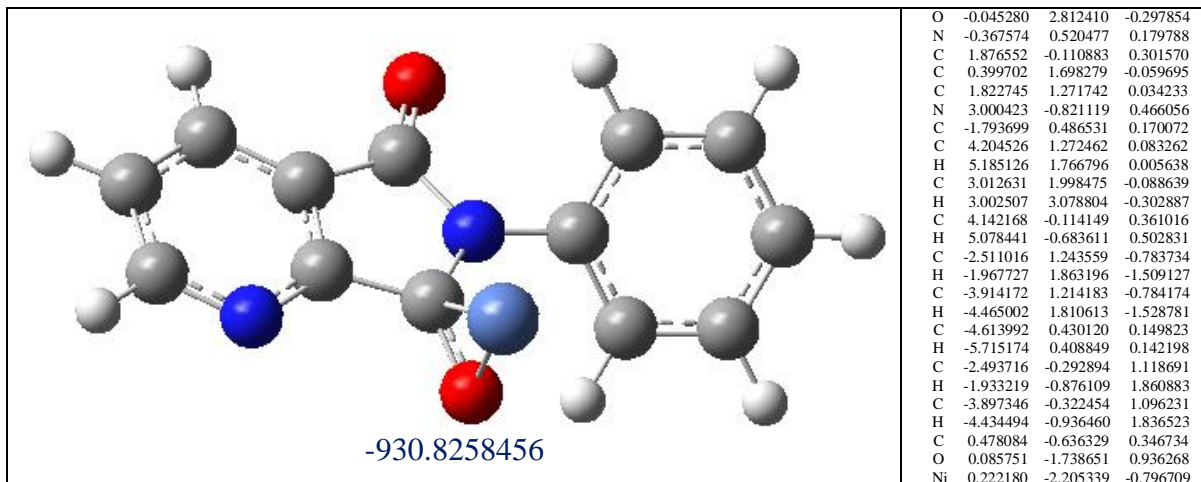
10-11



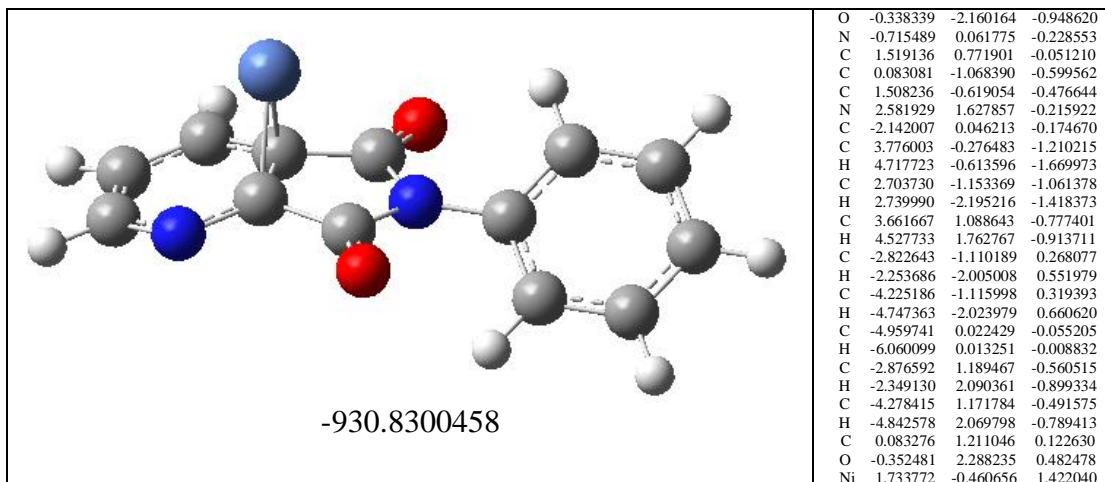
11



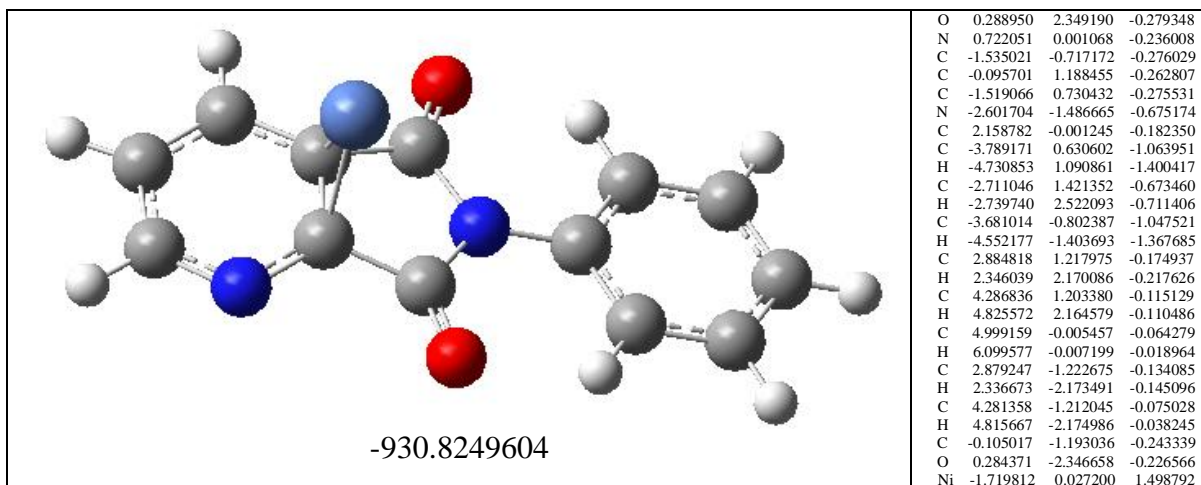
12



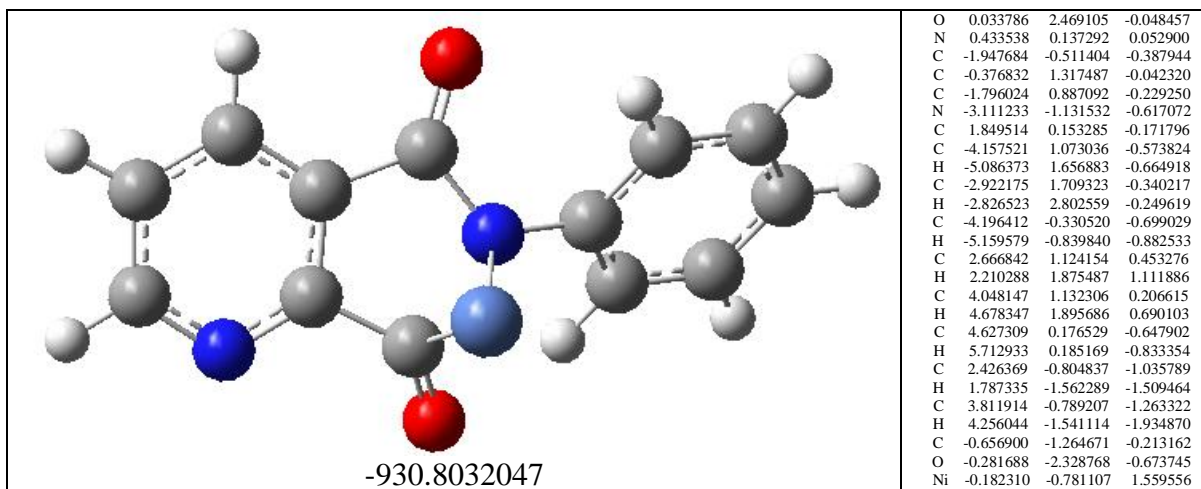
12isomer



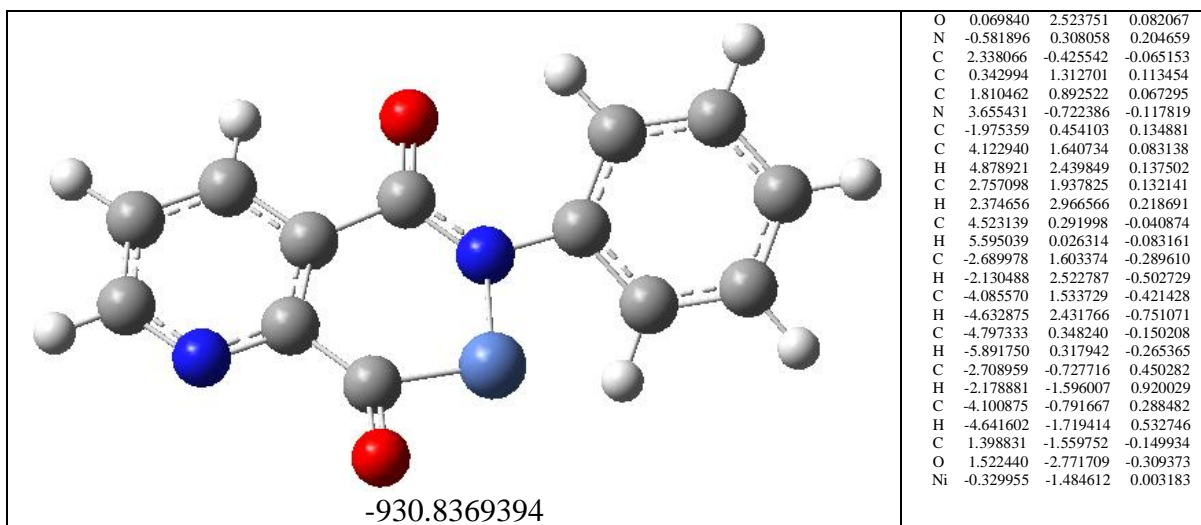
12isomerTS (rotation phenyl, barrier of 0.1 kcal/mol)



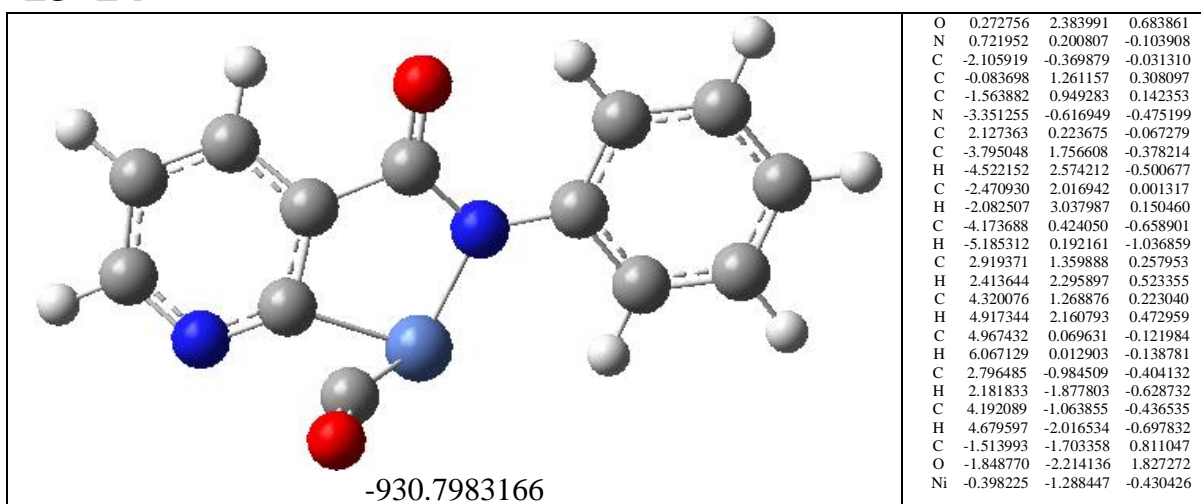
12-13



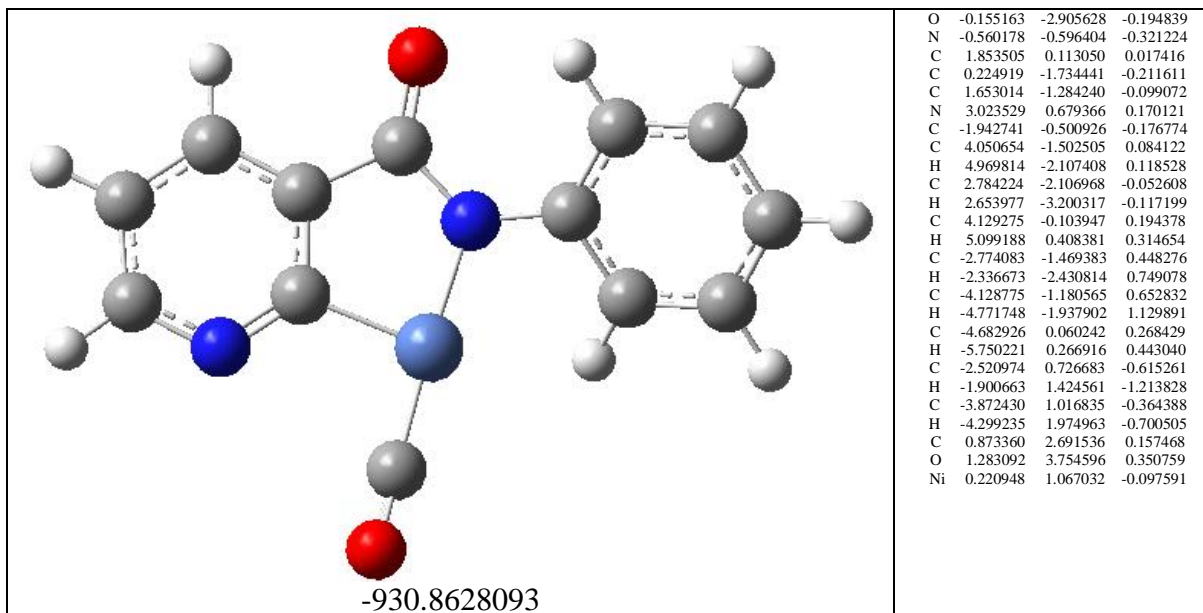
13

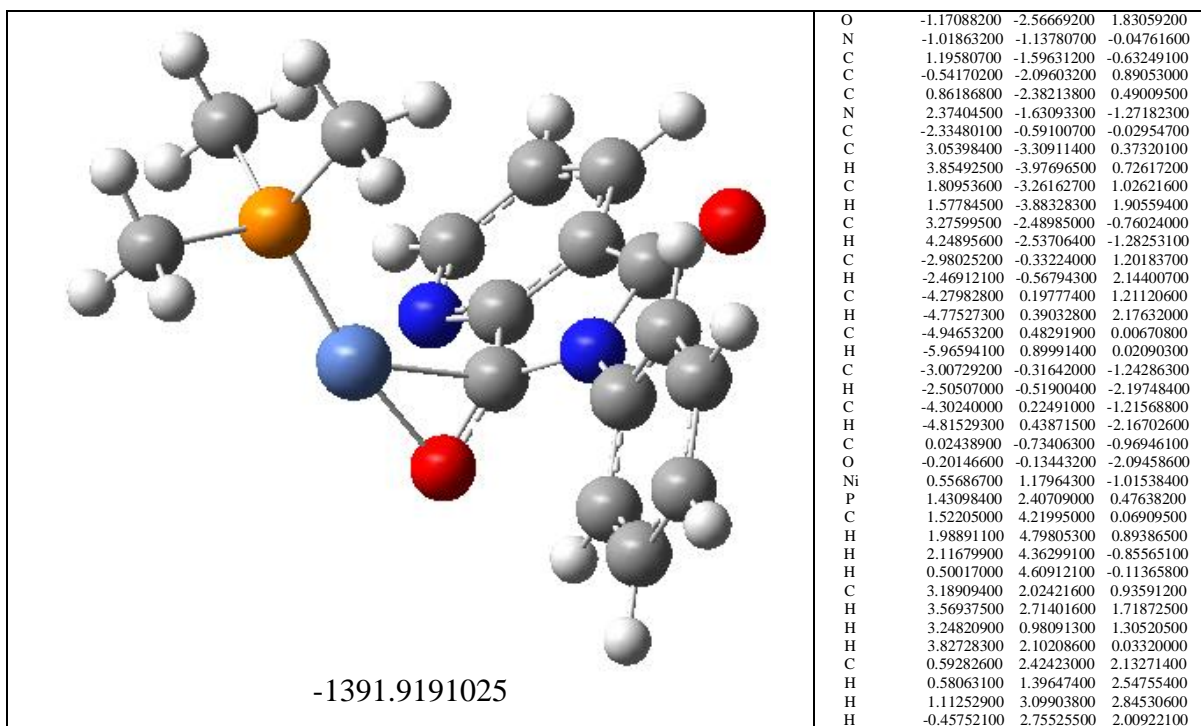


13-14

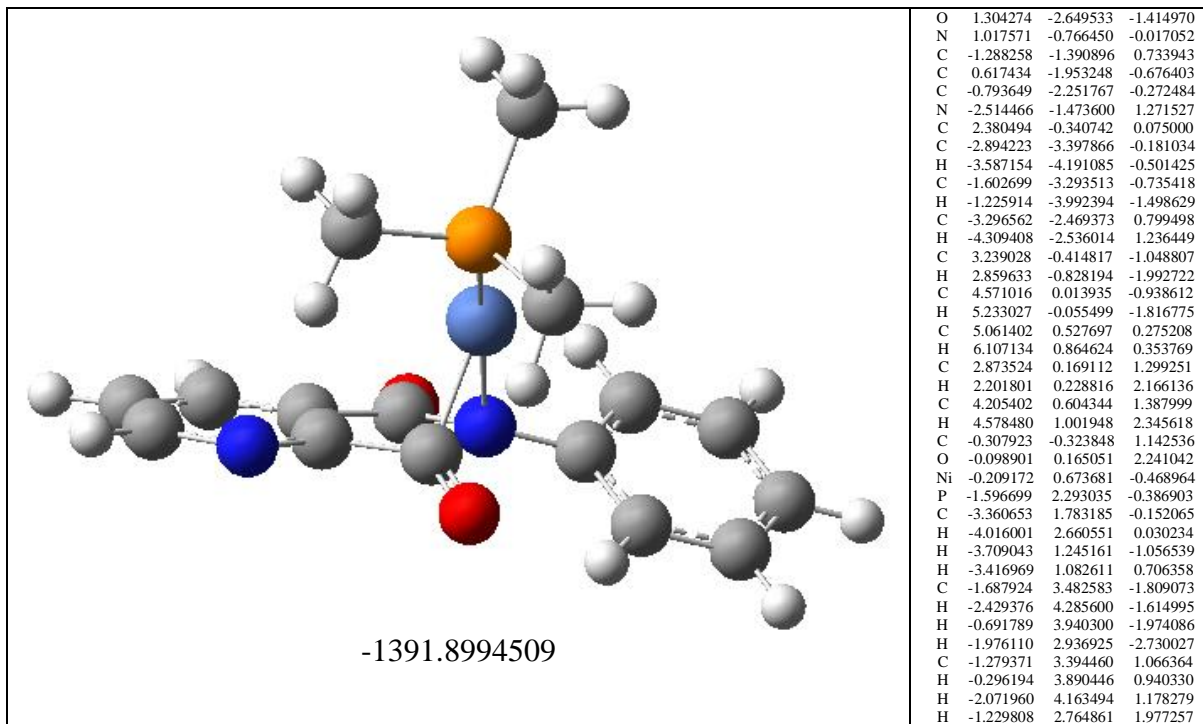


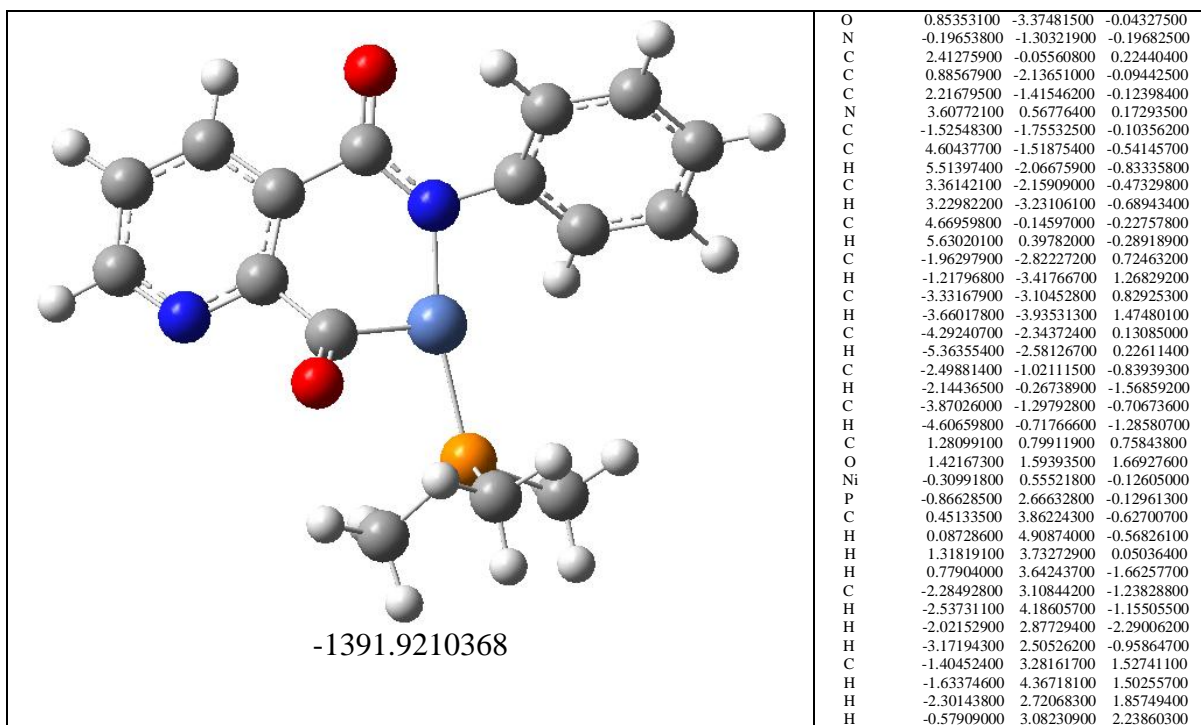
14



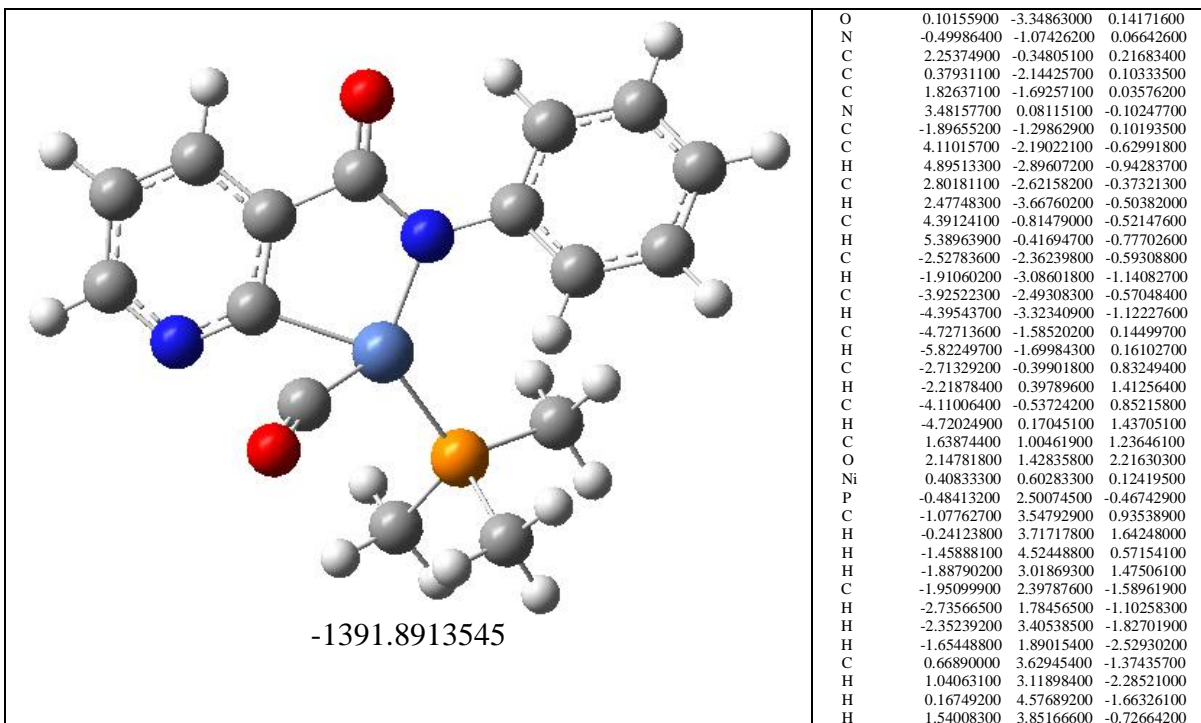


15-16

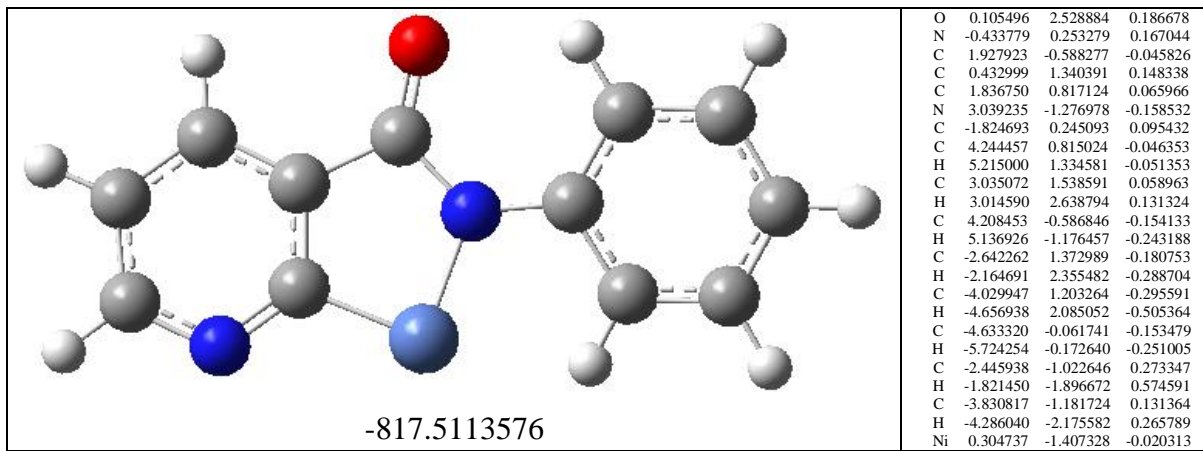


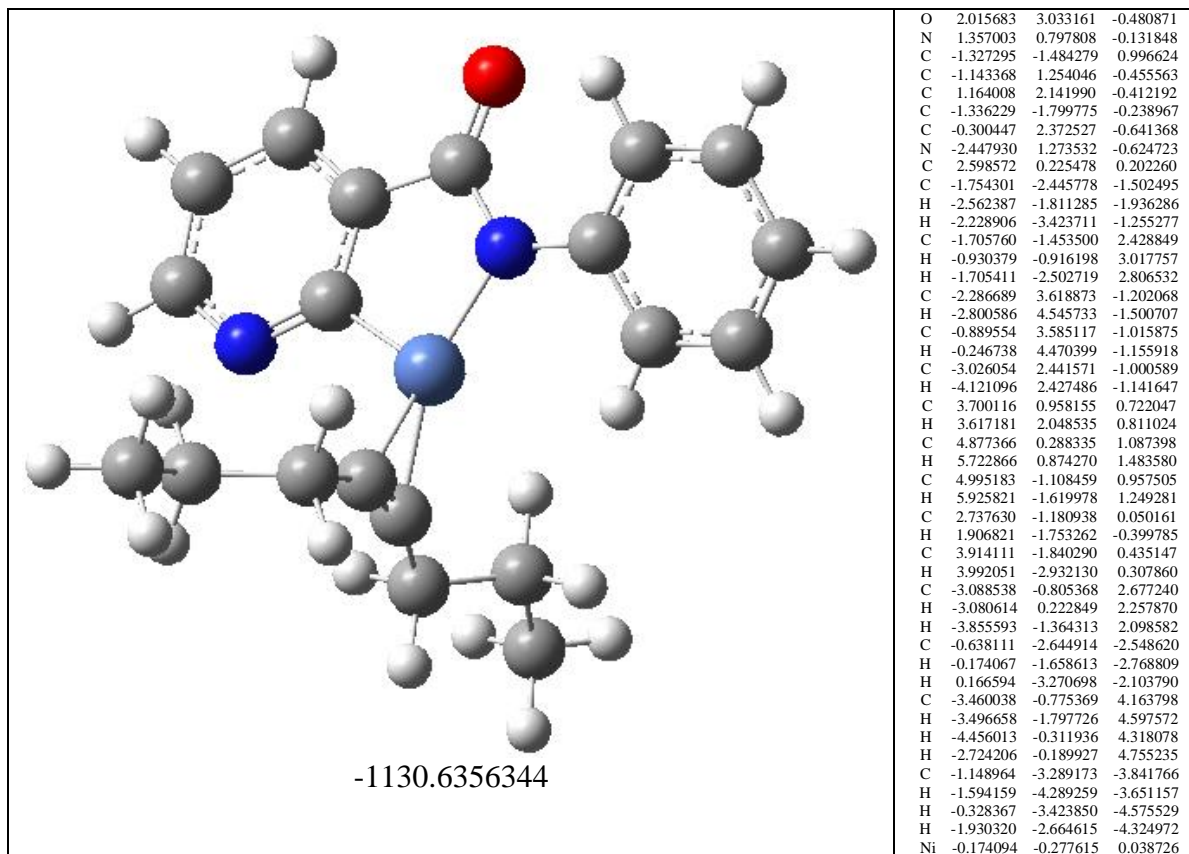


16-8

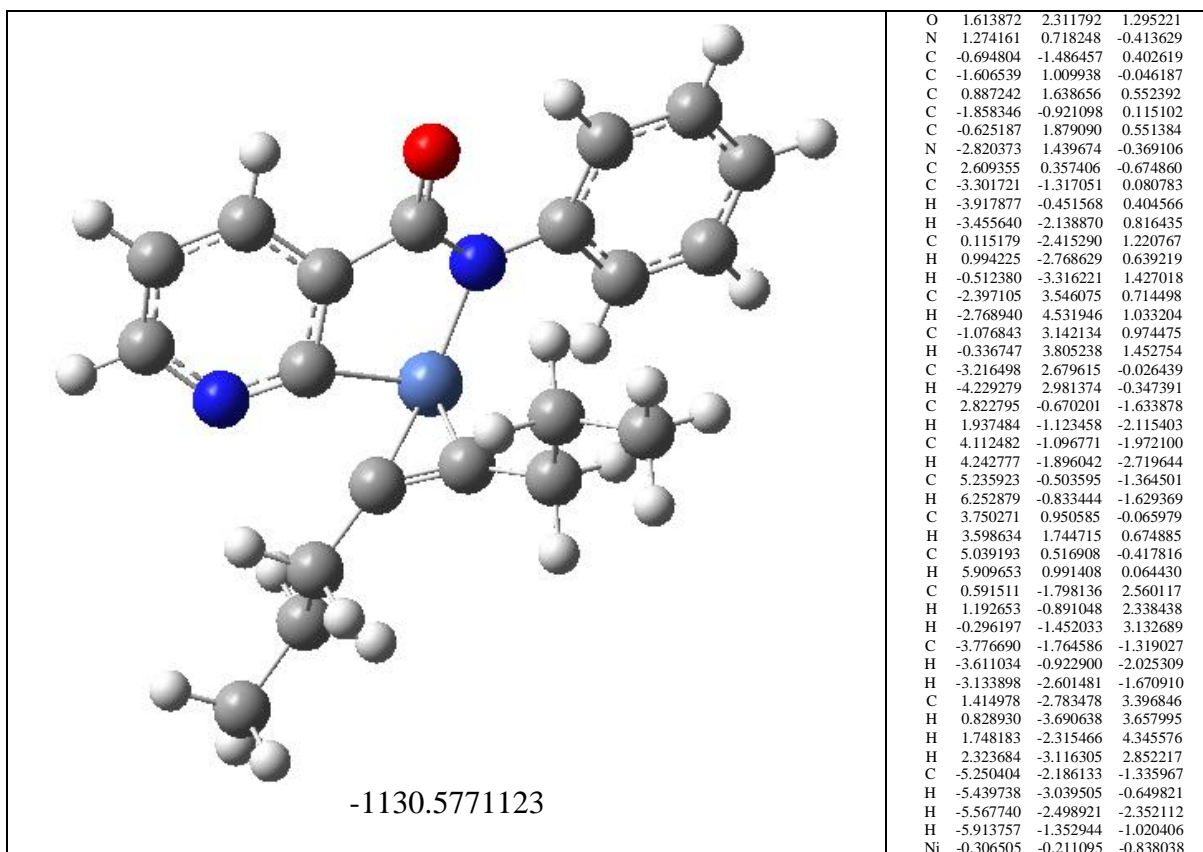


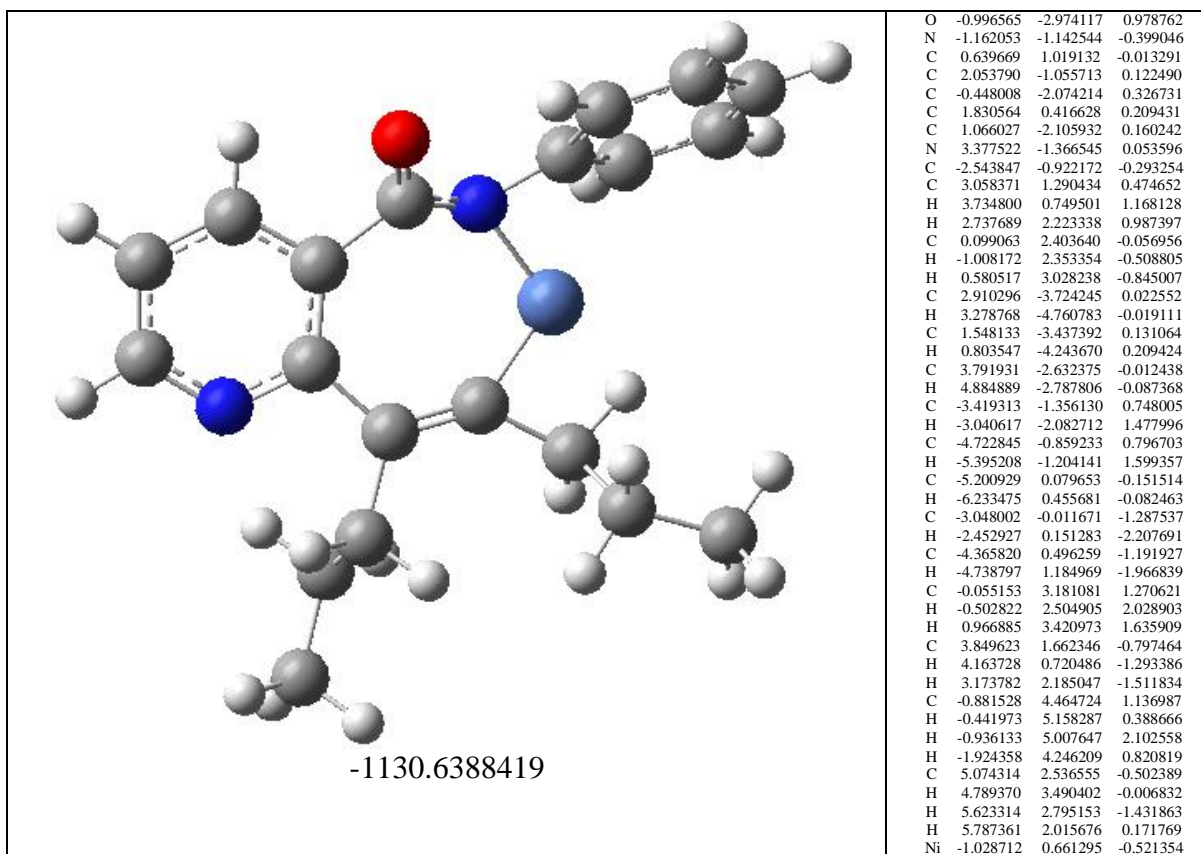
17



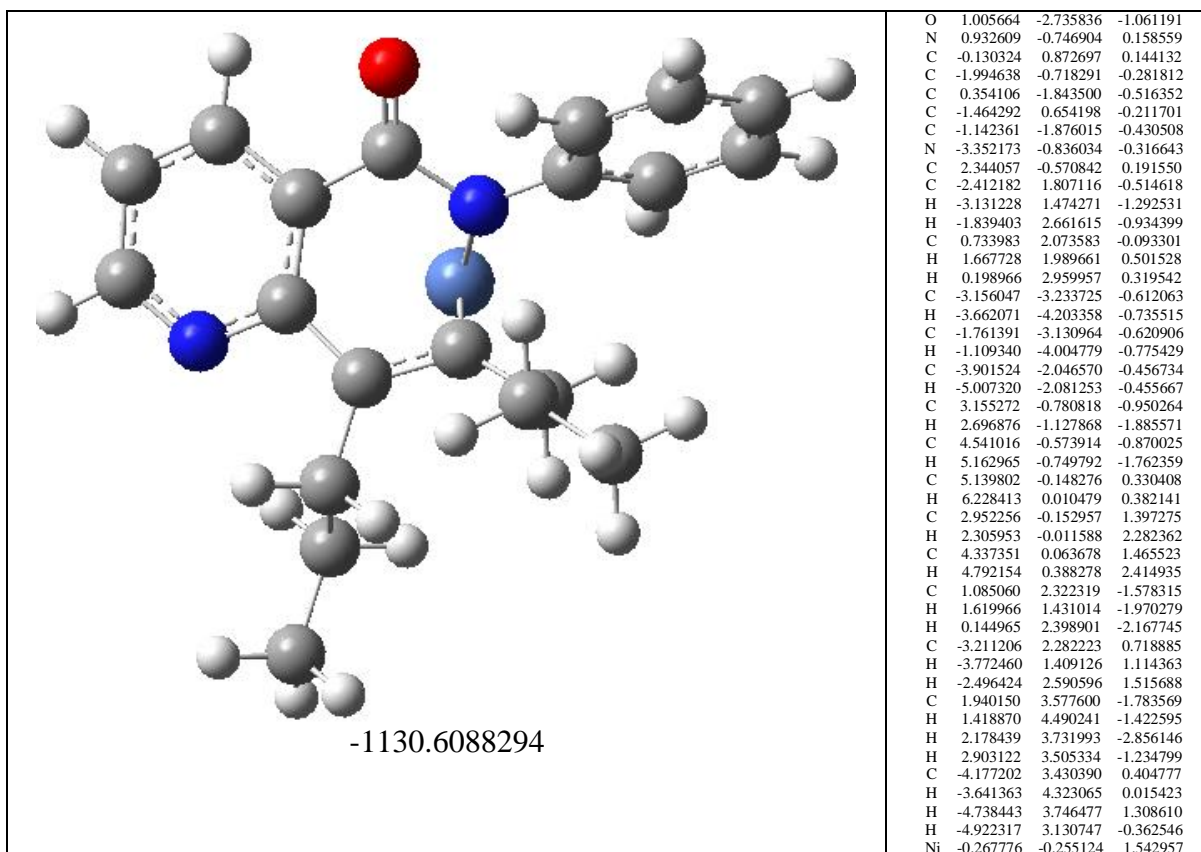


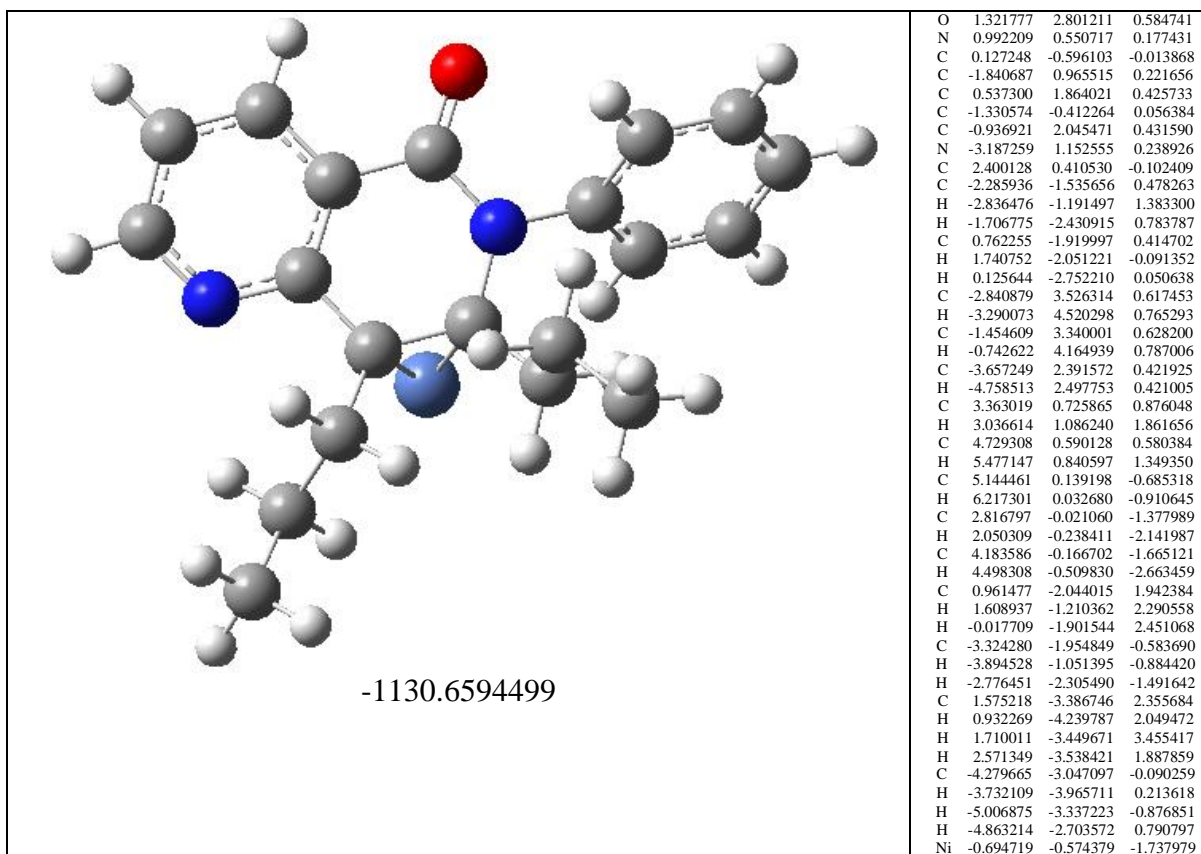
18-19

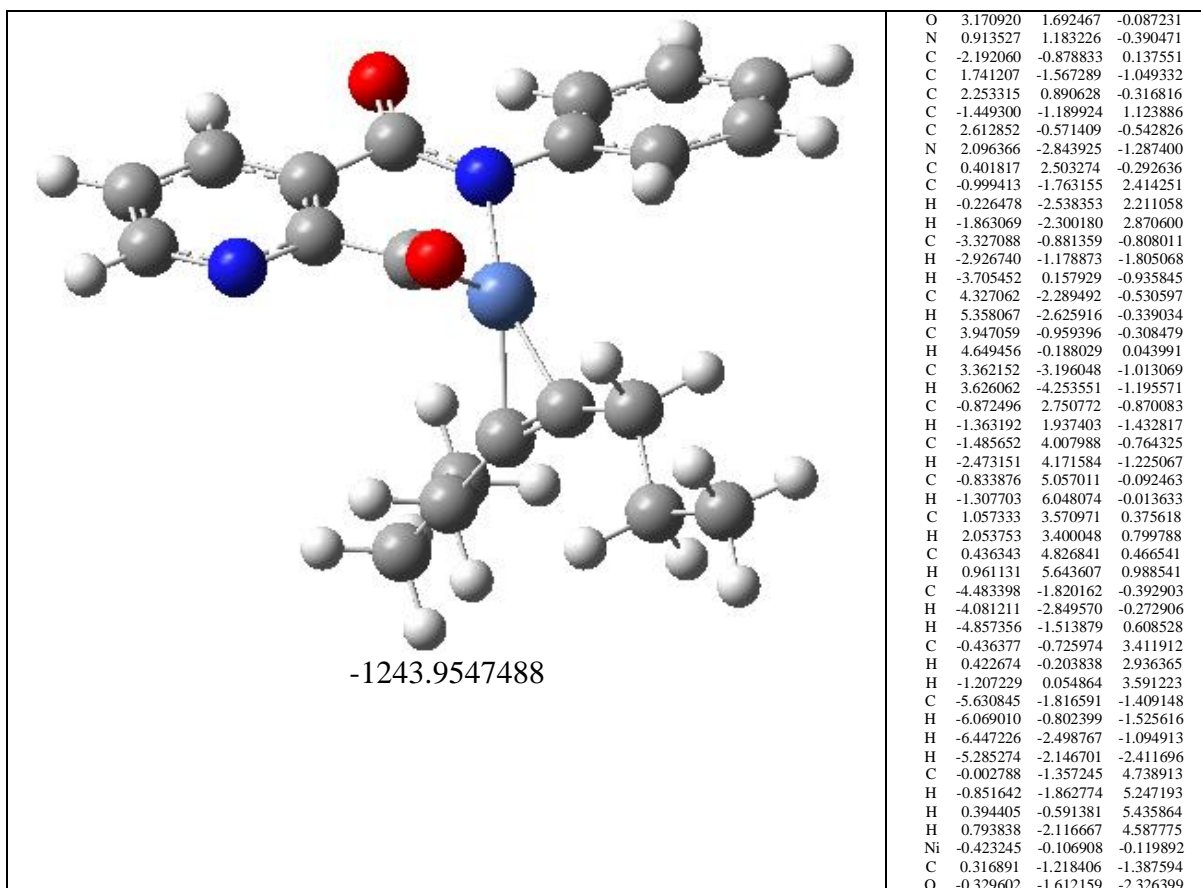




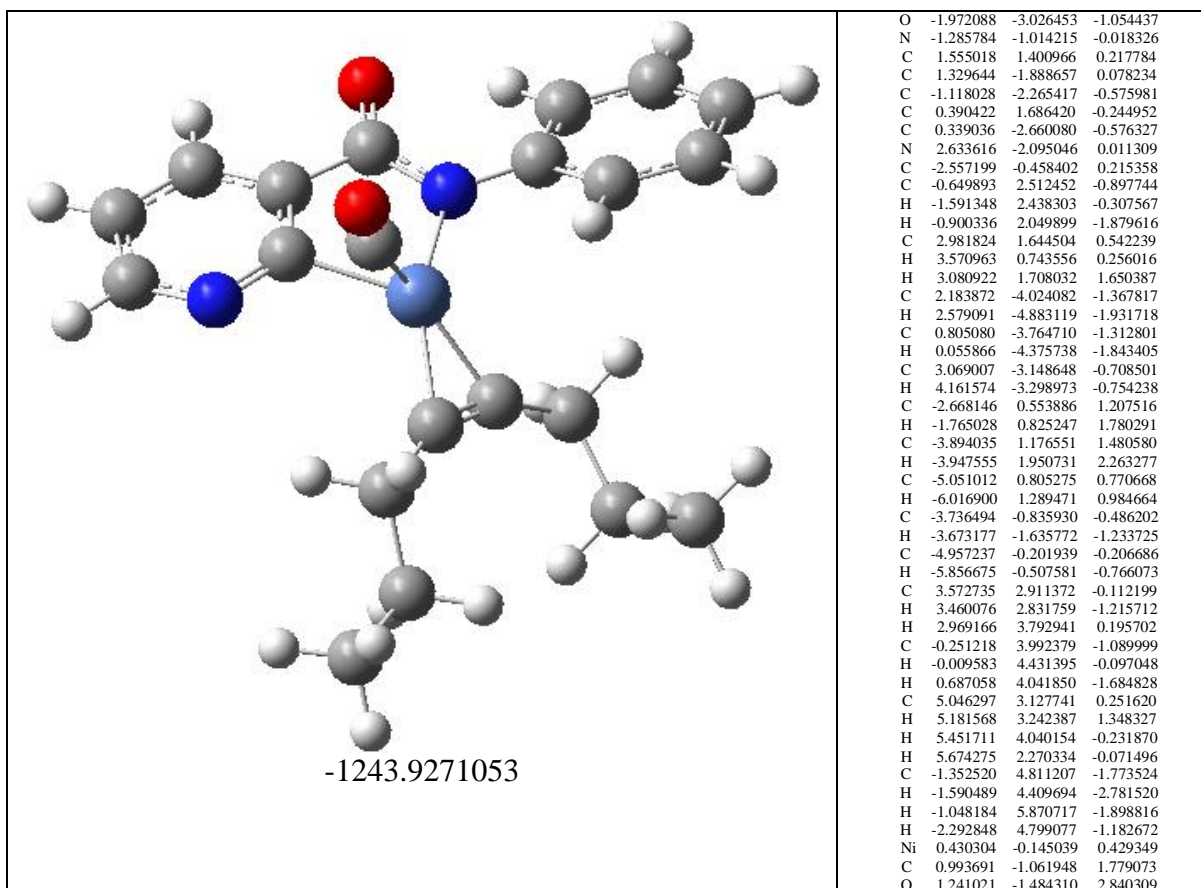
19-20

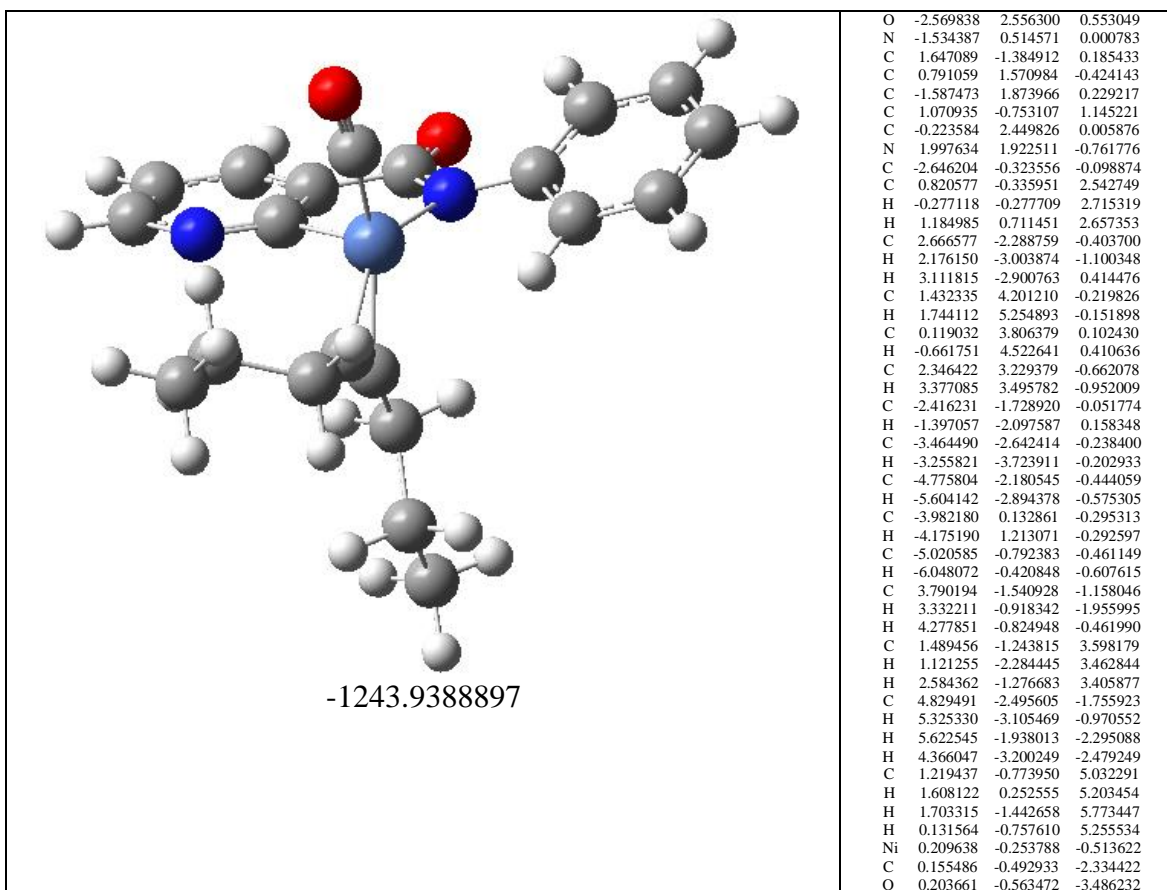




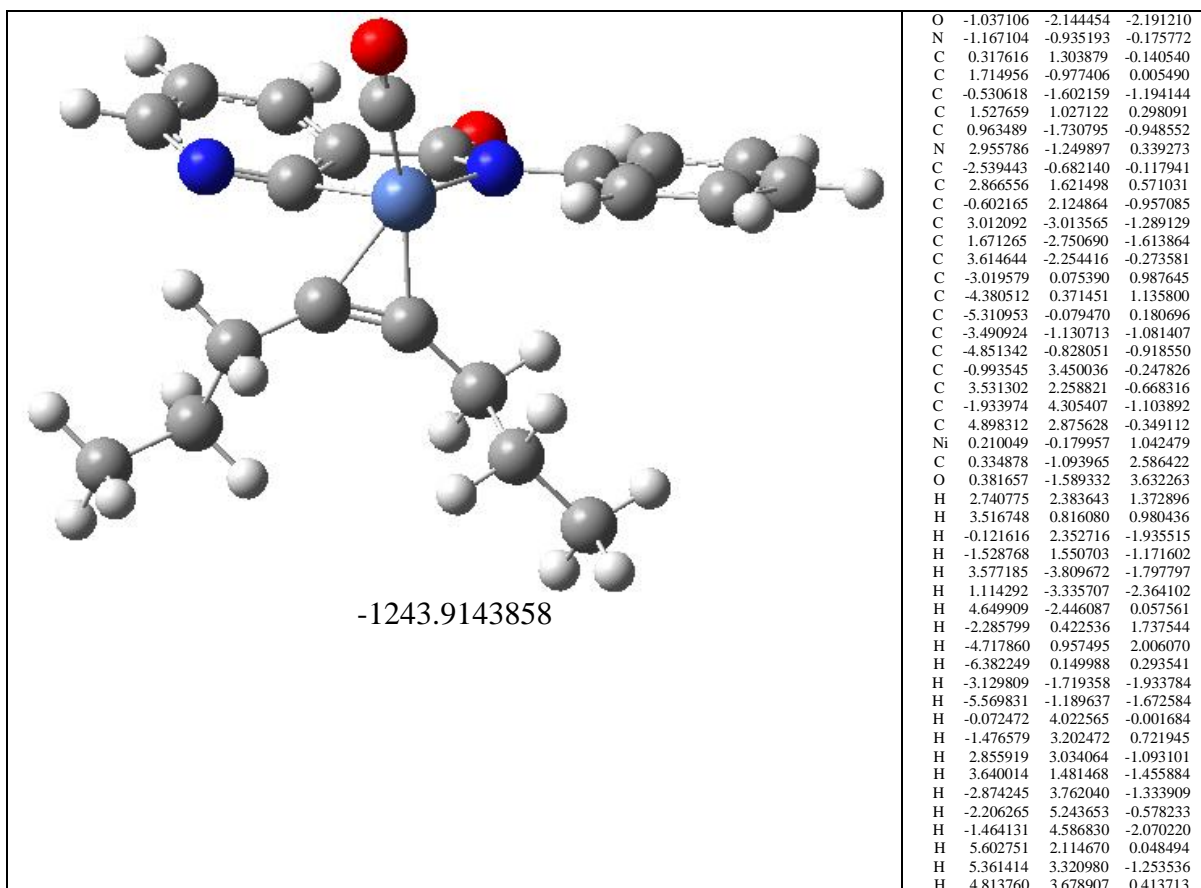


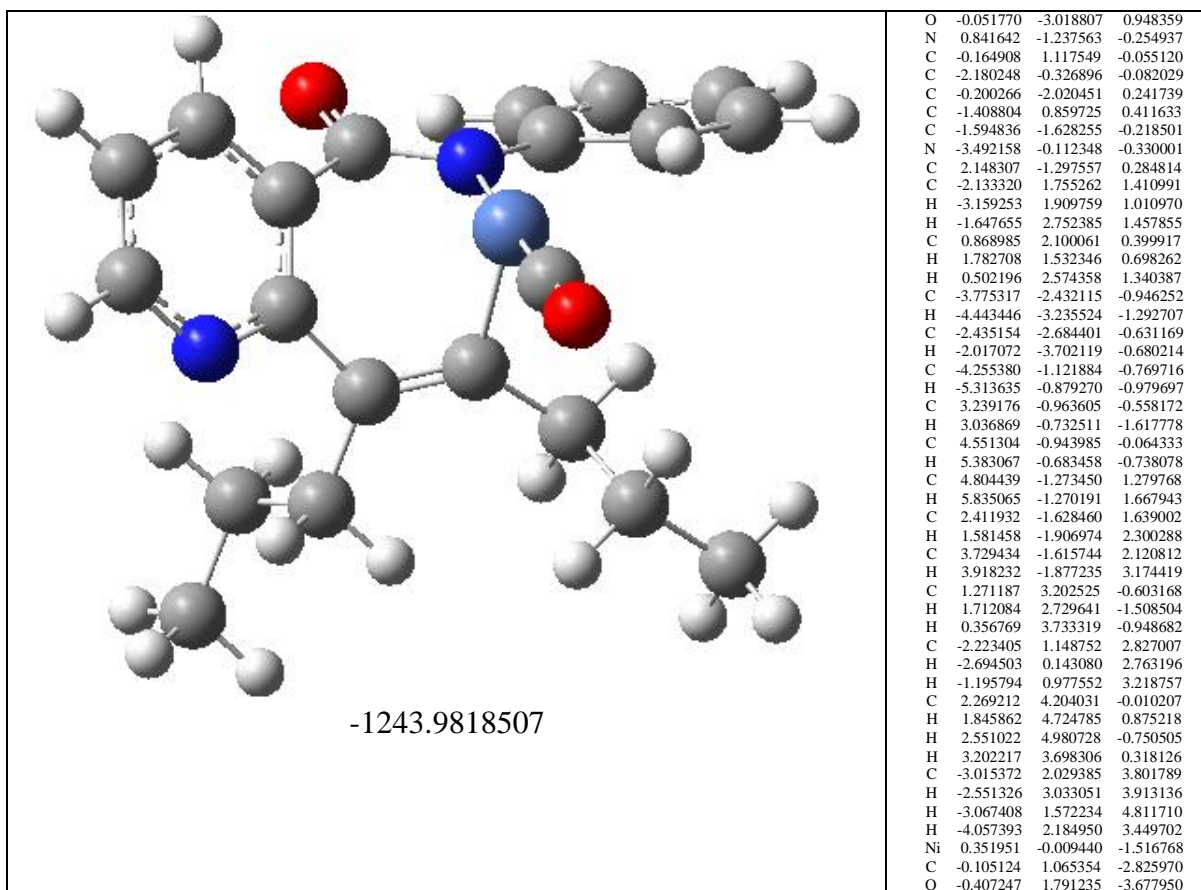
21-22



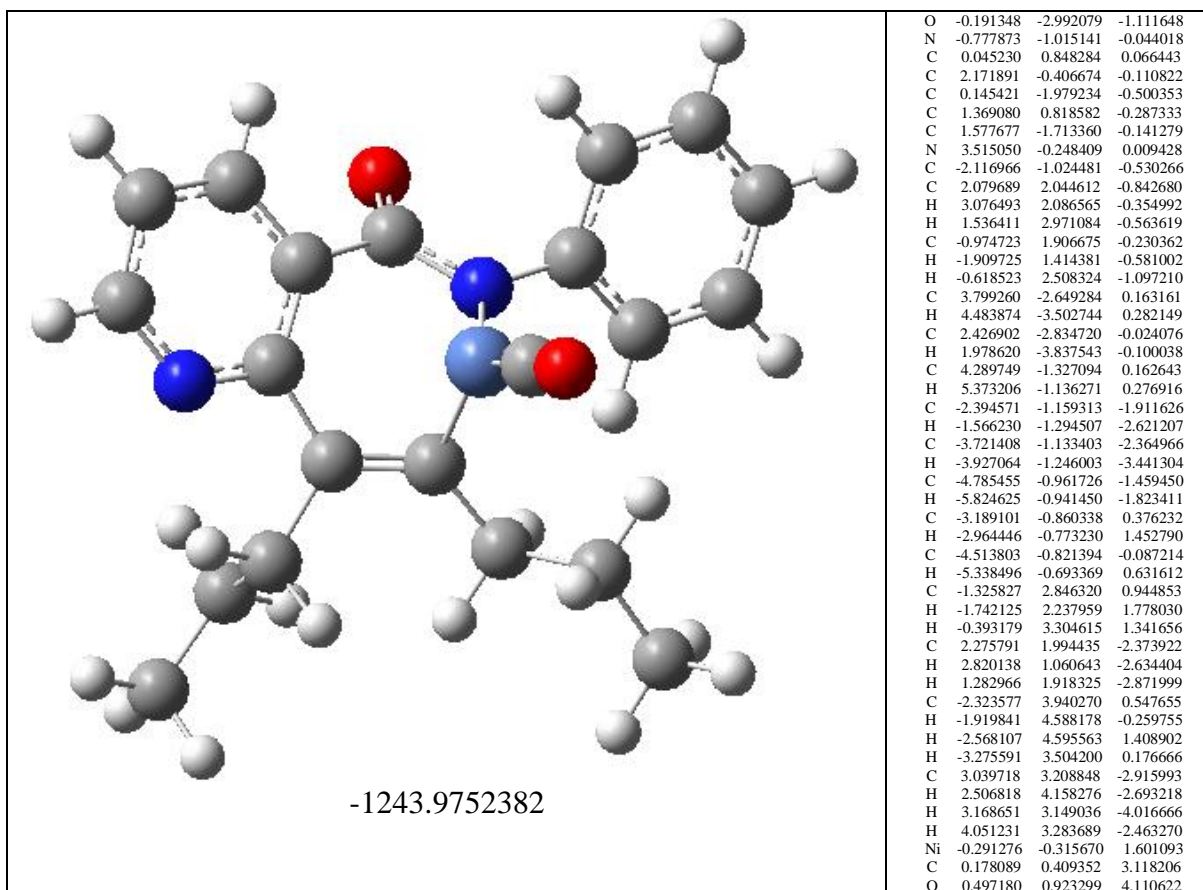


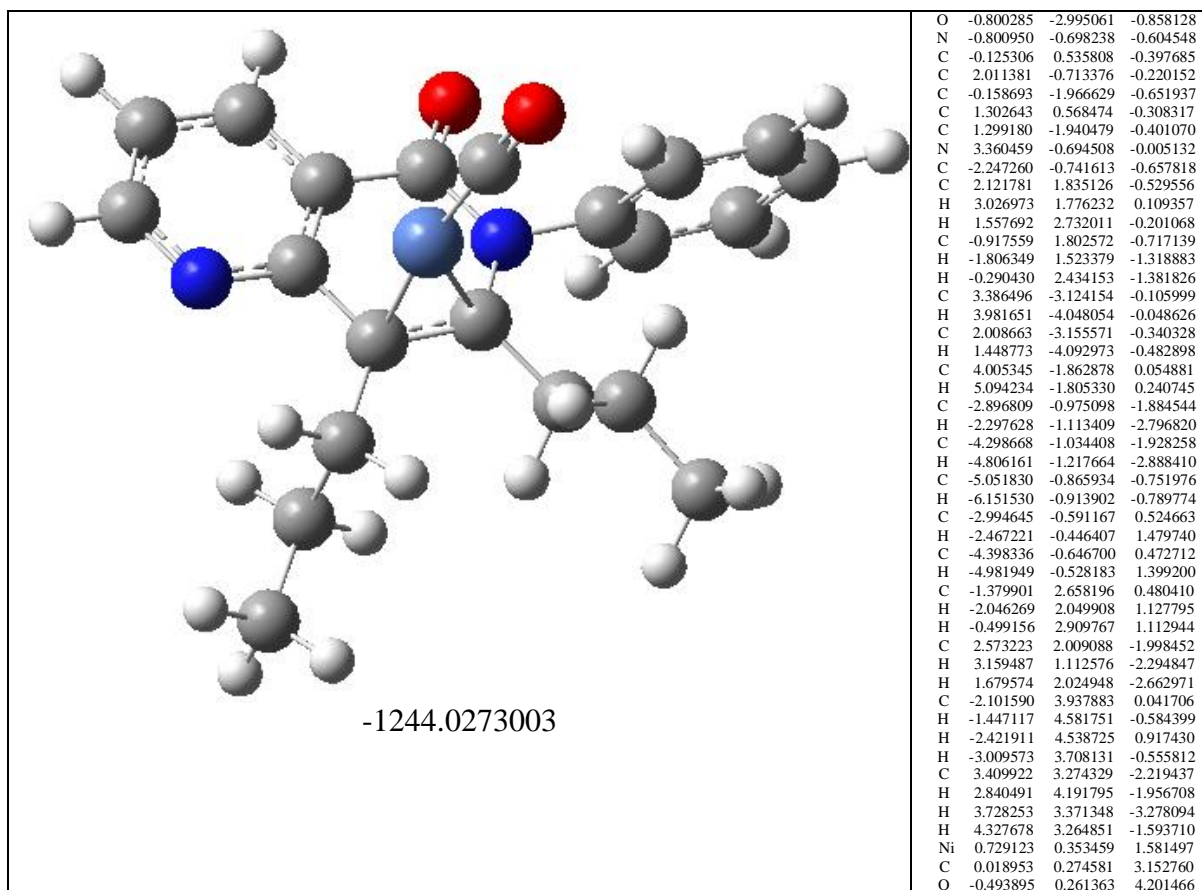
22-23

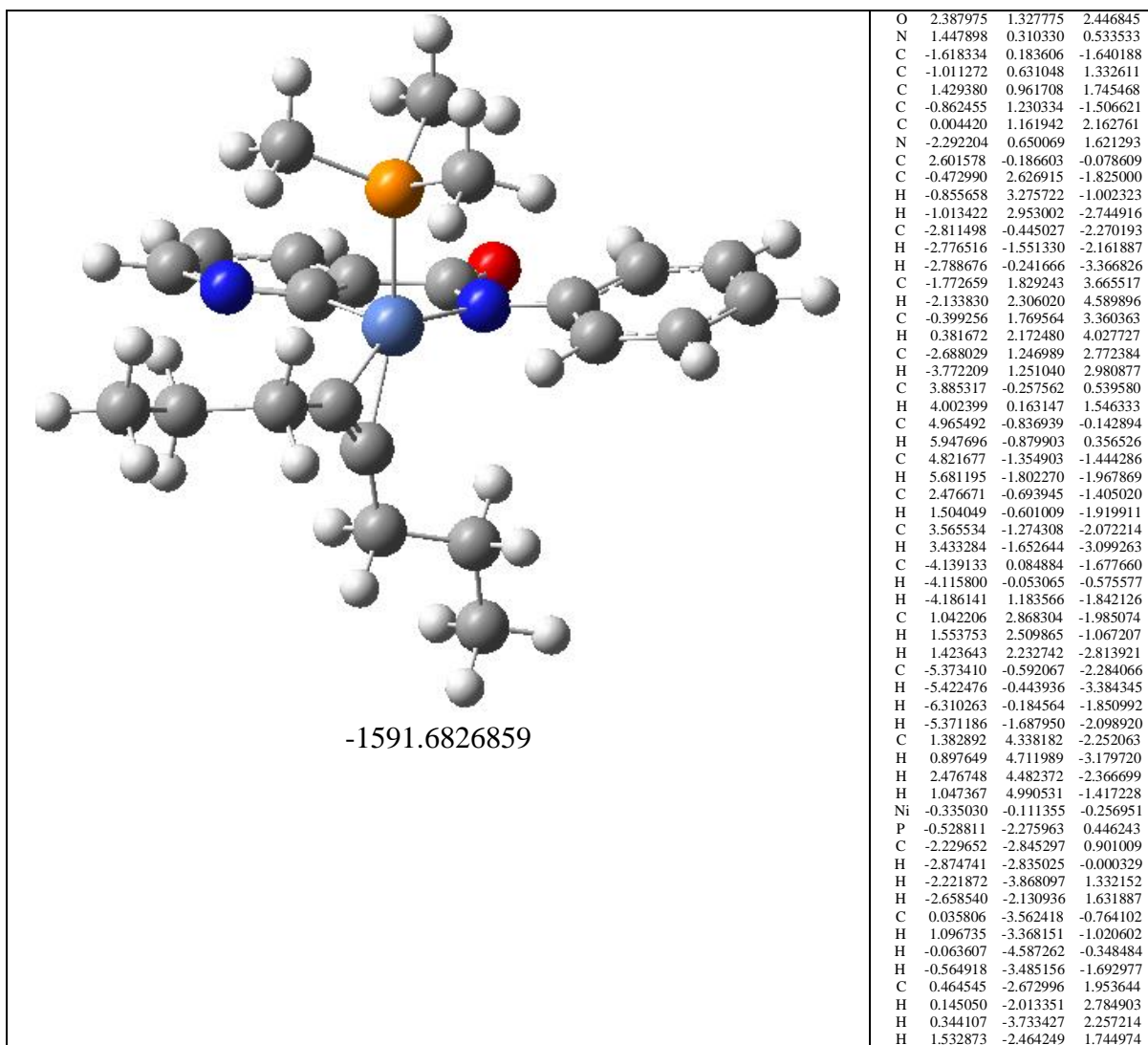




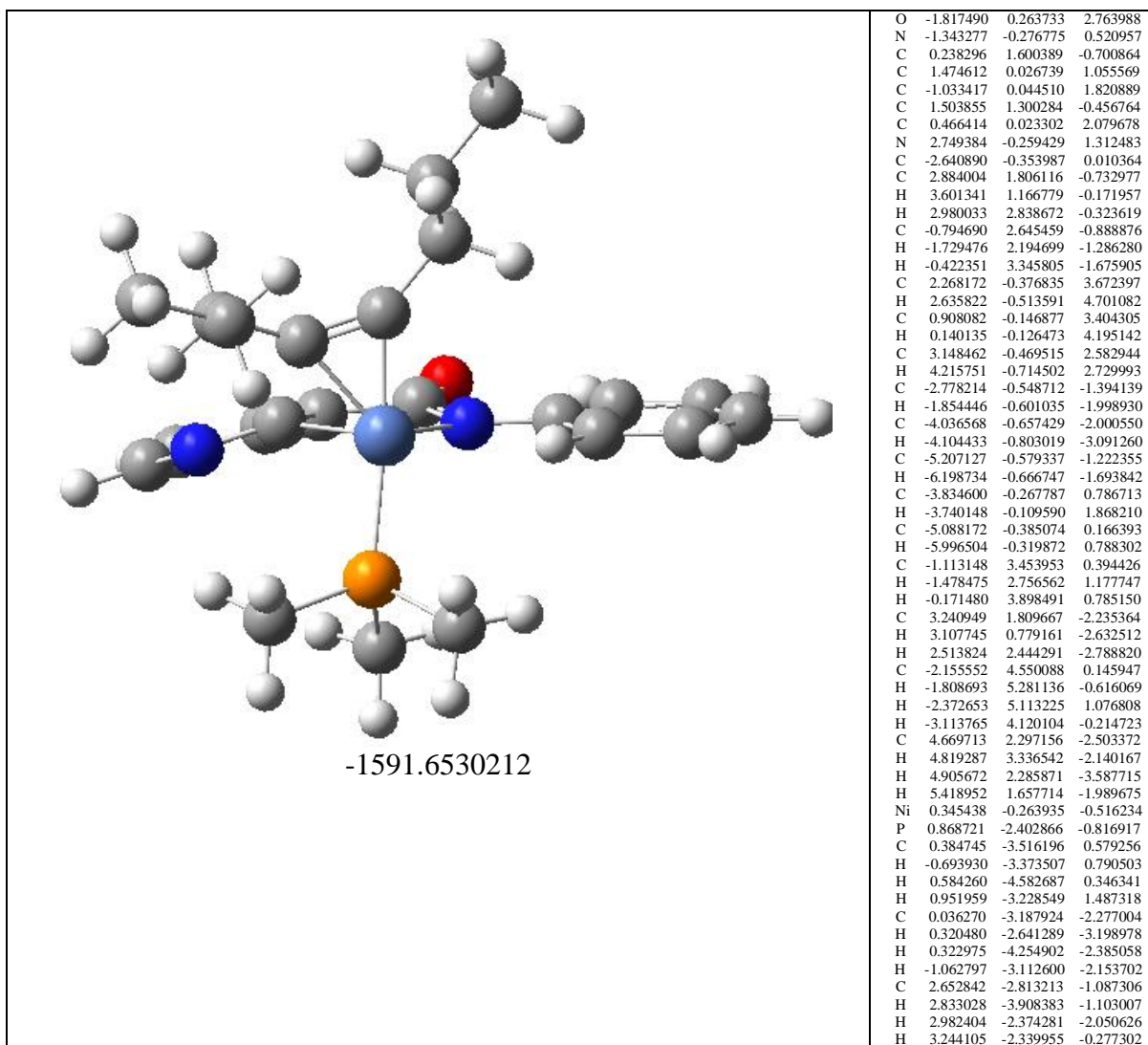
23-24

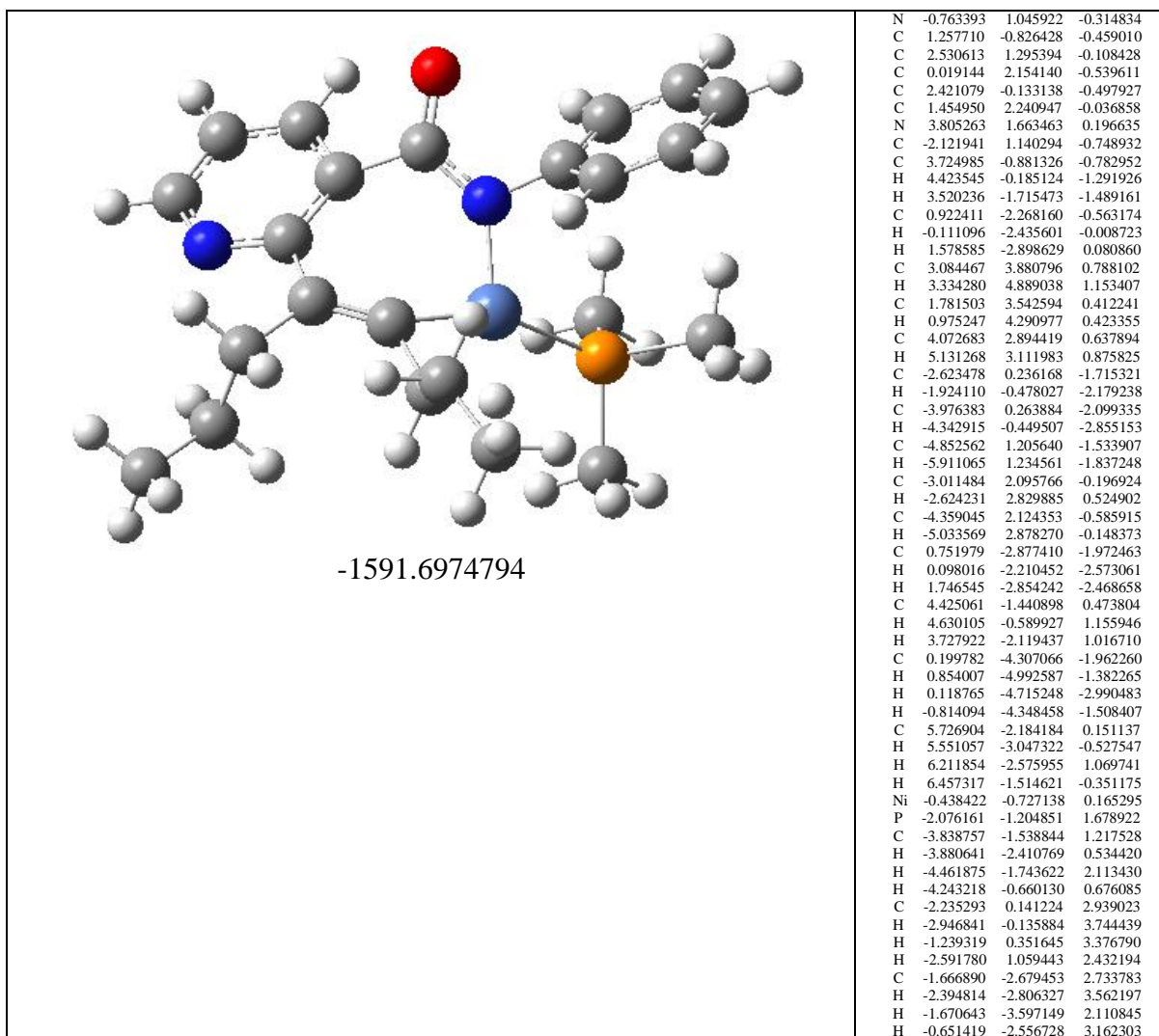




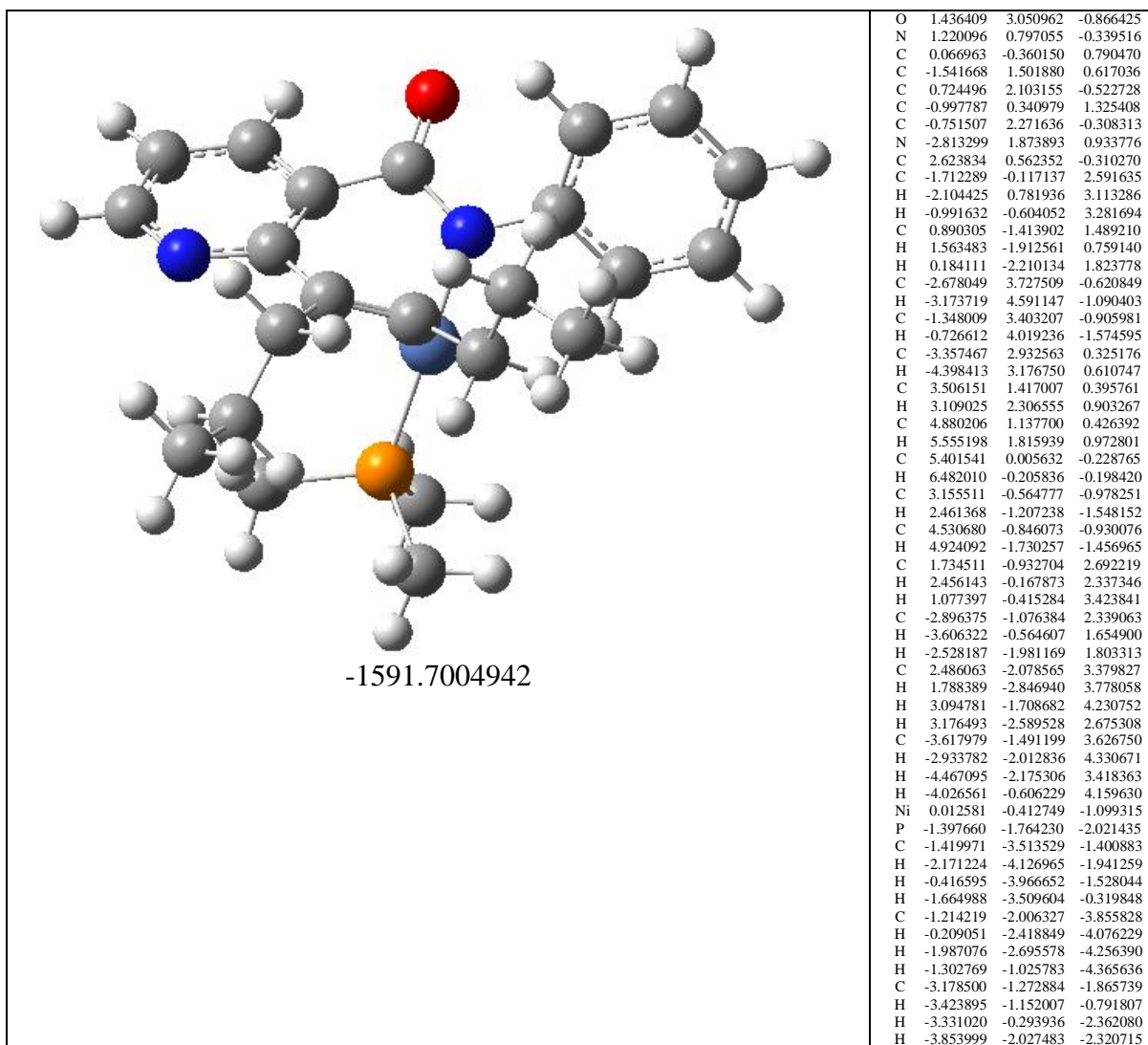


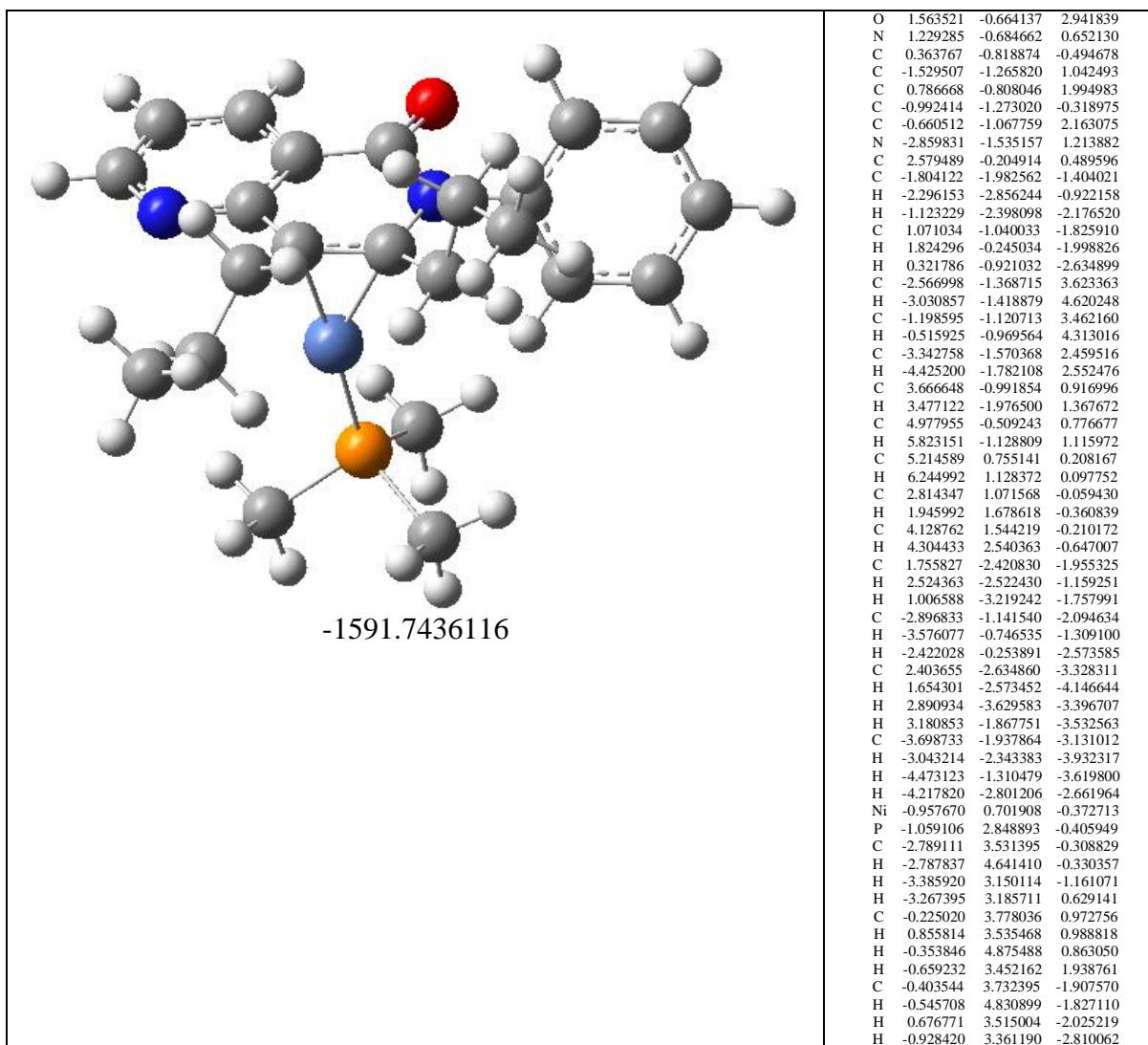
25-26

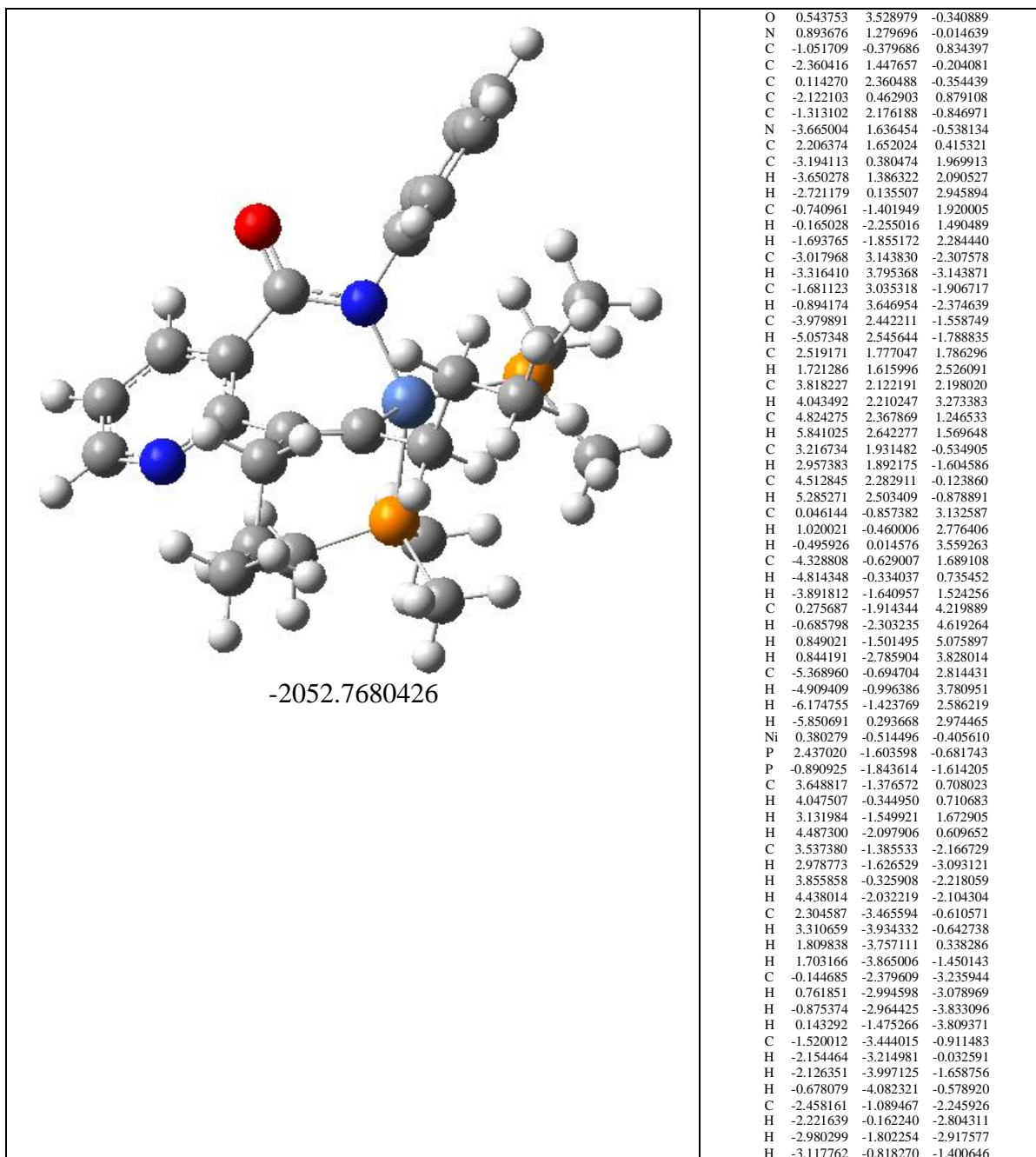




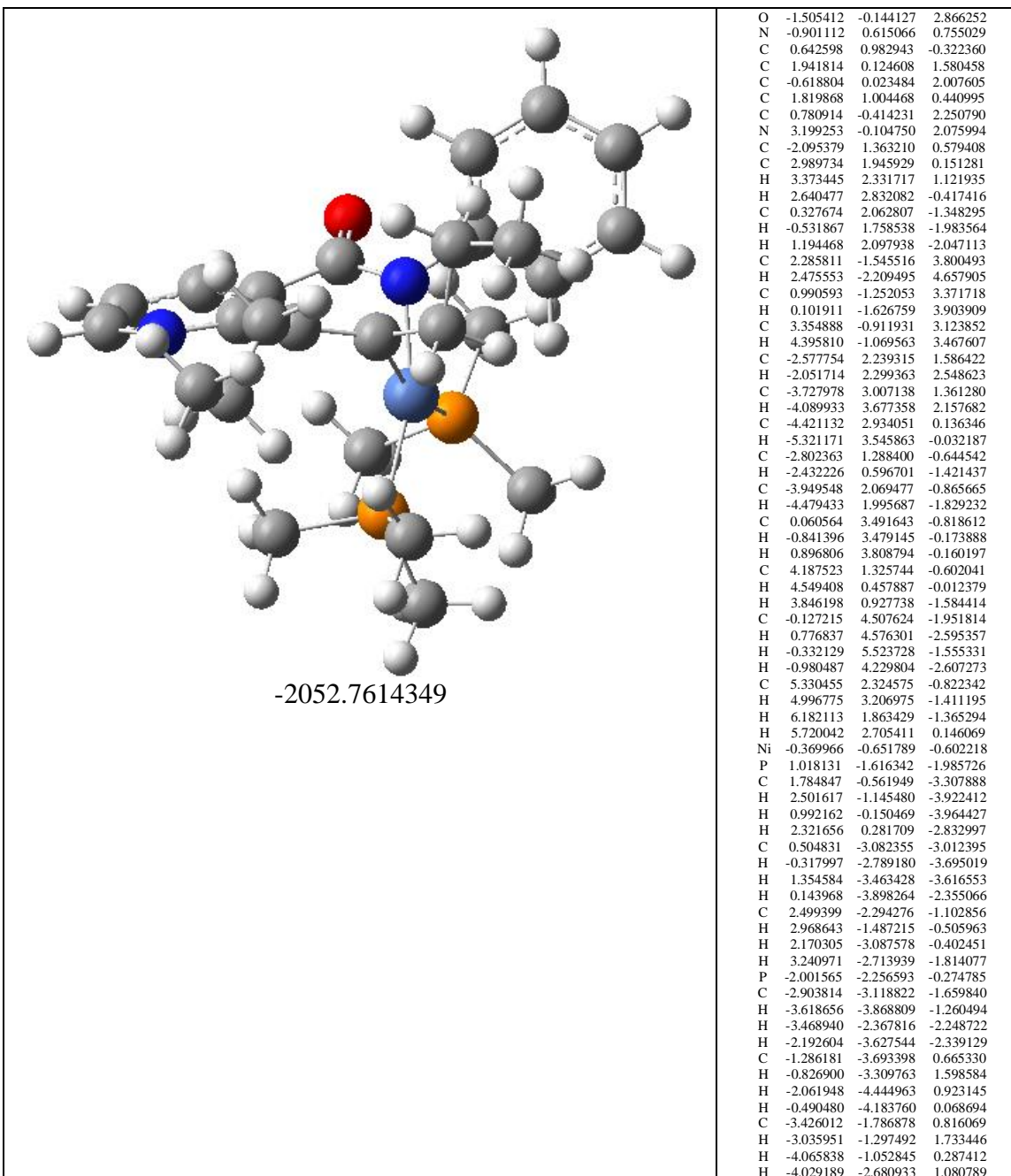
26-27

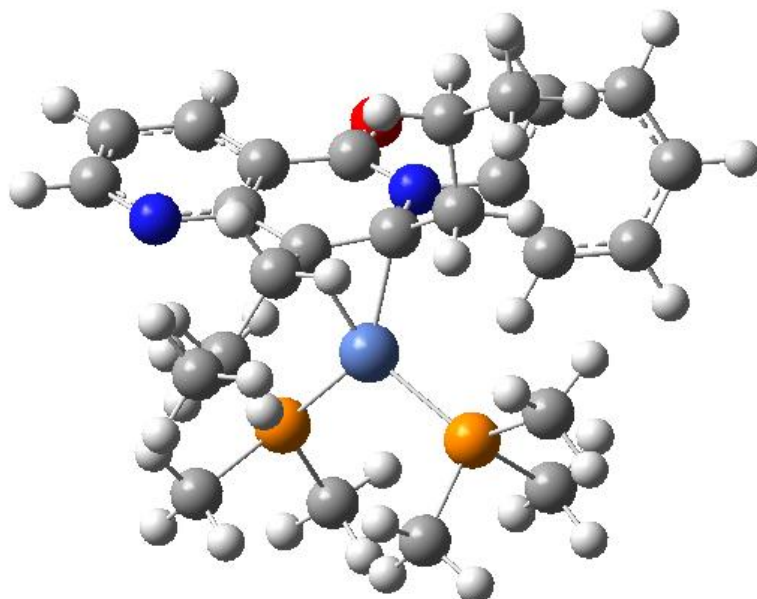






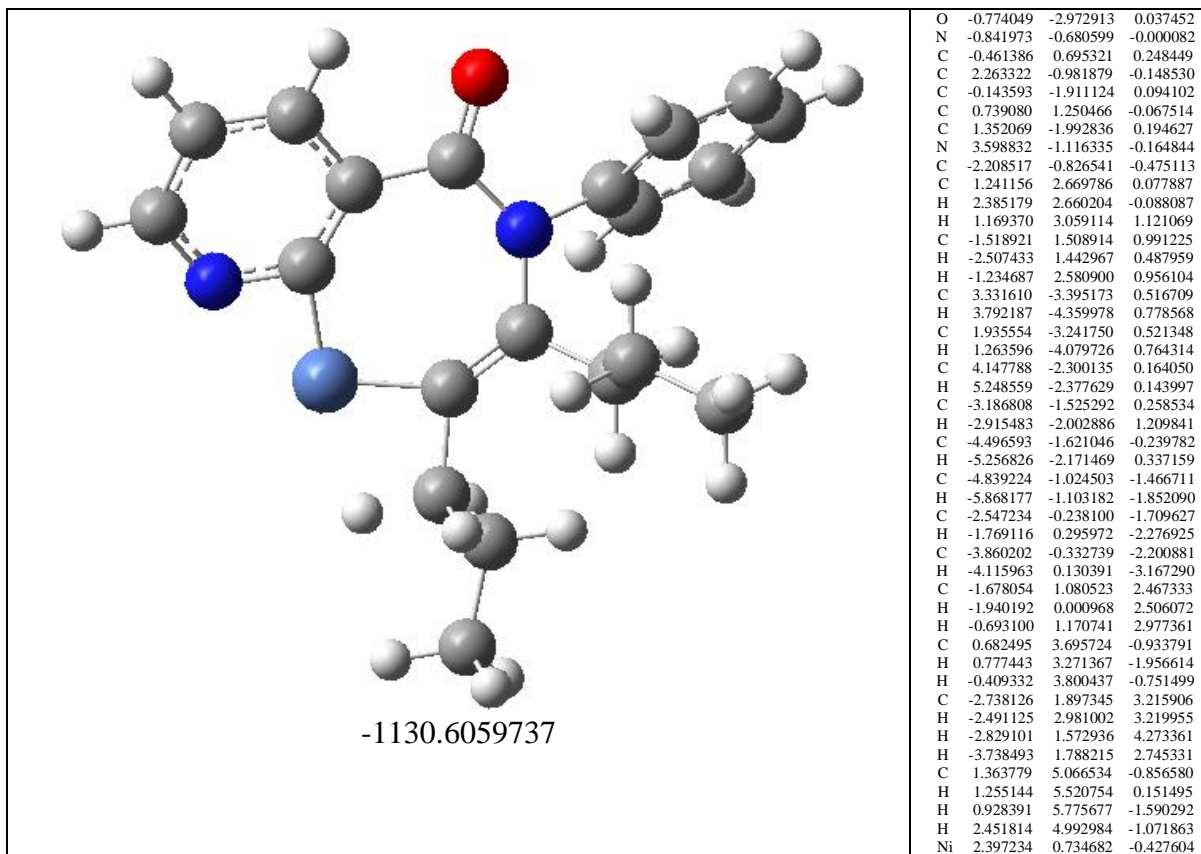
28-29

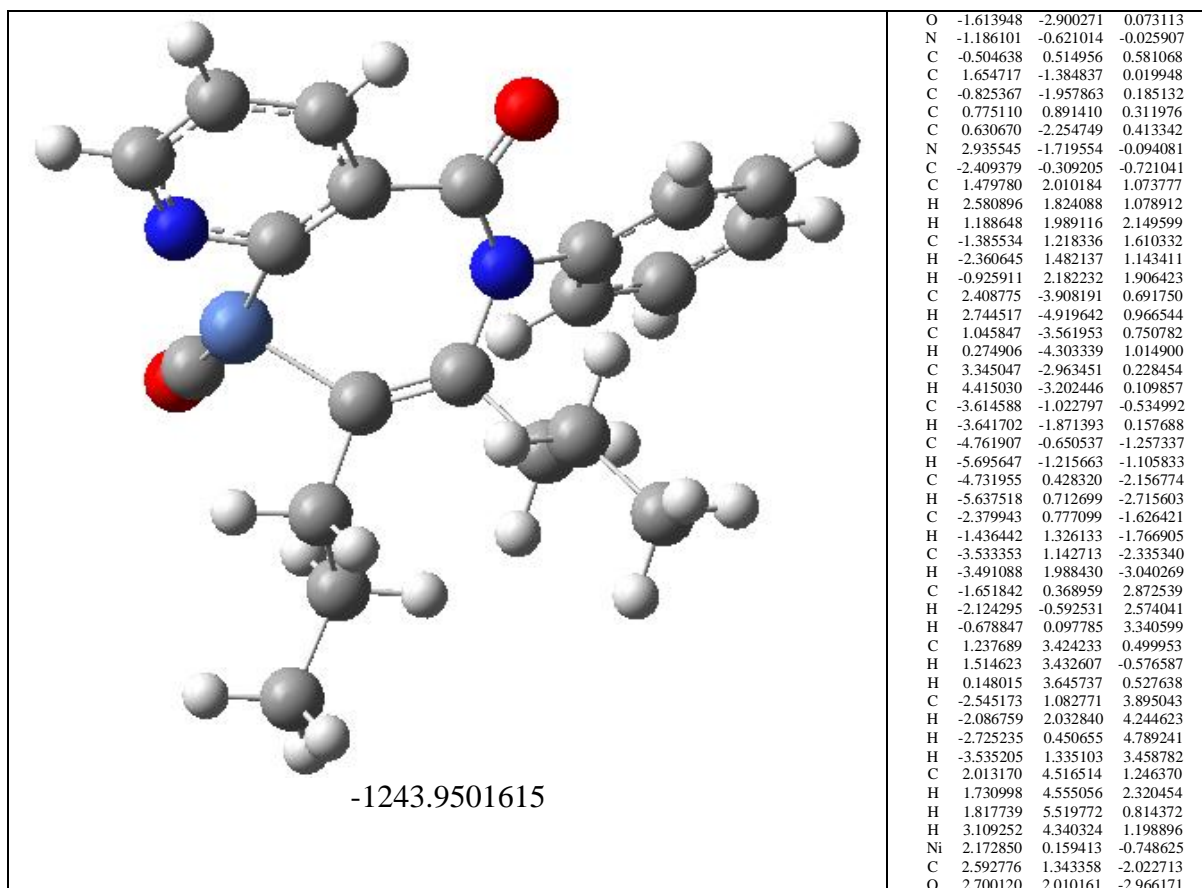


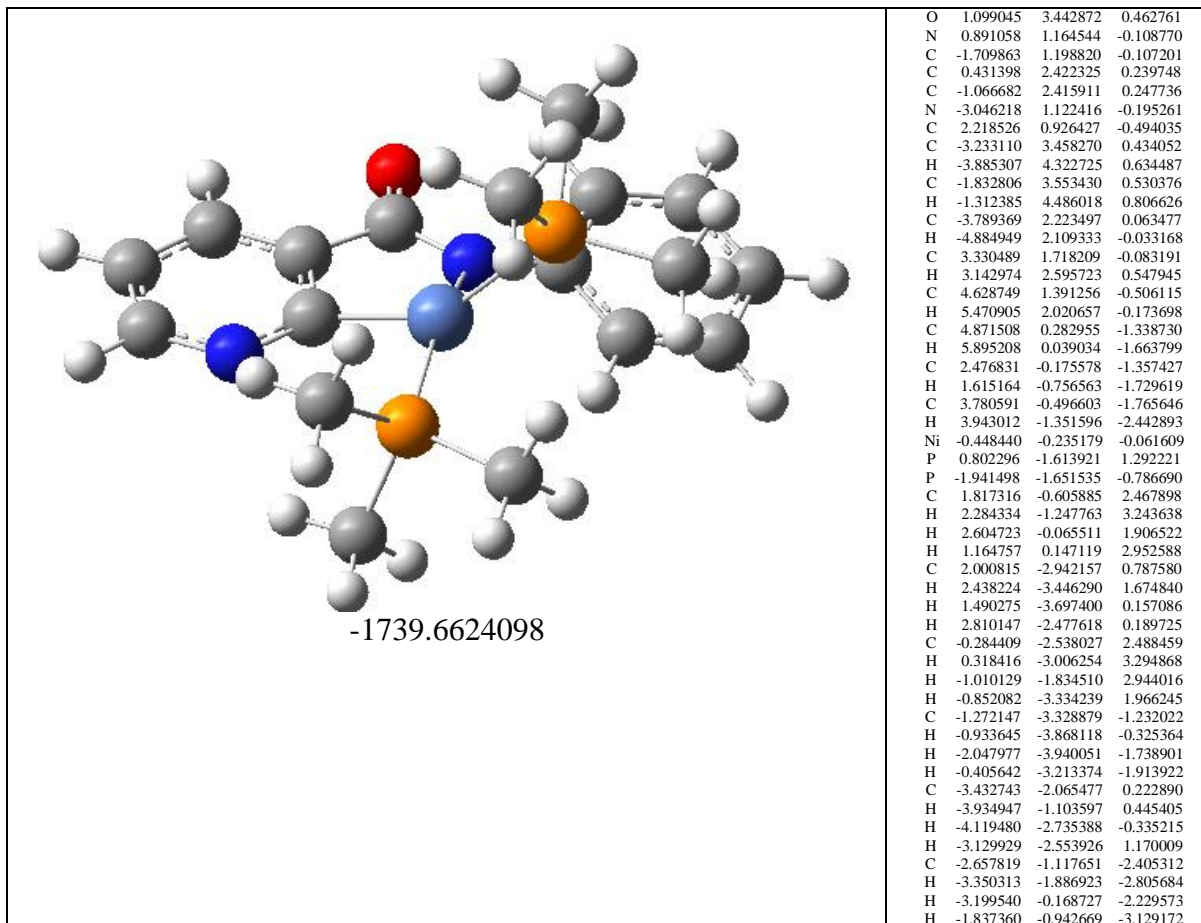


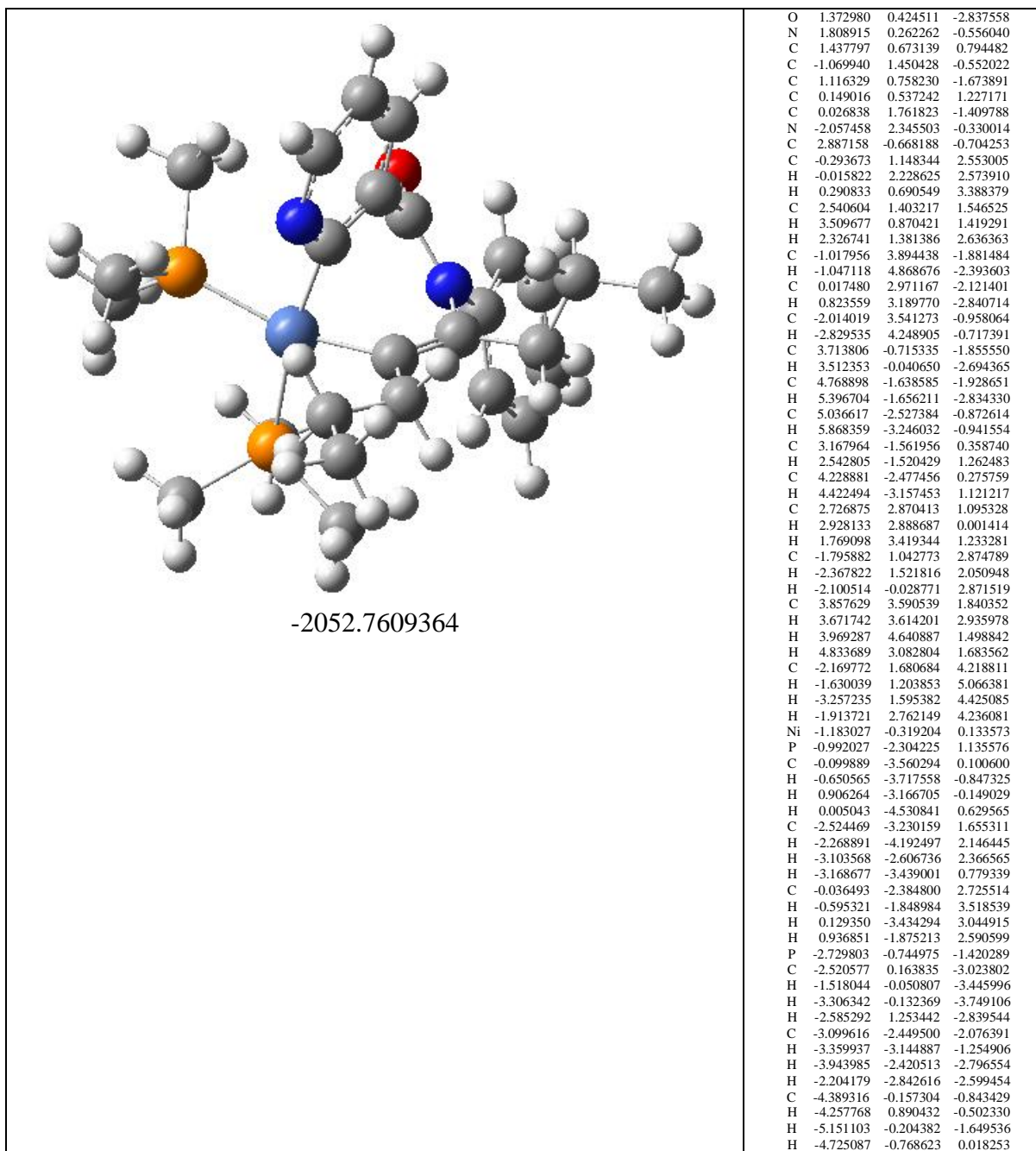
-2052.8113219

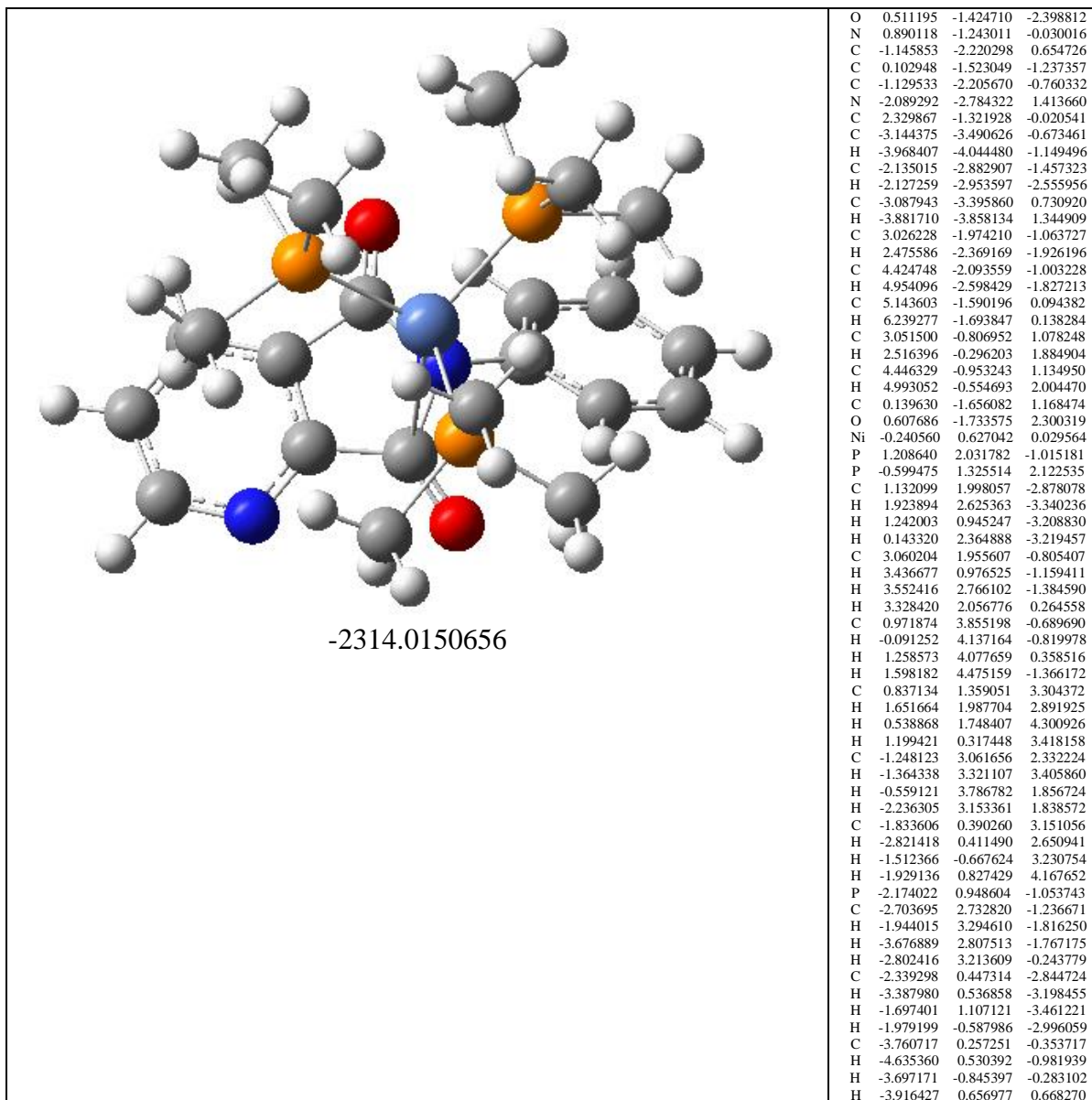
O	-1.353484	-2.794025	-1.934722
N	-1.286189	-1.224420	-0.243746
C	-0.622769	-0.517690	0.845410
C	1.380400	-1.964855	0.366290
C	-0.730844	-2.272814	-1.002857
C	0.779453	-0.825778	1.066099
C	0.635328	-2.687154	-0.612950
N	2.641076	-2.354520	0.711946
C	-2.620650	-0.832202	-0.636074
C	1.453804	-0.425795	2.385266
H	1.704634	-1.359564	2.939453
H	0.734320	0.130336	3.020416
C	-1.570728	-0.176870	1.992908
H	-2.486631	0.300535	1.585045
H	-1.101310	0.588010	2.641077
C	2.517815	-4.181440	-0.889242
H	3.014491	-5.046852	-1.353925
C	1.216434	-3.800970	-1.241791
H	0.621140	-4.338945	-1.996372
C	3.175975	-3.417433	0.097476
H	4.201140	-3.691998	0.412341
C	-3.717775	-1.677783	-0.381301
H	-3.551372	-2.645890	0.113545
C	-5.008668	-1.288732	-0.773852
H	-5.862621	-1.954486	-0.571282
C	-5.212138	-0.060865	-1.429350
H	-6.225845	0.239977	-1.738162
C	-2.820047	0.383402	-1.314445
H	-1.935574	1.002961	-1.532034
C	-4.113396	0.771938	-1.704056
H	-4.262548	1.723217	-2.240238
C	-1.970840	-1.395600	2.852863
H	-2.408775	-2.178104	2.195986
H	-1.053243	-1.848099	3.289253
C	2.740030	0.414470	2.273908
H	3.471976	-0.153396	1.663485
H	2.507649	1.347013	1.710355
C	-2.963028	-1.044385	3.968153
H	-2.543123	-0.284723	4.662512
H	-3.229243	-1.935680	4.573836
H	-3.906576	-0.628112	3.554516
C	3.346220	0.756209	3.640318
H	2.637117	1.334965	4.272369
H	4.272061	1.360824	3.539905
H	3.610756	-0.163907	4.204745
Ni	0.485504	0.844820	-0.118551
P	2.073057	0.992160	-1.664048
P	-0.335353	2.855858	0.169035
C	-1.468248	3.270259	1.594992
H	-1.657848	4.362978	1.656379
H	-2.435553	2.743041	1.475146
H	-1.007462	2.933131	2.545172
C	1.048950	4.055259	0.533628
H	1.732266	4.112415	-0.336843
H	0.663138	5.072733	0.756725
H	1.631924	3.688456	1.402237
C	-1.225430	3.793225	-1.183186
H	-1.382381	4.857186	-0.904599
H	-0.642683	3.746803	-2.124075
H	-2.214207	3.327009	-1.365720
C	1.780854	2.168304	-3.085888
H	0.812482	1.925320	-3.568821
H	1.730799	3.212997	-2.716929
H	2.587184	2.101389	-3.847144
C	3.788356	1.505894	-1.150123
H	4.477946	1.551374	-2.019655
H	3.753529	2.500284	-0.661381
H	4.176557	0.776083	-0.412614
C	2.457492	-0.555589	-2.620925
H	2.890274	-1.313006	-1.938863
H	1.523865	-0.974388	-3.046499
H	3.176441	-0.350357	-3.441964

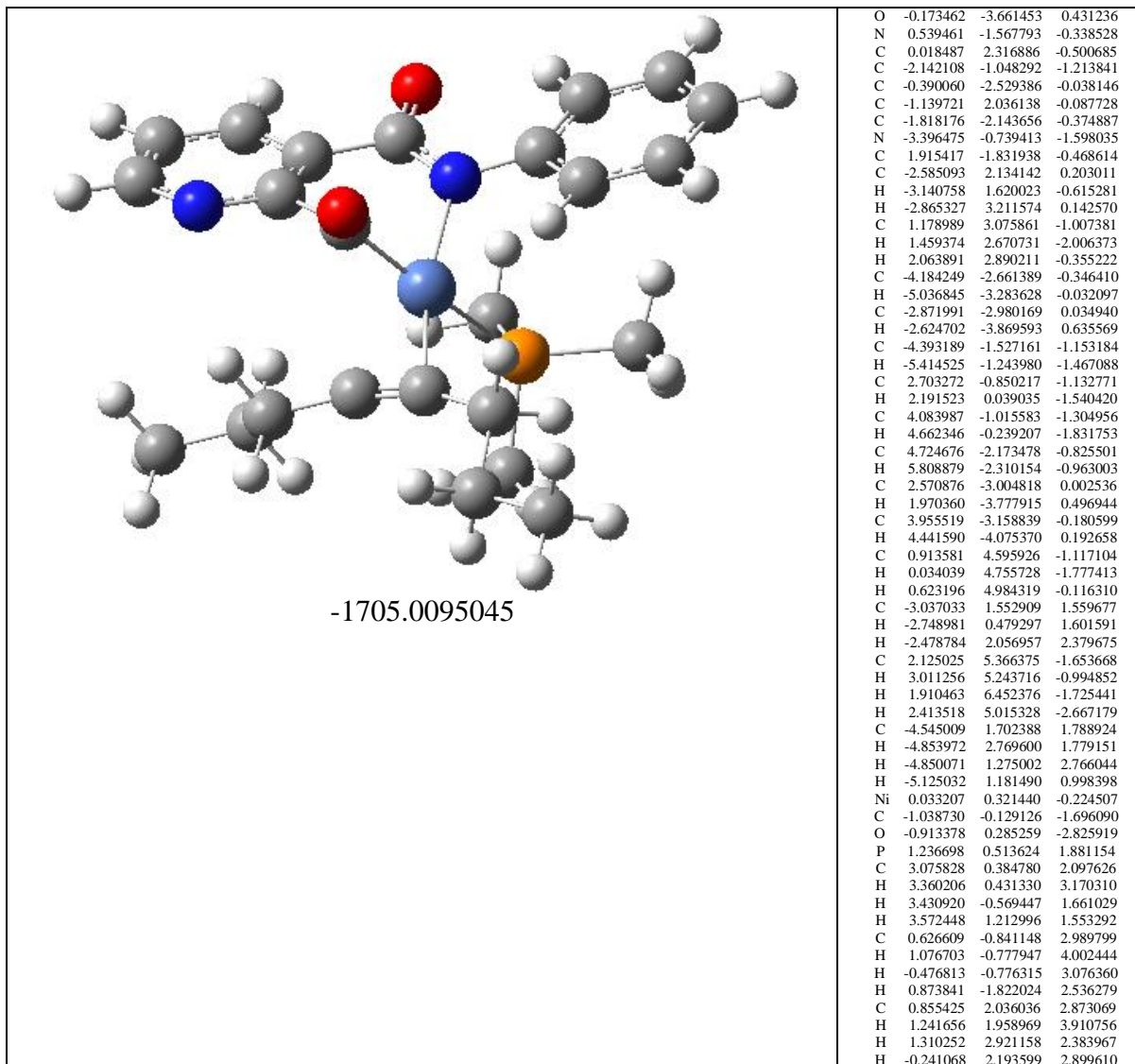




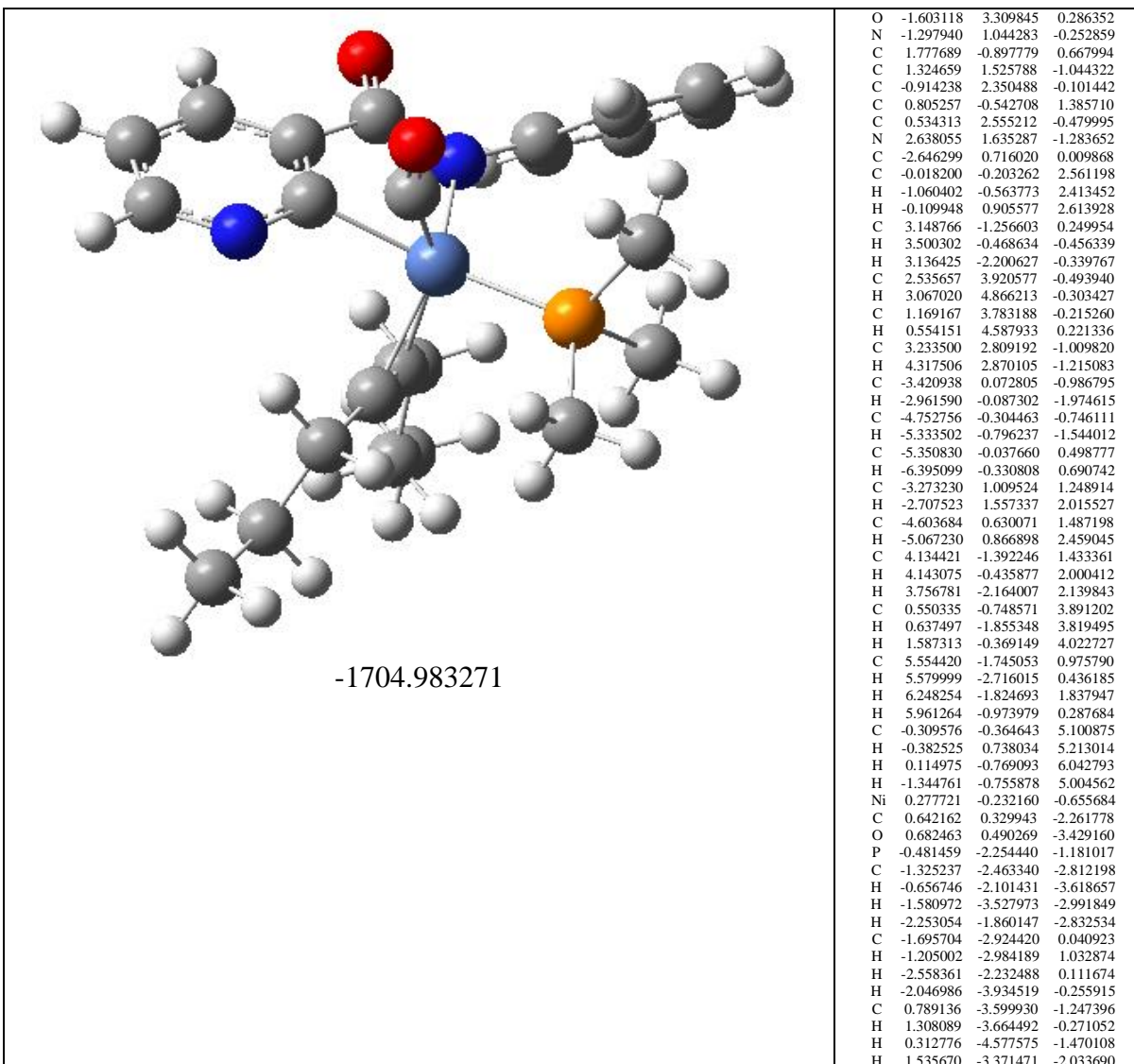


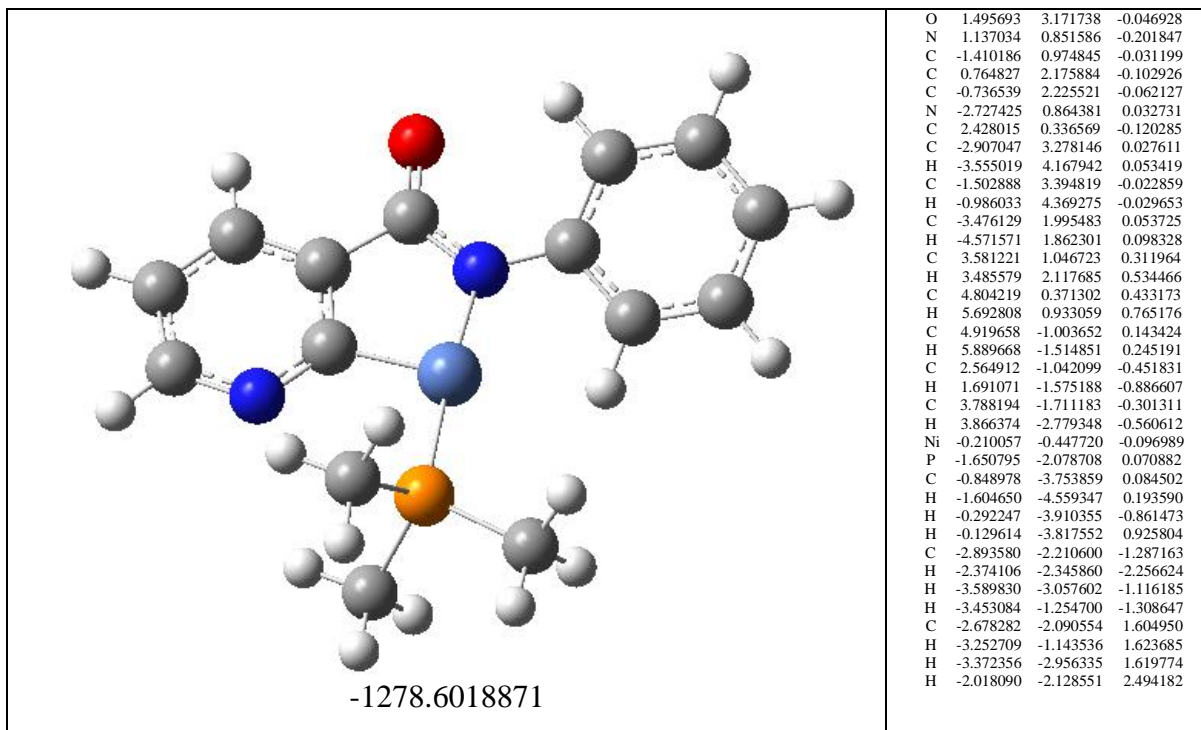






35-9





rmsd

Standard deviations for the distances and for the angles, $s_{n-1} = \sqrt{\frac{\sum_{i=1}^N (CV - EV)^2}{N - 1}}$, where CV means calculated value, EV experimental value (X-Ray data), and N is the number of distances or angles taken into account.

Table S2. Comparison between the X-Ray data of complex **2** and the corresponding computed geometry (Distances in Å and angles in degrees).

Distances	OPT	X-Ray
N2-O1	1.232	1.224
N2-C5	1.425	1.388
C5-C7	1.473	1.456
C7-C18	1.411	1.394
C18-C16	1.395	1.363
C16-C20	1.416	1.380
C20-N8	1.333	1.322
N8-C4	1.365	1.359
C4-C6	1.452	1.441
C6-C3	1.388	1.352
C3-N2	1.408	1.413
C7-C4	1.427	1.398
C6-C10	1.516	1.508
C10-C35	1.545	1.516
C35-C42	1.533	1.512
C3-C13	1.516	1.505
C13-C32	1.547	1.523
C32-C38	1.533	1.505
N2-C9	1.450	1.449
C9-C22	1.406	1.380
C22-C24	1.405	1.384
C24-C26	1.406	1.367
C26-C30	1.406	1.376
C30-C28	1.405	1.383
C28-C9	1.407	1.373

Angles	OPT	X-Ray
N2-C5-C7	114.5	115.3
C5-C7-C18	119.3	120.1
C7-C18-C16	118.9	119.3
C18-C16-C20	117.9	117.8
C16-C20-N8	124.3	125.3
C20-N8-C4	118.4	117.1
N8-C4-C6	118.7	118.5
C4-C6-C3	118.8	119.0
C6-C3-N2	120.9	120.6
C3-N2-C5	124.3	123.7
N2-C5-O1	120.8	121.1
C7-C5-O1	124.7	123.6
C3-C6-C10	123.1	122.4
C6-C3-C13	121.9	122.5
C5-N2-C9	114.2	115.7
C3-N2-C9	121.5	120.5
C5-C7-C4	121.6	120.9
N8-C4-C7	121.3	121.4
N2-C9-C22	119.9	119.5
C9-C22-C24	119.8	119.0
C24-C26-C30	119.9	120.2
C30-C28-C9	119.8	119.5
C6-C10-C35	113.8	114.3
C3-C13-C32	114.1	113.4
C10-C35-C42	113.0	112.4

- (ⁱ) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Scalmani, G.; Barone, V.; Mennucci, B.; Petersson, G. A.; Nakatsuji, H.; Caricato, M.; Li, X.; Hratchian, H. P.; Izmaylov, A. F.; Bloino, J.; Zheng, G.; Sonnenberg, J. L.; Hada, M.; Ehara, M.; Toyota, K.; Fukuda, R.; Hasegawa, J.; Ishida, M.; Nakajima, T.; Honda, Y.; Kitao, O.; Nakai, H.; Vreven, T.; Montgomery, Jr., J. A.; Peralta, J. E.; Ogliaro, F.; Bearpark, M.; Heyd, J. J.; Brothers, E.; Kudin, K. N.; Staroverov, V. N.; Kobayashi, R.; Normand, J.; Raghavachari, K.; Rendell, A.; Burant, J. C.; Iyengar, S. S.; Tomasi, J.; Cossi, M.; Rega, N.; Millam, J. M.; Klene, M.; Knox, J. E.; Cross, J. B.; Bakken, V.; Adamo, C.; Jaramillo, J.; Gomperts, R.; Stratmann, R. E.; Yazyev, O.; Austin, A. J.; Cammi, R.; Pomelli, C.; Ochterski, J. W.; Martin, R. L.; Morokuma, K.; Zakrzewski, V. G.; Voth, G. A.; Salvador, P.; Dannenberg, J. J.; Dapprich, S.; Daniels, A. D.; Farkas, Ö.; Foresman, J. B.; Ortiz, J. V.; Cioslowski, J.; Fox, D. J. Gaussian, Inc., Wallingford CT, 2009.
- (ⁱⁱ) (a) Perdew, J. P. *Phys. Rev. B* **1986**, *33*, 8822-8824. (b) Perdew, J. P. *Phys. Rev. B* **1986**, *34*, 7406-7406. (c) Becke, A. D. *Phys. Rev. A* **1988**, *38*, 3098-3100.
- (ⁱⁱⁱ) (a) Schaefer, A.; Horn, H.; Ahlrichs, R. *J. Chem. Phys.* **1992**, *97*, 2571-2577. (b) Schaefer, A.; Huber, C.; Ahlrichs, R. *J. Chem. Phys.* **1994**, *100*, 5829-5835.
- (^{iv}) (a) Haeusermann, U.; Dolg, M.; Stoll, H.; Preuss, H. *Mol. Phys.* **1993**, *78*, 1211-1224. (b) Kuechle, W.; Dolg, M.; Stoll, H.; Preuss, H. *J. Chem. Phys.* **1994**, *100*, 7535-7542. (c) Leininger, T.; Nicklass, A.; Stoll, H.; Dolg, M.; Schwerdtfeger, P. *J. Chem. Phys.* **1996**, *105*, 1052-1059.
- (v) Zhao, Y.; Truhlar, D. *Theor. Chem. Acc.* **2008**, *120*, 215-241.
- (vi) F. Weigend, R. Ahlrichs, *Phys. Chem. Chem. Phys.* **2005**, *7*, 3297-3305.
- (^{vii}) (a) Barone, V.; Cossi, M. *J. Phys. Chem. A* **1998**, *102*, 1995-2001. (b) Tomasi, J.; Persico, M. *Chem. Rev.* **1994**, *94*, 2027-2094.
- (^{viii}) (a) Kruszewski, J.; Krygowski, T. M. *Tetrahedron Lett.* **1972**, 3839-3842. (b) Krygowski, T. M. *J. Chem. Inf. Comp. Sci.* **1993**, *33*, 70-78.
- (^{ix}) (a) Poater, A.; Cavallo, L. *Inorg. Chem.* **2009**, *48*, 2340-2342. (b) Poater, A.; Ribas, X.; Llobet, A.; Cavallo, L.; Solà, M. *J. Am. Chem. Soc.* **2008**, *130*, 17710-17717.

-
- ^(c) Wiberg, K. B. *Tetrahedron* **1968**, *24*, 1083-1096.
- ^(d) (a) Mayer, I. *Chem. Phys. Lett.* **1983**, *97*, 270-277. (b) Mayer, I. *Int. J. Quantum Chem.* **1984**, *26*, 151-154.
- ^(e) (a) Bridgeman, A. J.; Harris, N.; Young, N. A. *Chem. Commun.* **2000**, 1241-1242. (b) Bridgeman, A. J. *J. Chem. Soc., Dalton Trans.* **1997**, 1323-1329.
- ^(f) (a) Bridgeman, A. J. *J. Chem. Soc., Dalton Trans.* **1996**, 2601-2607. (b) Poater, A.; Francesco, R.; Correa, A.; Cavallo, L. *Dalton Trans.* **2011**, *40*, 11066-11069. (c) Ramón, R. S.; Gaillard, S.; Poater, A.; Cavallo, L.; Slawin, A. M. Z.; Nolan, S. P. *Chem. Eur. J.* **2011**, *17*, 1238-1246. (d) Poater, A.; Cavallo, L. *J. Mol. Catal. A* **2010**, *324*, 75-79.